



Universiteit
Leiden

The Netherlands

Getting the electrons right for O₂-on-metal systems

Bree, R.A.B. van

Citation

Bree, R. A. B. van. (2025, October 21). *Getting the electrons right for O₂-on-metal systems*. Retrieved from <https://hdl.handle.net/1887/4279459>

Version: Publisher's Version

License: [Licence agreement concerning inclusion of doctoral thesis in the Institutional Repository of the University of Leiden](#)

Downloaded from: <https://hdl.handle.net/1887/4279459>

Note: To cite this publication please use the final published version (if applicable).

References

- (1) Ertl, G. Reactions at Surfaces: From Atoms to Complexity (Nobel Lecture). *Angew. Chemie Int. Ed.* **2008**, 47 (19), 3524–3535. <https://doi.org/10.1002/anie.200800480>.
- (2) Palmer, T.; Bonner, P. L. *Enzymes: Biochemistry, Biotechnology, Clinical Chemistry*; Elsevier, **2007**.
- (3) Braslavsky, S. E. Glossary of Terms Used in Photochemistry, 3rd Edition (IUPAC Recommendations 2006). *Pure Appl. Chem.* **2007**, 79 (3), 293–465. <https://doi.org/10.1351/pac200779030293>.
- (4) Smil, V. Detonator of the Population Explosion. *Nature* **1999**, 400 (6743), 415–415. <https://doi.org/10.1038/22672>.
- (5) Ertl, G. Primary Steps in Catalytic Synthesis of Ammonia. *J. Vac. Sci. Technol. A Vacuum, Surfaces, Film.* **1983**, 1 (2), 1247–1253. <https://doi.org/10.1116/1.572299>.
- (6) Ertl, G. Elementary Steps in Heterogeneous Catalysis. *Angew. Chemie Int. Ed. English* **1990**, 29 (11), 1219–1227. <https://doi.org/10.1002/anie.199012191>.
- (7) Bagot, P. A. J. Fundamental Surface Science Studies of Automobile Exhaust Catalysis. *Mater. Sci. Technol.* **2004**, 20 (6), 679–694. <https://doi.org/10.1179/026708304225016789>.
- (8) van Dishoeck, E. F. Astrochemistry: Overview and Challenges. *Proc. Int. Astron. Union* **2017**, 13 (S332), 3–22. <https://doi.org/10.1017/S1743921317011528>.
- (9) Catalyst. In *The IUPAC Compendium of Chemical Terminology*; International Union of Pure and Applied Chemistry (IUPAC): Research Triangle Park, NC, **2019**. <https://doi.org/10.1351/goldbook.C00876>.
- (10) Engel, T.; Reid, P. *Physical Chemistry*, 3rd ed.; Pearson Education Limited: Harlow, 2014.
- (11) Taylor, H. S. Catalysis. *Encyclopedia Britannica.* **2024**.
- (12) Wisniak, J. The History of Catalysis. From the Beginning to Nobel Prizes. *Educ. Química* **2010**, 21 (1), 60–69. [https://doi.org/10.1016/S0187-893X\(18\)30074-0](https://doi.org/10.1016/S0187-893X(18)30074-0).
- (13) Fulhame, E. *An Essay on Combustion: With a View to a New Art of Dying and Painting : Wherein the Phlogistic and Antiphlogistic Hypotheses Are Proved Erroneous*; author, **1794**.
- (14) Li, Y.; Somorjai, G. A. Nanoscale Advances in Catalysis and Energy Applications. *Nano Lett.* **2010**, 10 (7), 2289–2295. <https://doi.org/10.1021/nl101807g>.
- (15) Eley, D. D.; Haag, W. O.; Gates, B. C.; Knoezinger, H. *Advances in Catalysis*; ISSN; Elsevier Science, **1998**.
- (16) Cheng, K.; Kang, J.; King, D. L.; Subramanian, V.; Zhou, C.; Zhang, Q.; Wang, Y. Advances in Catalysis for Syngas Conversion to Hydrocarbons; **2017**; pp 125–208. <https://doi.org/10.1016/bs.acat.2017.09.003>.
- (17) Daza, Y. A.; Kuhn, J. N. CO₂ Conversion by Reverse Water Gas Shift Catalysis: Comparison of Catalysts, Mechanisms and Their Consequences for CO₂ Conversion to Liquid Fuels. *RSC Adv.* **2016**, 6 (55), 49675–49691. <https://doi.org/10.1039/c6ra05414e>.
- (18) Kaiser, P.; Unde, R. B.; Kern, C.; Jess, A. Production of Liquid Hydrocarbons with CO₂ as Carbon Source Based on Reverse Water-Gas Shift and Fischer-Tropsch Synthesis. *Chemie-Ingenieur-Technik* **2013**, 85 (4), 489–499. <https://doi.org/10.1002/cite.201200179>.
- (19) Birdja, Y. Y.; Pérez-Gallent, E.; Figueiredo, M. C.; Göttle, A. J.; Calle-Vallejo, F.; Koper, M. T. M. Advances and Challenges in Understanding the Electrocatalytic Conversion of Carbon Dioxide to Fuels. *Nat. Energy* **2019**, 4 (9), 732–745. <https://doi.org/10.1038/s41560-019-0450-y>.
- (20) Angelici, R. J. Heterogeneous Catalysis of the Hydrodesulfurization of Thiophenes in

- Petroleum: An Organometallic Perspective of the Mechanism. *Acc. Chem. Res.* **1988**, *21* (11), 387–394. <https://doi.org/10.1021/ar00155a001>.
- (21) Hettinger, W. P. Contribution to Catalytic Cracking in the Petroleum Industry. *Appl. Clay Sci.* **1991**, *5* (5–6), 445–468. [https://doi.org/10.1016/0169-1317\(91\)90017-4](https://doi.org/10.1016/0169-1317(91)90017-4).
 - (22) Satterfield, C. N. Heterogeneous Catalysis in Industrial Practice. 2nd Edition.
 - (23) Fatima, S. S.; Zuraiqi, K.; Zavabeti, A.; Krishnamurthi, V.; Kalantar-Zadeh, K.; Chiang, K.; Daeneke, T. Current State and Future Prospects of Liquid Metal Catalysis. *Nat. Catal.* **2023**, *6* (12), 1131–1139. <https://doi.org/10.1038/s41929-023-01083-3>.
 - (24) Schlögl, R. Heterogeneous Catalysis. *Angew. Chemie Int. Ed.* **2015**, *54* (11), 3465–3520. <https://doi.org/10.1002/anie.201410738>.
 - (25) Fechet, I.; Wang, Y.; Védrine, J. C. The Past, Present and Future of Heterogeneous Catalysis. *Catal. Today* **2012**, *189* (1), 2–27. <https://doi.org/10.1016/j.cattod.2012.04.003>.
 - (26) Smith, C.; Hill, A. K.; Torrente-Murciano, L. Current and Future Role of Haber–Bosch Ammonia in a Carbon-Free Energy Landscape. *Energy Environ. Sci.* **2020**, *13* (2), 331–344. <https://doi.org/10.1039/C9EE02873K>.
 - (27) Waugh, K. C. Methanol Synthesis. *Catal. Today* **1992**, *15* (1), 51–75. [https://doi.org/10.1016/0920-5861\(92\)80122-4](https://doi.org/10.1016/0920-5861(92)80122-4).
 - (28) Rostrup-Nielsen, J. R.; Sehested, J.; Nørskov, J. K. Hydrogen and Synthesis Gas by Steam- and CO₂ Reforming; Advances in Catalysis; Academic Press, **2002**; Vol. 47, pp 65–139. [https://doi.org/10.1016/S0360-0564\(02\)47006-X](https://doi.org/10.1016/S0360-0564(02)47006-X).
 - (29) Ali, K. A.; Abdullah, A. Z.; Mohamed, A. R. Recent Development in Catalytic Technologies for Methanol Synthesis from Renewable Sources: A Critical Review. *Renew. Sustain. Energy Rev.* **2015**, *44*, 508–518. <https://doi.org/10.1016/j.rser.2015.01.010>.
 - (30) Zhou, G.; Dai, B.; Xie, H.; Zhang, G.; Xiong, K.; Zheng, X. CeCu Composite Catalyst for CO Synthesis by Reverse Water-Gas Shift Reaction: Effect of Ce/Cu Mole Ratio. *J. CO₂ Util.* **2017**, *21* (April), 292–301. <https://doi.org/10.1016/j.jcou.2017.07.004>.
 - (31) Chen, C. S. S.; Cheng, W. H. H.; Lin, S. S. Mechanism of CO Formation in Reverse Water-Gas Shift Reaction over Cu/Al₂O₃ Catalyst. *Catal. Letters* **2000**, *68* (1–2), 45–48. <https://doi.org/10.1023/A:1019071117449>.
 - (32) Bustamante, F.; Enick, R. M. M.; Cugini, A. V. V.; Killmeyer, R. P. P.; Howard, B. H. H.; Rothenberger, K. S. S.; Ciocco, M. V. V.; Morreale, B. D. D.; Chattopadhyay, S.; Shi, S. High-Temperature Kinetics of the Homogeneous Reverse Water-Gas Shift Reaction. *AIChE J.* **2004**, *50* (5), 1028–1041. <https://doi.org/10.1002/aic.10099>.
 - (33) Iglesia, E. Design, Synthesis, and Use of Cobalt-Based Fischer-Tropsch Synthesis Catalysts. *Appl. Catal. A Gen.* **1997**, *161* (1–2), 59–78. [https://doi.org/10.1016/S0926-860X\(97\)00186-5](https://doi.org/10.1016/S0926-860X(97)00186-5).
 - (34) Dry, M. E. The Fischer – Tropsch Process : 1950 – 2000. **2002**, *71*, 227–241.
 - (35) Jager, B.; Espinoza, R. Advances in Low Temperature Fischer-Tropsch Synthesis. *Catal. Today* **1995**, *23* (1), 17–28. [https://doi.org/10.1016/0920-5861\(94\)00136-P](https://doi.org/10.1016/0920-5861(94)00136-P).
 - (36) Schulz, H. Short History and Present Trends of Fischer-Tropsch Synthesis. *Appl. Catal. A Gen.* **1999**, *186* (1–2), 3–12. [https://doi.org/10.1016/S0926-860X\(99\)00160-X](https://doi.org/10.1016/S0926-860X(99)00160-X).
 - (37) Weckhuysen, B. M. Determining the Active Site in a Catalytic Process: Operando Spectroscopy Is More than a Buzzword. *Phys. Chem. Chem. Phys.* **2003**, *5* (20), 4351. <https://doi.org/10.1039/b309650p>.
 - (38) Bentrup, U. Combining in Situ Characterization Methods in One Set-up: Looking with More Eyes into the Intricate Chemistry of the Synthesis and Working of Heterogeneous Catalysts. *Chem. Soc. Rev.* **2010**, *39* (12), 4718. <https://doi.org/10.1039/b919711g>.
 - (39) Zambelli, T.; Barth, J. V. V.; Wintterlin, J.; Ertl, G. Complex Pathways in Dissociative Adsorption of Oxygen on Platinum. *Nature* **1997**, *390* (6659), 495–497.

- <https://doi.org/10.1038/37329>.
- (40) Diebold, U. The Surface Science of Titanium Dioxide. *Surf. Sci. Rep.* **2003**, *48* (5–8), 53–229. [https://doi.org/10.1016/S0167-5729\(02\)00100-0](https://doi.org/10.1016/S0167-5729(02)00100-0).
 - (41) Diebold, U.; Li, S.-C.; Schmid, M. Oxide Surface Science. *Annu. Rev. Phys. Chem.* **2010**, *61* (1), 129–148. <https://doi.org/10.1146/annurev.physchem.012809.103254>.
 - (42) Oura, K.; Lifshits, V. G.; Saranin, A. A.; Zotov, A. V.; Katayama, M. *Surface Science: An Introduction*; Advanced Texts in Physics; Springer Berlin Heidelberg, **2013**.
 - (43) Markovic, N. Surface Science Studies of Model Fuel Cell Electrocatalysts. *Surf. Sci. Rep.* **2002**, *45* (4–6), 117–229. [https://doi.org/10.1016/S0167-5729\(01\)00022-X](https://doi.org/10.1016/S0167-5729(01)00022-X).
 - (44) Somorjai, G. A. Modern Surface Science and Surface Technologies: An Introduction. *Chem. Rev.* **1996**, *96* (4), 1223–1236. <https://doi.org/10.1021/cr950234e>.
 - (45) Zaera, F. The Surface Chemistry of Metal-Based Hydrogenation Catalysis. *ACS Catal.* **2017**, *7* (8), 4947–4967. <https://doi.org/10.1021/acscatal.7b01368>.
 - (46) Kiselev, V. F.; Krylov, O. V. *Adsorption and Catalysis on Transition Metals and Their Oxides*; Springer Series in Surface Sciences; Springer Berlin Heidelberg, **2012**.
 - (47) Kittel, C.; McEuen, P. *Introduction to Solid State Physics*; Wiley, **2018**.
 - (48) Britannica, T. E. of E. The Fermi-Level. *Encyclopedia Britannica*; **2006**.
 - (49) Cantor, B. The Fermi Level. In *The Equations of Materials*; Oxford University Press, **2020**; pp 267–300. <https://doi.org/10.1093/oso/9780198851875.003.0013>.
 - (50) Somorjai, G. A.; York, R. L.; Butcher, D.; Park, J. Y. The Evolution of Model Catalytic Systems; Studies of Structure, Bonding and Dynamics from Single Crystal Metal Surfaces to Nanoparticles, and from Low Pressure (<10^{−3} Torr) to High Pressure (>10^{−3} Torr) to Liquid Interfaces. *Phys. Chem. Chem. Phys.* **2007**, *9* (27), 3500–3513. <https://doi.org/10.1039/B618805B>.
 - (51) Auras, S. V.; van Bree, R. A. B.; Bashlakov, D. L.; van Lent, R.; Juurlink, L. B. F. It's Not Just the Defects – a Curved Crystal Study of H₂O Desorption from Ag. *Phys. Chem. Chem. Phys.* **2019**, *21* (28), 15422–15430. <https://doi.org/10.1039/C9CP02609F>.
 - (52) Ashcroft, N. W.; Mermin, N. D. *Solid State Physics*; Cengage Learning, **2011**.
 - (53) Horuiti, J.; Polanyi, M. Catalytic Interchange of Hydrogen between Water and Ethylene and between Water and Benzene. *Nature* **1934**, *134* (3384), 377–378. <https://doi.org/10.1038/134377b0>.
 - (54) Horiuti, I.; Polanyi, M. Exchange Reactions of Hydrogen on Metallic Catalysts. *Trans. Faraday Soc.* **1934**, *30*, 1164. <https://doi.org/10.1039/tf9343001164>.
 - (55) Kroes, G.-J.; Díaz, C. Quantum and Classical Dynamics of Reactive Scattering of H₂ from Metal Surfaces. *Chem. Soc. Rev.* **2016**, *45* (13), 3658–3700. <https://doi.org/10.1039/C5CS00336A>.
 - (56) Kroes, G.-J. Computational Approaches to Dissociative Chemisorption on Metals: Towards Chemical Accuracy. *Phys. Chem. Chem. Phys.* **2021**, *23* (15), 8962–9048. <https://doi.org/10.1039/D1CP00044F>.
 - (57) Muller, P. Glossary of Terms Used in Physical Organic Chemistry (IUPAC Recommendations 1994). *Pure Appl. Chem.* **1994**, *66* (5), 1077–1184. <https://doi.org/10.1351/pac199466051077>.
 - (58) Jackson, J. M.; Howarth, A. Exchange of Energy between Diatomic Gas Molecules and a Solid Surface. *Proc. R. Soc. London. Ser. A - Math. Phys. Sci.* **1935**, *152* (877), 515–529. <https://doi.org/10.1098/rspa.1935.0206>.
 - (59) Shenvi, N.; Roy, S.; Tully, J. C. Dynamical Steering and Electronic Excitation in NO Scattering from a Gold Surface. *Science (80-.)*. **2009**, *326* (5954), 829–832. <https://doi.org/10.1126/science.1179240>.
 - (60) Bunermann, O.; Jiang, H.; Dorenkamp, Y.; Kandratsenka, A.; Janke, S. M.; Auerbach, D. J.; Wodtke, A. M. Electron-Hole Pair Excitation Determines the Mechanism of Hydrogen

- Atom Adsorption. *Science* (80-.). **2015**, 350 (6266), 1346–1349.
<https://doi.org/10.1126/science.aad4972>.
- (61) Blanco-Rey, M.; Juaristi, J. I.; Díez Muiño, R.; Busnengo, H. F.; Kroes, G. J.; Alducin, M. Electronic Friction Dominates Hydrogen Hot-Atom Relaxation on Pd(100). *Phys. Rev. Lett.* **2014**, 112 (10), 1–5. <https://doi.org/10.1103/PhysRevLett.112.103203>.
- (62) Estermann, I.; Stern, O. Beugung von Molekularstrahlen. *Zeitschrift fur Phys.* **1930**, 61 (1–2), 95–125. <https://doi.org/10.1007/BF01340293>.
- (63) Singh, R. C. V. Raman and the Discovery of the Raman Effect. *Phys. Perspect.* **2002**, 4 (4), 399–420. <https://doi.org/10.1007/s000160200002>.
- (64) Heisenberg, W. Ober Den Anschaulichen Inhalt Der Quantentheoretischen Kinematik Und Mechanik. *Zeitschrift fur Phys. Phys.* **1927**, 43 (3–4), 172–198.
<https://doi.org/10.1007/BF01397280>.
- (65) Cao, K.; van Lent, R.; Kleyn, A. W.; Kurahashi, M.; Juurlink, L. B. F. Steps on Pt Stereodynamically Filter Sticking of O₂. *Proc. Natl. Acad. Sci. U. S. A.* **2019**, 116 (28), 13862–13866. <https://doi.org/10.1073/pnas.1902846116>.
- (66) Kroes, G.-J. Toward a Database of Chemically Accurate Barrier Heights for Reactions of Molecules with Metal Surfaces. *J. Phys. Chem. Lett.* **2015**, 6 (20), 4106–4114.
<https://doi.org/10.1021/acs.jpclett.5b01344>.
- (67) Michelsen, H. A.; Auerbach, D. J. A Critical Examination of Data on the Dissociative Adsorption and Associative Desorption of Hydrogen at Copper Surfaces. *J. Chem. Phys.* **1991**, 94 (11), 7502–7520. <https://doi.org/10.1063/1.460182>.
- (68) Michelsen, H. A.; Rettner, C. T.; Auerbach, D. J.; Zare, R. N. Effect of Rotation on the Translational and Vibrational Energy Dependence of the Dissociative Adsorption of D₂ on Cu(111). *J. Chem. Phys.* **1993**, 98 (10), 8294–8307.
<https://doi.org/10.1063/1.464535>.
- (69) King, D. A.; Wells, M. G. Molecular Beam Investigation of Adsorption Kinetics on Bulk Metal Targets: Nitrogen on Tungsten. *Surf. Sci.* **1972**, 29 (2), 454–482.
[https://doi.org/10.1016/0039-6028\(72\)90232-4](https://doi.org/10.1016/0039-6028(72)90232-4).
- (70) Kleyn, A. W. Molecular Beams and Chemical Dynamics at Surfaces. *Chem. Soc. Rev.* **2003**, 32 (2), 87–95. <https://doi.org/10.1039/b105760j>.
- (71) Luntz, A. C.; Brown, J. K.; Williams, M. D. Molecular Beam Studies of H₂ and D₂ Dissociative Chemisorption on Pt(111). *J. Chem. Phys.* **1990**, 93 (7), 5240–5246.
<https://doi.org/10.1063/1.459669>.
- (72) Balooch, M.; Cardillo, M. J.; Miller, D. R.; Stickney, R. E. Molecular Beam Study of the Apparent Activation Barrier Associated with Adsorption and Desorption of Hydrogen on Copper. *Surf. Sci.* **1974**, 46 (2), 358–392. [https://doi.org/10.1016/0039-6028\(74\)90315-X](https://doi.org/10.1016/0039-6028(74)90315-X).
- (73) Kurahashi, M.; Yamauchi, Y. Production of a Single Spin-Rotational State [(J,M)=(2,2)] Selected Molecular Oxygen (^Σg⁻) Beam by a Hexapole Magnet. *Rev. Sci. Instrum.* **2009**, 80 (8), 083103. <https://doi.org/10.1063/1.3206299>.
- (74) Yoder, B. L.; Bisson, R.; Beck, R. D. Steric Effects in the Chemisorption of Vibrationally Excited Methane on Ni(100). *Science* (80-.). **2010**, 329 (5991), 553–556.
<https://doi.org/10.1126/science.1191751>.
- (75) Vattuone, L.; Savio, L.; Pirani, F.; Cappelletti, D.; Okada, M.; Rocca, M. Interaction of Rotationally Aligned and of Oriented Molecules in Gas Phase and at Surfaces. *Prog. Surf. Sci.* **2010**, 85 (1–4), 92–160. <https://doi.org/10.1016/j.progsurf.2009.12.001>.
- (76) Chadwick, H.; Somers, M. F.; Stewart, A. C.; Alkoby, Y.; Carter, T. J. D.; Butkovicova, D.; Alexandrowicz, G. Stopping Molecular Rotation Using Coherent Ultra-Low-Energy Magnetic Manipulations. *Nat. Commun.* **2022**, 13 (1), 2287.
<https://doi.org/10.1038/s41467-022-29830-3>.

- (77) Hanbicki, A. T.; Baddorf, A. P.; Plummer, E. W.; Hammer, B.; Scheffler, M. The Interaction of Hydrogen with the (110) Surface of NiAl. *Surf. Sci.* **1995**, *331–333*, 811–817. [https://doi.org/10.1016/0039-6028\(95\)00376-2](https://doi.org/10.1016/0039-6028(95)00376-2).
- (78) Smith, R. S.; Huang, C.; Wong, E. K. L.; Kay, B. D. The Molecular Volcano: Abrupt CCl₄ Desorption Driven by the Crystallization of Amorphous Solid Water. *Phys. Rev. Lett.* **1997**, *79* (5), 909–912. <https://doi.org/10.1103/PhysRevLett.79.909>.
- (79) Henzler, M. LEED-Investigation of Step Arrays on Cleaved Germanium (111) Surfaces. *Surf. Sci.* **1970**, *19* (1), 159–171. [https://doi.org/10.1016/0039-6028\(70\)90115-9](https://doi.org/10.1016/0039-6028(70)90115-9).
- (80) Van Hove, M. A.; Somorjai, G. A. A New Microfacet Notation for High-Miller-Index Surfaces of Cubic Materials with Terrace, Step and Kink Structures. *Surf. Sci.* **1980**, *92* (2–3), 489–518. [https://doi.org/10.1016/0039-6028\(80\)90219-8](https://doi.org/10.1016/0039-6028(80)90219-8).
- (81) Chang, C. C. Auger Electron Spectroscopy. *Surf. Sci.* **1971**, *25* (1), 53–79. [https://doi.org/10.1016/0039-6028\(71\)90210-X](https://doi.org/10.1016/0039-6028(71)90210-X).
- (82) Mathieu, H. J. *Surface Analysis – The Principal Techniques*; Vickerman, J. C., Gilmore, I. S., Eds.; Wiley, **2009**. <https://doi.org/10.1002/9780470721582>.
- (83) Wagner, C. D.; Zatko, D. A.; Raymond, R. H. Use of the Oxygen KLL Auger Lines in Identification of Surface Chemical States by Electron Spectroscopy for Chemical Analysis. *Anal. Chem.* **1980**, *52* (9), 1445–1451. <https://doi.org/10.1021/ac50059a017>.
- (84) Mathieu, H. J.; Landolt, D. Quantitative Auger Electron Spectroscopy Analysis of Ag-Pd and Ni-Pd Alloys. *Surf. Sci.* **1975**, *53* (1), 228–240. [https://doi.org/10.1016/0039-6028\(75\)90126-0](https://doi.org/10.1016/0039-6028(75)90126-0).
- (85) Hall, P. M.; Morabito, J. M. Matrix Effects in Quantitative Auger Analysis of Dilute Alloys. *Surf. Sci.* **1979**, *83* (2), 391–405. [https://doi.org/10.1016/0039-6028\(79\)90052-9](https://doi.org/10.1016/0039-6028(79)90052-9).
- (86) Hall, P. M.; Morabito, J. M.; Conley, D. K. Relative Sensitivity Factors for Quantitative Auger Analysis of Binary Alloys. *Surf. Sci.* **1977**, *62* (1), 1–20. [https://doi.org/10.1016/0039-6028\(77\)90424-1](https://doi.org/10.1016/0039-6028(77)90424-1).
- (87) Langeron, J. P.; Minel, L.; Vignes, J. L.; Bouquet, S.; Pellerin, F.; Lorang, G.; Ailloud, P.; Le Héricy, J. An Important Step in Quantitative Auger Analysis: The Use of Peak to Background Ratio. *Surf. Sci.* **1984**, *138* (2–3), 610–628. [https://doi.org/10.1016/0039-6028\(84\)90269-3](https://doi.org/10.1016/0039-6028(84)90269-3).
- (88) Shimizu, R. Quantitative Analysis by Auger Electron Spectroscopy. *Jpn. J. Appl. Phys.* **1983**, *22* (11R), 1631. <https://doi.org/10.1143/JJAP.22.1631>.
- (89) Liday, J.; Hofmann, S.; Harman, R. Quantitative Auger Electron Analysis of Homogeneous Binary Alloys of Cr, Fe and Ni. *Vacuum* **1992**, *43* (4), 331–339. [https://doi.org/10.1016/0042-207X\(92\)90166-T](https://doi.org/10.1016/0042-207X(92)90166-T).
- (90) Siegbahn, K.; Edvarson, K. β -Ray Spectroscopy in the Precision Range of 1 : 105. *Nucl. Phys.* **1956**, *1* (8), 137–159. [https://doi.org/10.1016/S0029-5582\(56\)80022-9](https://doi.org/10.1016/S0029-5582(56)80022-9).
- (91) Andrade, J. D. X-Ray Photoelectron Spectroscopy (XPS). In *Surface and Interfacial Aspects of Biomedical Polymers*; Springer US: Boston, MA, **1985**; pp 105–195. https://doi.org/10.1007/978-1-4684-8610-0_5.
- (92) Powell, C. J.; Seah, M. P. Precision, Accuracy, and Uncertainty in Quantitative Surface Analyses by Auger-Electron Spectroscopy and x-Ray Photoelectron Spectroscopy. *J. Vac. Sci. Technol. A Vacuum, Surfaces, Film.* **1990**, *8* (2), 735–763. <https://doi.org/10.1116/1.576956>.
- (93) Fadley, C. S. X-Ray Photoelectron Spectroscopy: Progress and Perspectives. *J. Electron Spectros. Relat. Phenomena* **2010**, *178–179* (C), 2–32. <https://doi.org/10.1016/j.elspec.2010.01.006>.
- (94) Greczynski, G.; Hultman, L. X-Ray Photoelectron Spectroscopy: Towards Reliable Binding Energy Referencing. *Prog. Mater. Sci.* **2020**, *107*, 100591. <https://doi.org/10.1016/j.pmatsci.2019.100591>.

- (95) Schatz, G. C.; Wodtke, A. M.; Yang, X. Spiers Memorial Lecture: New Directions in Molecular Scattering. *Faraday Discuss.* **2024**, *251*, 9–62. <https://doi.org/10.1039/D4FD00015C>.
- (96) Laidler, K. J. A Glossary of Terms Used in Chemical Kinetics, Including Reaction Dynamics (IUPAC Recommendations 1996). *Pure Appl. Chem.* **1996**, *68* (1), 149–192. <https://doi.org/10.1351/pac199668010149>.
- (97) Binnig, G.; Rohrer, H.; Gerber, C.; Weibel, E. Surface Studies by Scanning Tunneling Microscopy. *Phys. Rev. Lett.* **1982**, *49* (1), 57–61. <https://doi.org/10.1103/PhysRevLett.49.57>.
- (98) Binnig, G.; Rohrer, H. Scanning Tunneling Microscopy. *Surf. Sci.* **1983**, *126* (1–3), 236–244. [https://doi.org/10.1016/0039-6028\(83\)90716-1](https://doi.org/10.1016/0039-6028(83)90716-1).
- (99) Chen, C. J. Scanning Tunneling Microscopy: A Chemical Perspective. *Scanning Microsc.* **1993**, *7* (3), 793–804.
- (100) Chen, C. J. *Introduction to Scanning Tunneling Microscopy*; Oxford University Press/Oxford, **2007**. <https://doi.org/10.1093/acprof:oso/9780199211500.001.0001>.
- (101) Laegsgaard, E.; Österlund, L.; Thostrup, P.; Rasmussen, P. B.; Stensgaard, I.; Besenbacher, F. A High-Pressure Scanning Tunneling Microscope. *Rev. Sci. Instrum.* **2001**, *72* (9), 3537–3542. <https://doi.org/10.1063/1.1389497>.
- (102) Rößler, M.; Geng, P.; Wintterlin, J. A High-Pressure Scanning Tunneling Microscope for Studying Heterogeneous Catalysis. *Rev. Sci. Instrum.* **2005**, *76* (2). <https://doi.org/10.1063/1.1841951>.
- (103) Tao, F.; Tang, D.; Salmeron, M.; Somorjai, G. A. A New Scanning Tunneling Microscope Reactor Used for High-Pressure and High-Temperature Catalysis Studies. *Rev. Sci. Instrum.* **2008**, *79* (8). <https://doi.org/10.1063/1.2960569>.
- (104) Herbschleb, C. T.; van der Tuijn, P. C.; Roobol, S. B.; Navarro, V.; Bakker, J. W.; Liu, Q.; Stoltz, D.; Cañas-Ventura, M. E.; Verdoes, G.; van Spronsen, M. A.; Bergman, M.; Crama, L.; Taminiau, I.; Ofitserov, A.; van Baarle, G. J. C.; Frenken, J. W. M. The ReactorSTM: Atomically Resolved Scanning Tunneling Microscopy under High-Pressure, High-Temperature Catalytic Reaction Conditions. *Rev. Sci. Instrum.* **2014**, *85* (8). <https://doi.org/10.1063/1.4891811>.
- (105) Prabhu, M. K.; Boden, D.; Rost, M. J.; Meyer, J.; Groot, I. M. N. Structural Characterization of a Novel Two-Dimensional Material: Cobalt Sulfide Sheets on Au(111). *J. Phys. Chem. Lett.* **2020**, *11* (21), 9038–9044. <https://doi.org/10.1021/acs.jpclett.0c02268>.
- (106) Wenzel, S.; Groot, I. M. N. ZnO (1010) Is Unstable in Moderate Pressures of Water. *Surf. Sci.* **2022**, *715*, 121940. <https://doi.org/10.1016/j.susc.2021.121940>.
- (107) Wenzel, S.; Boden, D.; van Lent, R.; Motaei, E.; Prabhu, M. K.; Achour, H.; Groot, I. M. N. Spectroscopic Investigation of a Co(0001) Model Catalyst during Exposure to H₂ and CO at near-Ambient Pressures. *Phys. Chem. Chem. Phys.* **2023**, *25* (37), 25094–25104. <https://doi.org/10.1039/D3CP02739B>.
- (108) Chadwick, H.; Alexandrowicz, G. Temperature Dependent Stereodynamics in Surface Scattering Measured through Subtle Changes in the Molecular Wave Function. *Faraday Discuss.* **2024**, *251*, 76–91. <https://doi.org/10.1039/D4FD00007B>.
- (109) Smits, B.; Somers, M. F. The Quantum Dynamics of H₂ on Cu(111) at a Surface Temperature of 925 K: Comparing State-of-the-Art Theory to State-of-the-Art Experiments 2. *J. Chem. Phys.* **2023**, *158* (1), 014704. <https://doi.org/10.1063/5.0134817>.
- (110) Smits, B.; Somers, M. F. The Quantum Dynamics of H₂ on Cu(111) at a Surface Temperature of 925 K: Comparing State-of-the-Art Theory to State-of-the-Art Experiments. *J. Chem. Phys.* **2022**, *157* (13), 134704.

- <https://doi.org/10.1063/5.0112036>.
- (111) Wodtke, A. M. Chemistry in a Computer: Advancing the in Silico Dream. *Science* (80-.). **2006**, 312 (5770), 64–65. <https://doi.org/10.1126/science.1124924>.
 - (112) Cramer, C. J. *Essentials of Computational Chemistry Theories and Models*, 2nd ed.; John Wiley & Sons, Ltd, 2017.
 - (113) Doblhoff-Dier, K.; Meyer, J.; Hoggan, P. E.; Kroes, G.-J.; Wagner, L. K. Diffusion Monte Carlo for Accurate Dissociation Energies of 3d Transition Metal Containing Molecules. *J. Chem. Theory Comput.* **2016**, 12 (6), 2583–2597. <https://doi.org/10.1021/acs.jctc.6b00160>.
 - (114) Doblhoff-Dier, K.; Meyer, J.; Hoggan, P. E.; Kroes, G.-J. Quantum Monte Carlo Calculations on a Benchmark Molecule–Metal Surface Reaction: H 2 + Cu(111). *J. Chem. Theory Comput.* **2017**, 13 (7), 3208–3219. <https://doi.org/10.1021/acs.jctc.7b00344>.
 - (115) Powell, A. D.; Gerrits, N.; Tchakoua, T.; Somers, M. F.; Busnengo, H. F.; Meyer, J.; Kroes, G.; Doblhoff-Dier, K. Best-of-Both-Worlds Predictive Approach to Dissociative Chemisorption on Metals. *J. Phys. Chem. Lett.* **2024**, 307–315. <https://doi.org/10.1021/acs.jpcllett.3c02972>.
 - (116) Zhao, Q.; Zhang, X.; Martinez, J. M. P.; Carter, E. A. Benchmarking an Embedded Adaptive Sampling Configuration Interaction Method for Surface Reactions: H2Desorption from and CH4Dissociation on Cu(111). *J. Chem. Theory Comput.* **2020**, 16 (11), 7078–7088. <https://doi.org/10.1021/acs.jctc.0c00341>.
 - (117) Born, M.; Oppenheimer, R. Zur Quantentheorie Der Molekeln. *Ann. Phys.* **1927**, 389 (20), 457–484. <https://doi.org/10.1002/andp.19273892002>.
 - (118) Filot, I. A. W. *Elements of Electronic Structure Theory*, 1st ed.; **2022**.
 - (119) Sherrill, C. D. The Born-Oppenheimer Approximation. *Sch. Chem. Biochem. Georg. Inst. Technol.* **2005**, 242.
 - (120) Dewar, M. J. S.; Kelemen, J. LCAO MO Theory Illustrated by Its Application to H2. *J. Chem. Educ.* **1971**, 48 (8), 494. <https://doi.org/10.1021/ed048p494>.
 - (121) Hohenberg, P.; Kohn, W. Inhomogeneous Electron Gas. *Phys. Rev.* **1964**, 136 (3B), B864–B871. <https://doi.org/10.1103/PhysRev.136.B864>.
 - (122) Kohn, W.; Sham, L. J. Self-Consistent Equations Including Exchange and Correlation Effects. *Phys. Rev.* **1965**, 140 (4A), A1133–A1138. <https://doi.org/10.1103/PhysRev.140.A1133>.
 - (123) Gerrits, N.; Shakouri, K.; Behler, J.; Kroes, G.-J. Accurate Probabilities for Highly Activated Reaction of Polyatomic Molecules on Surfaces Using a High-Dimensional Neural Network Potential: CHD 3 + Cu(111). *J. Phys. Chem. Lett.* **2019**, 10 (8), 1763–1768. <https://doi.org/10.1021/acs.jpcllett.9b00560>.
 - (124) Gerrits, N.; Smeets, E. W. F.; Vuckovic, S.; Powell, A. D.; Doblhoff-Dier, K.; Kroes, G.-J. Density Functional Theory for Molecule–Metal Surface Reactions: When Does the Generalized Gradient Approximation Get It Right, and What to Do If It Does Not. *J. Phys. Chem. Lett.* **2020**, 11 (24), 10552–10560. <https://doi.org/10.1021/acs.jpcllett.0c02452>.
 - (125) Gerrits, N.; Geweke, J.; Smeets, E. W. F.; Voss, J.; Wodtke, A. M.; Kroes, G.-J. Closing the Gap Between Experiment and Theory: Reactive Scattering of HCl from Au(111). *J. Phys. Chem. C* **2020**, 124 (29), 15944–15960. <https://doi.org/10.1021/acs.jpcc.0c03756>.
 - (126) Smeets, E. W. F.; Kroes, G.-J. Designing New SRP Density Functionals Including Non-Local VdW-DF2 Correlation for H 2 + Cu(111) and Their Transferability to H 2 + Ag(111), Au(111) and Pt(111). *Phys. Chem. Chem. Phys.* **2021**, 23 (13), 7875–7901. <https://doi.org/10.1039/D0CP05173J>.
 - (127) Tchakoua, T.; Smeets, E. W. F.; Somers, M.; Kroes, G.-J. Toward a Specific Reaction Parameter Density Functional for H 2 + Ni(111): Comparison of Theory with Molecular Beam Sticking Experiments. *J. Phys. Chem. C* **2019**, 123 (33), 20420–20433.

- <https://doi.org/10.1021/acs.jpcc.9b05928>.
- (128) Smeets, E. W. F.; Voss, J.; Kroes, G.-J. Specific Reaction Parameter Density Functional Based on the Meta-Generalized Gradient Approximation: Application to H₂ + Cu(111) and H₂ + Ag(111). *J. Phys. Chem. A* **2019**, *123* (25), 5395–5406. <https://doi.org/10.1021/acs.jpca.9b02914>.
 - (129) Tchakoua, T.; Gerrits, N.; Smeets, E. W. F.; Kroes, G.-J. SBH17: Benchmark Database of Barrier Heights for Dissociative Chemisorption on Transition Metal Surfaces. *J. Chem. Theory Comput.* **2023**, *19* (1), 245–270. <https://doi.org/10.1021/acs.jctc.2c00824>.
 - (130) Smeets, E. W. F.; Kroes, G. J. Performance of Made Simple Meta-GGA Functionals with RVV10 Nonlocal Correlation for H₂ + Cu(111), D₂ + Ag(111), H₂ + Au(111), and D₂ + Pt(111). *J. Phys. Chem. C* **2021**, *125* (17), 8993–9010. <https://doi.org/10.1021/acs.jpcc.0c11034>.
 - (131) Nattino, F.; Costanzo, F.; Kroes, G.-J. N₂ Dissociation on W(110): An Ab Initio Molecular Dynamics Study on the Effect of Phonons. *J. Chem. Phys.* **2015**, *142* (10). <https://doi.org/10.1063/1.4913979>.
 - (132) Migliorini, D.; Nattino, F.; Kroes, G.-J. Application of van Der Waals Functionals to the Calculation of Dissociative Adsorption of N₂ on W(110) for Static and Dynamic Systems. *J. Chem. Phys.* **2016**, *144* (8), 084702. <https://doi.org/10.1063/1.4942198>.
 - (133) Nattino, F.; Migliorini, D.; Kroes, G.-J.; Dombrowski, E.; High, E. A.; Killelea, D. R.; Utz, A. L. Chemically Accurate Simulation of a Polyatomic Molecule-Metal Surface Reaction. *J. Phys. Chem. Lett.* **2016**, *7* (13), 2402–2406. <https://doi.org/10.1021/acs.jpcllett.6b01022>.
 - (134) Migliorini, D.; Chadwick, H.; Nattino, F.; Gutiérrez-González, A.; Dombrowski, E.; High, E. A.; Guo, H.; Utz, A. L.; Jackson, B.; Beck, R. D.; Kroes, G. J. Surface Reaction Barriometry: Methane Dissociation on Flat and Stepped Transition-Metal Surfaces. *J. Phys. Chem. Lett.* **2017**, *8* (17), 4177–4182. <https://doi.org/10.1021/acs.jpcllett.7b01905>.
 - (135) Díaz, C.; Pijper, E.; Olsen, R. A.; Busnengo, H. F.; Auerbach, D. J.; Kroes, G. J. Chemically Accurate Simulation of a Prototypical Surface Reaction: H₂ Dissociation on Cu(111). *Science* (80-.). **2009**, *326* (5954), 832–834. <https://doi.org/10.1126/science.1178722>.
 - (136) Füchsel, G.; Del Cueto, M.; Díaz, C.; Kroes, G. J. Enigmatic HCl + Au(111) Reaction: A Puzzle for Theory and Experiment. *J. Phys. Chem. C* **2016**, *120* (45), 25760–25779. <https://doi.org/10.1021/acs.jpcc.6b07453>.
 - (137) Peverati, R.; Truhlar, D. G. Quest for a Universal Density Functional: The Accuracy of Density Functionals across a Broad Spectrum of Databases in Chemistry and Physics. *Philos. Trans. R. Soc. A Math. Phys. Eng. Sci.* **2014**, *372* (2011), 20120476. <https://doi.org/10.1098/rsta.2012.0476>.
 - (138) Goerigk, L.; Kruse, H.; Grimme, S. Benchmarking Density Functional Methods against the S66 and S66x8 Datasets for Non-Covalent Interactions. *ChemPhysChem* **2011**, *12* (17), 3421–3433. <https://doi.org/10.1002/cphc.201100826>.
 - (139) Mardirossian, N.; Head-Gordon, M. Thirty Years of Density Functional Theory in Computational Chemistry: An Overview and Extensive Assessment of 200 Density Functionals. *Mol. Phys.* **2017**, *115* (19), 2315–2372. <https://doi.org/10.1080/00268976.2017.1333644>.
 - (140) Spiegel, M.; Gamian, A.; Sroka, Z. A Statistically Supported Antioxidant Activity DFT Benchmark—The Effects of Hartree–Fock Exchange and Basis Set Selection on Accuracy and Resources Uptake. *Molecules* **2021**, *26* (16), 5058. <https://doi.org/10.3390/molecules26165058>.
 - (141) Marcel Swart. *Average results for the 2010–2023 editions of the Annual DFT Popularity Poll*. <https://www.marcelswart.eu/dft-poll/results.html> (accessed 2024-08-28).
 - (142) Nour Ghassemi, E.; Wijzenbroek, M.; Somers, M. F.; Kroes, G.-J. Chemically Accurate Simulation of Dissociative Chemisorption of D₂ on Pt(1 1 1). *Chem. Phys. Lett.* **2017**,

- 683, 329–335. <https://doi.org/10.1016/j.cplett.2016.12.059>.
- (143) Ghassemi, E. N.; Somers, M. F.; Kroes, G.-J. Assessment of Two Problems of Specific Reaction Parameter Density Functional Theory: Sticking and Diffraction of H₂ on Pt(111). *J. Phys. Chem. C* **2019**, *123* (16), 10406–10418. <https://doi.org/10.1021/acs.jpcc.9b00981>.
- (144) Boereboom, J. M.; Wijzenbroek, M.; Somers, M. F.; Kroes, G. J. Towards a Specific Reaction Parameter Density Functional for Reactive Scattering of H₂ from Pd(111). *J. Chem. Phys.* **2013**, *139* (24). <https://doi.org/10.1063/1.4851355>.
- (145) Gerrits, N.; Smeets, E. W. F.; Vuckovic, S.; Powell, A. D.; Doblhoff-Dier, K.; Kroes, G.-J. Correction to “Density Functional Theory for Molecule–Metal Surface Reactions: When Does the Generalized Gradient Approximation Get It Right, and What to Do If It Does Not.” *J. Phys. Chem. Lett.* **2022**, *13* (45), 10575–10576. <https://doi.org/10.1021/acs.jpcllett.2c03223>.
- (146) Cohen, A. J.; Mori-Sánchez, P.; Yang, W. Insights into Current Limitations of Density Functional Theory. *Science* (80-.). **2008**, *321* (5890), 792–794. <https://doi.org/10.1126/science.1158722>.
- (147) Kim, M.-C.; Sim, E.; Burke, K. Understanding and Reducing Errors in Density Functional Calculations. *Phys. Rev. Lett.* **2013**, *111* (7), 073003. <https://doi.org/10.1103/PhysRevLett.111.073003>.
- (148) Song, S.; Vuckovic, S.; Sim, E.; Burke, K. Density Sensitivity of Empirical Functionals. *J. Phys. Chem. Lett.* **2021**, *12* (2), 800–807. <https://doi.org/10.1021/acs.jpcllett.0c03545>.
- (149) Vuckovic, S.; Song, S.; Kozłowski, J.; Sim, E.; Burke, K. Density Functional Analysis: The Theory of Density-Corrected DFT. *J. Chem. Theory Comput.* **2019**, *15* (12), 6636–6646. <https://doi.org/10.1021/acs.jctc.9b00826>.
- (150) Kaplan, A. D.; Shahi, C.; Bhetwal, P.; Sah, R. K.; Perdew, J. P. Understanding Density-Driven Errors for Reaction Barrier Heights. *J. Chem. Theory Comput.* **2023**, *19* (2), 532–543. <https://doi.org/10.1021/acs.jctc.2c00953>.
- (151) Kanungo, B.; Kaplan, A. D.; Shahi, C.; Gavini, V.; Perdew, J. P. Unconventional Error Cancellation Explains the Success of Hartree–Fock Density Functional Theory for Barrier Heights. *J. Phys. Chem. Lett.* **2024**, *15* (1), 323–328. <https://doi.org/10.1021/acs.jpcllett.3c03088>.
- (152) Perdew, J. P.; Ernzerhof, M.; Burke, K. Rationale for Mixing Exact Exchange with Density Functional Approximations. *J. Chem. Phys.* **1996**, *105* (22), 9982–9985. <https://doi.org/10.1063/1.472933>.
- (153) Adamo, C.; Barone, V. Toward Reliable Density Functional Methods without Adjustable Parameters: The PBE0 Model. *J. Chem. Phys.* **1999**, *110* (13), 6158–6170. <https://doi.org/10.1063/1.478522>.
- (154) Perdew, J. P.; Burke, K.; Ernzerhof, M. Generalized Gradient Approximation Made Simple. *Phys. Rev. Lett.* **1996**, *77* (18), 3865–3868. <https://doi.org/10.1103/PhysRevLett.77.3865>.
- (155) Henderson, T. M.; Izmaylov, A. F.; Scuseria, G. E.; Savin, A. The Importance of Middle-Range Hartree-Fock-Type Exchange for Hybrid Density Functionals. *J. Chem. Phys.* **2007**, *127* (22), 221103. <https://doi.org/10.1063/1.2822021>.
- (156) Krukau, A. V.; Vydrov, O. A.; Izmaylov, A. F.; Scuseria, G. E. Influence of the Exchange Screening Parameter on the Performance of Screened Hybrid Functionals. *J. Chem. Phys.* **2006**, *125* (22), 224106. <https://doi.org/10.1063/1.2404663>.
- (157) Brütting, M.; Bahmann, H.; Kümmel, S. Hybrid Functionals with Local Range Separation: Accurate Atomization Energies and Reaction Barrier Heights. *J. Chem. Phys.* **2022**, *156* (10), 104109. <https://doi.org/10.1063/5.0082957>.
- (158) Gao, W.; Abtew, T. A.; Cai, T.; Sun, Y. Y.; Zhang, S.; Zhang, P. On the Applicability of

- Hybrid Functionals for Predicting Fundamental Properties of Metals. *Solid State Commun.* **2016**, 234–235, 10–13. <https://doi.org/10.1016/j.ssc.2016.02.014>.
- (159) Smits, B.; Somers, M. F. Beyond the Static Corrugation Model: Dynamic Surfaces with the Embedded Atom Method. *J. Chem. Phys.* **2021**, 154 (7), 074710. <https://doi.org/10.1063/5.0036611>.
- (160) Smits, B.; Litjens, L. G. B.; Somers, M. F. Accurate Description of the Quantum Dynamical Surface Temperature Effects on the Dissociative Chemisorption of H₂ from Cu(111). *J. Chem. Phys.* **2022**, 156 (21), 214706. <https://doi.org/10.1063/5.0094985>.
- (161) Kroes, G.; Meyer, J. Best-of-Both-Worlds Computational Approaches to Difficult-to-Model Dissociation Reactions on Metal Surfaces. *Chem. Sci.* **2025**, 16 (2), 480–506. <https://doi.org/10.1039/D4SC06004K>.
- (162) Goikoetxea, I.; Beltrán, J.; Meyer, J.; Iñaki Juaristi, J.; Alducin, M.; Reuter, K. Non-Adiabatic Effects during the Dissociative Adsorption of O₂ at Ag(111)? A First-Principles Divide and Conquer Study. *New J. Phys.* **2012**, 14 (1), 013050. <https://doi.org/10.1088/1367-2630/14/1/013050>.
- (163) Daru, J.; Forbert, H.; Behler, J.; Marx, D. Coupled Cluster Molecular Dynamics of Condensed Phase Systems Enabled by Machine Learning Potentials: Liquid Water Benchmark. *Phys. Rev. Lett.* **2022**, 129 (22), 226001. <https://doi.org/10.1103/PhysRevLett.129.226001>.
- (164) Jensen, F. *Introduction to Computational Chemistry*, 2nd ed.; John Wiley & Sons, Ltd, 2007.
- (165) Schrödinger, E. An Undulatory Theory of the Mechanics of Atoms and Molecules. *Phys. Rev.* **1926**, 28 (6), 1049–1070. <https://doi.org/10.1103/PhysRev.28.1049>.
- (166) Born, M. Zur Quantenmechanik Der Stossvorgänge. *Zeitschrift für Phys.* **1926**, 37 (12), 863–867. <https://doi.org/10.1007/BF01397477>.
- (167) Dirac, P. A. M. A New Notation for Quantum Mechanics. *Math. Proc. Cambridge Philos. Soc.* **1939**, 35 (3), 416–418. <https://doi.org/10.1017/S0305004100021162>.
- (168) Hartree, D. R. The Wave Mechanics of an Atom with a Non-Coulomb Central Field. Part I. Theory and Methods. *Math. Proc. Cambridge Philos. Soc.* **1928**, 24 (1), 89–110. <https://doi.org/10.1017/S0305004100011919>.
- (169) Nienhaus, H.; Bergh, H. S.; Gergen, B.; Majumdar, A.; Weinberg, W. H.; McFarland, E. W. Direct Detection of Electron-Hole Pairs Generated by Chemical Reactions on Metal Surfaces. *Surf. Sci.* **2000**, 445 (2–3), 335–342. [https://doi.org/10.1016/S0039-6028\(99\)01078-X](https://doi.org/10.1016/S0039-6028(99)01078-X).
- (170) Spiering, P.; Meyer, J. Testing Electronic Friction Models: Vibrational De-Excitation in Scattering of H₂ and D₂ from Cu(111). *J. Phys. Chem. Lett.* **2018**, 9 (7), 1803–1808. <https://doi.org/10.1021/acs.jpclett.7b03182>.
- (171) Spiering, P.; Shakouri, K.; Behler, J.; Kroes, G. J.; Meyer, J. Orbital-Dependent Electronic Friction Significantly Affects the Description of Reactive Scattering of N₂ from Ru(0001). *J. Phys. Chem. Lett.* **2019**, 10 (11), 2957–2962. <https://doi.org/10.1021/acs.jpclett.9b00523>.
- (172) Pauli, W. Über Den Zusammenhang Des Abschlusses Der Elektronengruppen Im Atom Mit Der Komplexstruktur Der Spektren. *Zeitschrift für Phys.* **1925**, 31 (1), 765–783. <https://doi.org/10.1007/BF02980631>.
- (173) Atkins, P.; Friedman, R. *Molecular Quantum Mechanics*, 4th ed.; Oxford University Press: New York, NY, **2005**.
- (174) Slater, J. C. The Theory of Complex Spectra. *Phys. Rev.* **1929**, 34 (10), 1293–1322. <https://doi.org/10.1103/PhysRev.34.1293>.
- (175) Boys, S. F. Electronic Wave Functions - I. A General Method of Calculation for the Stationary States of Any Molecular System. *Proc. R. Soc. London. Ser. A. Math. Phys. Sci.*

- 1950**, 200 (1063), 542–554. <https://doi.org/10.1098/rspa.1950.0036>.
- (176) Slater, J. C. Atomic Shielding Constants. *Phys. Rev.* **1930**, 36 (1), 57–64. <https://doi.org/10.1103/PhysRev.36.57>.
- (177) Averill, F. W.; Ellis, D. E. An Efficient Numerical Multicenter Basis Set for Molecular Orbital Calculations: Application to FeCl₄. *J. Chem. Phys.* **1973**, 59 (12), 6412–6418. <https://doi.org/10.1063/1.1680020>.
- (178) Thomas, L. H. The Calculation of Atomic Fields. *Math. Proc. Cambridge Philos. Soc.* **1927**, 23 (5), 542–548. <https://doi.org/10.1017/S0305004100011683>.
- (179) Fermi, E. Un Metodo Statistico per La Determinazione Di Alcune Priorieta Dell'atome. *Rend. Accad. Naz. Lincei* **1927**, 6 (602–607), 32.
- (180) Dirac, P. A. M. Note on Exchange Phenomena in the Thomas Atom. *Math. Proc. Cambridge Philos. Soc.* **1930**, 26 (3), 376–385. <https://doi.org/10.1017/S0305004100016108>.
- (181) Bloch, F. Über Die Quantenmechanik Der Elektronen in Kristallgittern. *Zeitschrift für Phys.* **1929**, 52 (7–8), 555–600. <https://doi.org/10.1007/BF01339455>.
- (182) Lundqvist, B. I.; Andersson, Y.; Shao, H.; Chan, S.; Langreth, D. C. Density Functional Theory Including Van Der Waals Forces. *Int. J. Quantum Chem.* **1995**, 56 (4), 247–255. <https://doi.org/10.1002/qua.560560410>.
- (183) Dion, M.; Rydberg, H.; Schröder, E.; Langreth, D. C.; Lundqvist, B. I. Van Der Waals Density Functional for General Geometries. *Phys. Rev. Lett.* **2004**, 92 (24), 246401. <https://doi.org/10.1103/PhysRevLett.92.246401>.
- (184) Sabatini, R.; Gorni, T.; de Gironcoli, S. Nonlocal van Der Waals Density Functional Made Simple and Efficient. *Phys. Rev. B* **2013**, 87 (4), 041108. <https://doi.org/10.1103/PhysRevB.87.041108>.
- (185) Grimme, S. Semiempirical GGA-type Density Functional Constructed with a Long-range Dispersion Correction. *J. Comput. Chem.* **2006**, 27 (15), 1787–1799. <https://doi.org/10.1002/jcc.20495>.
- (186) Grimme, S.; Antony, J.; Ehrlich, S.; Krieg, H. A Consistent and Accurate Ab Initio Parametrization of Density Functional Dispersion Correction (DFT-D) for the 94 Elements H-Pu. *J. Chem. Phys.* **2010**, 132 (15), 154104. <https://doi.org/10.1063/1.3382344>.
- (187) Vydrov, O. A.; Van Voorhis, T. Nonlocal van Der Waals Density Functional: The Simpler the Better. *J. Chem. Phys.* **2010**, 133 (24), 244103. <https://doi.org/10.1063/1.3521275>.
- (188) Perdew, J. P.; Chevary, J. A.; Vosko, S. H.; Jackson, K. A.; Pederson, M. R.; Singh, D. J.; Fiolhais, C. Atoms, Molecules, Solids, and Surfaces: Applications of the Generalized Gradient Approximation for Exchange and Correlation. *Phys. Rev. B* **1992**, 46 (11), 6671–6687. <https://doi.org/10.1103/PhysRevB.46.6671>.
- (189) Heyd, J.; Scuseria, G. E.; Ernzerhof, M. Hybrid Functionals Based on a Screened Coulomb Potential. *J. Chem. Phys.* **2003**, 118 (18), 8207–8215. <https://doi.org/10.1063/1.1564060>.
- (190) Heyd, J.; Scuseria, G. E.; Ernzerhof, M. Erratum: “Hybrid Functionals Based on a Screened Coulomb Potential” [J. Chem. Phys. 118, 8207 (2003)]. *J. Chem. Phys.* **2006**, 124 (21), 219906. <https://doi.org/10.1063/1.2204597>.
- (191) Berland, K.; Hyldgaard, P. Exchange Functional That Tests the Robustness of the Plasmon Description of the van Der Waals Density Functional. *Phys. Rev. B - Condens. Matter Mater. Phys.* **2014**, 89 (3), 1–8. <https://doi.org/10.1103/PhysRevB.89.035412>.
- (192) Berland, K.; Jiao, Y.; Lee, J. H.; Rangel, T.; Neaton, J. B.; Hyldgaard, P. Assessment of Two Hybrid van Der Waals Density Functionals for Covalent and Non-Covalent Binding of Molecules. *J. Chem. Phys.* **2017**, 146 (23). <https://doi.org/10.1063/1.4986522>.
- (193) Shukla, V.; Jiao, Y.; Lee, J.; Schröder, E.; Neaton, J. B.; Hyldgaard, P. Accurate

- Nonempirical Range-Separated Hybrid van Der Waals Density Functional for Complex Molecular Problems, Solids, and Surfaces. *Phys. Rev. X* **2022**, 12 (4), 041003. <https://doi.org/10.1103/PhysRevX.12.041003>.
- (194) Shukla, V.; Jiao, Y.; Frostenson, C. M.; Hyldgaard, P. VdW-DF-Ahcx: A Range-Separated van Der Waals Density Functional Hybrid. *J. Phys. Condens. Matter* **2022**, 34 (2), 025902. <https://doi.org/10.1088/1361-648X/ac2ad2>.
- (195) Lee, K.; Murray, É. D.; Kong, L.; Lundqvist, B. I.; Langreth, D. C. Higher-Accuracy van Der Waals Density Functional. *Phys. Rev. B - Condens. Matter Mater. Phys.* **2010**, 82 (8), 3–6. <https://doi.org/10.1103/PhysRevB.82.081101>.
- (196) Vanderbilt, D. Soft Self-Consistent Pseudopotentials in a Generalized Eigenvalue Formalism. *Phys. Rev. B* **1990**, 41 (11), 7892–7895. <https://doi.org/10.1103/PhysRevB.41.7892>.
- (197) Laasonen, K.; Car, R.; Lee, C.; Vanderbilt, D. Implementation of Ultrasoft Pseudopotentials in Ab Initio Molecular Dynamics. *Phys. Rev. B* **1991**, 43 (8), 6796–6799. <https://doi.org/10.1103/PhysRevB.43.6796>.
- (198) Pyykkö, P. Relativistic Effects in Chemistry: More Common Than You Thought. *Annu. Rev. Phys. Chem.* **2012**, 63 (1), 45–64. <https://doi.org/10.1146/annurev-physchem-032511-143755>.
- (199) Blöchl, P. E. Projector Augmented-Wave Method. *Phys. Rev. B* **1994**, 50 (24), 17953–17979. <https://doi.org/10.1103/PhysRevB.50.17953>.
- (200) Kresse, G.; Joubert, D. From Ultrasoft Pseudopotentials to the Projector Augmented-Wave Method. *Phys. Rev. B* **1999**, 59 (3), 1758–1775. <https://doi.org/10.1103/PhysRevB.59.1758>.
- (201) Busnengo, H. F.; Salin, A.; Dong, W. Representation of the 6D Potential Energy Surface for a Diatomic Molecule near a Solid Surface. *J. Chem. Phys.* **2000**, 112 (17), 7641–7651. <https://doi.org/10.1063/1.481377>.
- (202) Olsen, R. A.; Busnengo, H. F.; Salin, A.; Somers, M. F.; Kroes, G. J.; Baerends, E. J. Constructing Accurate Potential Energy Surfaces for a Diatomic Molecule Interacting with a Solid Surface: H₂+Pt(111) and H₂+Cu(100). *J. Chem. Phys.* **2002**, 116 (9), 3841–3855. <https://doi.org/10.1063/1.1446852>.
- (203) Wijzenbroek, M. Hydrogen Dissociation on Metal Surfaces, Leiden University, **2016**. <https://hdl.handle.net/1887/39935>.
- (204) Smeets, E. W. F. Development of Highly Accurate Density Functionals for H₂ Dissociation on Transition Metals, Leiden University, **2021**. <https://hdl.handle.net/1887/3193529>.
- (205) Verlet, L. Computer “Experiments” on Classical Fluids. I. Thermodynamical Properties of Lennard-Jones Molecules. *Phys. Rev.* **1967**, 159 (1), 98–103. <https://doi.org/10.1103/PhysRev.159.98>.
- (206) Bulirsch, R.; Stoer, J. Numerical Treatment of Ordinary Differential Equations by Extrapolation Methods. *Numer. Math.* **1966**, 8 (1), 1–13. <https://doi.org/10.1007/BF02165234>.
- (207) Stoer, J.; Bulirsch, R. *Introduction to Numerical Analysis*; Texts in Applied Mathematics; Springer New York: New York, NY, **2002**; Vol. 12. <https://doi.org/10.1007/978-0-387-21738-3>.
- (208) Richardson, L. F. IX. The Approximate Arithmetical Solution by Finite Differences of Physical Problems Involving Differential Equations, with an Application to the Stresses in a Masonry Dam. *Philos. Trans. R. Soc. London. Ser. A, Contain. Pap. a Math. or Phys. Character* **1911**, 210 (459–470), 307–357. <https://doi.org/10.1098/rsta.1911.0009>.
- (209) Smits, B. Surface Temperature and the Dynamics of H₂ on Cu(111), Leiden University, **2023**. <https://hdl.handle.net/1887/3628423>.

- (210) Karplus, M.; Porter, R. N.; Sharma, R. D. Exchange Reactions with Activation Energy. I. Simple Barrier Potential for (H, H₂). *J. Chem. Phys.* **1965**, *43* (9), 3259–3287. <https://doi.org/10.1063/1.1697301>.
- (211) Raff, L. M.; Karplus, M. Theoretical Investigations of Reactive Collisions in Molecular Beams: K+CH₃I and Related Systems. *J. Chem. Phys.* **1966**, *44* (3), 1212–1229. <https://doi.org/10.1063/1.1726808>.
- (212) Sabbe, M. K.; Reyniers, M.-F.; Reuter, K. First-Principles Kinetic Modeling in Heterogeneous Catalysis: An Industrial Perspective on Best-Practice, Gaps and Needs. *Catal. Sci. Technol.* **2012**, *2* (10), 2010–2024. <https://doi.org/10.1039/c2cy20261a>.
- (213) Wolcott, C. A.; Medford, A. J.; Studt, F.; Campbell, C. T. Degree of Rate Control Approach to Computational Catalyst Screening. *J. Catal.* **2015**, *330*, 197–207. <https://doi.org/10.1016/j.jcat.2015.07.015>.
- (214) Gartland, P. O. Adsorption of Oxygen on Clean Single Crystal Faces of Aluminium. *Surf. Sci.* **1977**, *62* (1), 183–196. [https://doi.org/10.1016/0039-6028\(77\)90436-8](https://doi.org/10.1016/0039-6028(77)90436-8).
- (215) Roberts, M. W. Chemisorption and Reaction Pathways at Metal Surfaces: The Role of Surface Oxygen. *Chem. Soc. Rev.* **1989**, *18*, 451–475. <https://doi.org/10.1039/cs9891800451>.
- (216) Lundgren, E.; Mikkelsen, A.; Andersen, J. N.; Kresse, G.; Schmid, M.; Varga, P. Surface Oxides on Close-Packed Surfaces of Late Transition Metals. *J. Phys. Condens. Matter* **2006**, *18* (30), R481–R499. <https://doi.org/10.1088/0953-8984/18/30/R01>.
- (217) Stampfl, C.; Soon, A.; Piccinin, S.; Shi, H.; Zhang, H. Bridging the Temperature and Pressure Gaps: Close-Packed Transition Metal Surfaces in an Oxygen Environment. *J. Phys. Condens. Matter* **2008**, *20* (18), 184021. <https://doi.org/10.1088/0953-8984/20/18/184021>.
- (218) Montemore, M. M.; van Spronsen, M. A.; Madix, R. J.; Friend, C. M. O₂ Activation by Metal Surfaces: Implications for Bonding and Reactivity on Heterogeneous Catalysts. *Chem. Rev.* **2018**, *118* (5), 2816–2862. <https://doi.org/10.1021/acs.chemrev.7b00217>.
- (219) Jacobsen, J.; Hammer, B.; Jacobsen, K. W.; Nørskov, J. K. Electronic Structure, Total Energies, and STM Images of Clean and Oxygen-Covered Al(111). *Phys. Rev. B* **1995**, *52* (20), 14954–14962. <https://doi.org/10.1103/PhysRevB.52.14954>.
- (220) Darling, G. R.; Holloway, S. The Dissociation of Diatomic Molecules at Surfaces. *Reports Prog. Phys.* **1995**, *58* (12), 1595–1672. <https://doi.org/10.1088/0034-4885/58/12/001>.
- (221) Wijzenbroek, M.; Somers, M. F. Static Surface Temperature Effects on the Dissociation of H₂ and D₂ on Cu(111). *J. Chem. Phys.* **2012**, *137* (5), 054703. <https://doi.org/10.1063/1.4738956>.
- (222) Batra, I. P.; Kleinman, L. Chemisorption of Oxygen on Aluminum Surfaces. *J. Electron Spectros. Relat. Phenomena* **1984**, *33* (3), 175–241. [https://doi.org/10.1016/0368-2048\(84\)80020-1](https://doi.org/10.1016/0368-2048(84)80020-1).
- (223) Kerkar, M.; Fisher, D.; Woodruff, D. P.; Cowie, B. Adsorption Site Determination for Oxygen on Al(111) Using Normal Incidence Standing X-Ray Wavefield Absorption. *Surf. Sci.* **1992**, *271* (1–2), 45–56. [https://doi.org/10.1016/0039-6028\(92\)90860-9](https://doi.org/10.1016/0039-6028(92)90860-9).
- (224) Österlund, L.; Zoric-acute, I.; Kasemo, B. Dissociative Sticking of O₂ on Al(111). *Phys. Rev. B* **1997**, *55* (23), 15452–15455. <https://doi.org/10.1103/PhysRevB.55.15452>.
- (225) Zhukov, V.; Popova, I.; Yates, J. T. Initial Stages of Al(111) Oxidation with Oxygen—Temperature Dependence of the Integral Reactive Sticking Coefficient. *Surf. Sci.* **1999**, *441* (2–3), 251–264. [https://doi.org/10.1016/S0039-6028\(99\)00614-7](https://doi.org/10.1016/S0039-6028(99)00614-7).
- (226) Liu, H.-R.; Xiang, H.; Gong, X. G. First Principles Study of Adsorption of O₂ on Al Surface with Hybrid Functionals. *J. Chem. Phys.* **2011**, *135* (21), 214702. <https://doi.org/10.1063/1.3665032>.
- (227) Kurahashi, M.; Yamauchi, Y. Steric Effect in O₂ Sticking on Al(111): Preference for

- Parallel Geometry. *Phys. Rev. Lett.* **2013**, *110* (24), 246102.
<https://doi.org/10.1103/PhysRevLett.110.246102>.
- (228) Libisch, F.; Huang, C.; Liao, P.; Pavone, M.; Carter, E. A. Origin of the Energy Barrier to Chemical Reactions of O₂ on Al(111): Evidence for Charge Transfer, Not Spin Selection. *Phys. Rev. Lett.* **2012**, *109* (19), 198303.
<https://doi.org/10.1103/PhysRevLett.109.198303>.
- (229) Cheng, J.; Libisch, F.; Carter, E. A. Dissociative Adsorption of O₂ on Al(111): The Role of Orientational Degrees of Freedom. *J. Phys. Chem. Lett.* **2015**, *6* (9), 1661–1665.
<https://doi.org/10.1021/acs.jpclett.5b00597>.
- (230) Yin, R.; Zhang, Y.; Libisch, F.; Carter, E. A.; Guo, H.; Jiang, B. Dissociative Chemisorption of O₂ on Al(111): Dynamics on a Correlated Wave-Function-Based Potential Energy Surface. *J. Phys. Chem. Lett.* **2018**, *9* (12), 3271–3277.
<https://doi.org/10.1021/acs.jpclett.8b01470>.
- (231) Behler, J.; Delley, B.; Lorenz, S.; Reuter, K.; Scheffler, M. Dissociation of O₂ at Al(111): The Role of Spin Selection Rules. *Phys. Rev. Lett.* **2005**, *94* (3), 036104.
<https://doi.org/10.1103/PhysRevLett.94.036104>.
- (232) Fan, X. L.; Lau, W. M.; Liu, Z. F. Comment on “Dissociation of O₂ at Al(111): The Role of Spin Selection Rules.” *Phys. Rev. Lett.* **2006**, *96* (7), 079801.
<https://doi.org/10.1103/PhysRevLett.96.079801>.
- (233) Behler, J.; Reuter, K.; Scheffler, M. Behler, Reuter, and Scheffler Reply. *Phys. Rev. Lett.* **2006**, *96* (7), 79802. <https://doi.org/10.1103/PhysRevLett.96.079802>.
- (234) Carbogno, C.; Behler, J.; Groß, A.; Reuter, K. Fingerprints for Spin-Selection Rules in the Interaction Dynamics of O₂ at Al(111). *Phys. Rev. Lett.* **2008**, *101* (9), 096104.
<https://doi.org/10.1103/PhysRevLett.101.096104>.
- (235) Carbogno, C.; Behler, J.; Reuter, K.; Groß, A. Signatures of Nonadiabatic O₂- Dissociation at Al(111): First-Principles Fewest-Switches Study. *Phys. Rev. B* **2010**, *81* (3), 035410.
<https://doi.org/10.1103/PhysRevB.81.035410>.
- (236) Sasaki, T.; Ohno, T. Calculations of the Potential-Energy Surface for Dissociation Process of O₂ on the Al(111) Surface. *Phys. Rev. B* **1999**, *60* (11), 7824–7827.
<https://doi.org/10.1103/PhysRevB.60.7824>.
- (237) Honkala, K.; Laasonen, K. Oxygen Molecule Dissociation on the Al(111) Surface. *Phys. Rev. Lett.* **2000**, *84* (4), 705–708. <https://doi.org/10.1103/PhysRevLett.84.705>.
- (238) Yourdshahyan, Y.; Razaznejad, B.; Lundqvist, B. I. Adiabatic Potential-Energy Surfaces for Oxygen on Al(111). *Phys. Rev. B* **2002**, *65* (7), 075416.
<https://doi.org/10.1103/PhysRevB.65.075416>.
- (239) Hammer, B.; Hansen, L. B.; Nørskov, J. K. Improved Adsorption Energetics within Density-Functional Theory Using Revised Perdew-Burke-Ernzerhof Functionals. *Phys. Rev. B* **1999**, *59* (11), 7413–7421. <https://doi.org/10.1103/PhysRevB.59.7413>.
- (240) Behler, J.; Reuter, K.; Scheffler, M. Nonadiabatic Effects in the Dissociation of Oxygen Molecules at the Al(111) Surface. *Phys. Rev. B* **2008**, *77* (11), 115421.
<https://doi.org/10.1103/PhysRevB.77.115421>.
- (241) Huang, C.; Pavone, M.; Carter, E. A. Quantum Mechanical Embedding Theory Based on a Unique Embedding Potential. *J. Chem. Phys.* **2011**, *134* (15), 154110.
<https://doi.org/10.1063/1.3577516>.
- (242) Andersson, K.; Malmqvist, P. A.; Roos, B. O.; Sadlej, A. J.; Wolinski, K. Second-Order Perturbation Theory with a CASSCF Reference Function. *J. Phys. Chem.* **1990**, *94* (14), 5483–5488. <https://doi.org/10.1021/j100377a012>.
- (243) Andersson, K.; Malmqvist, P.-Å.; Roos, B. O. Second-Order Perturbation Theory with a Complete Active Space Self-Consistent Field Reference Function. *J. Chem. Phys.* **1992**, *96* (2), 1218–1226. <https://doi.org/10.1063/1.462209>.

- (244) Perdew, J. P.; Parr, R. G.; Levy, M.; Balduz, J. L. Density-Functional Theory for Fractional Particle Number: Derivative Discontinuities of the Energy. *Phys. Rev. Lett.* **1982**, *49* (23), 1691–1694. <https://doi.org/10.1103/PhysRevLett.49.1691>.
- (245) Perdew, J. P.; Zunger, A. Self-Interaction Correction to Density-Functional Approximations for Many-Electron Systems. *Phys. Rev. B* **1981**, *23* (10), 5048–5079. <https://doi.org/10.1103/PhysRevB.23.5048>.
- (246) Menezes, F.; Kats, D.; Werner, H.-J. Local Complete Active Space Second-Order Perturbation Theory Using Pair Natural Orbitals (PNO-CASPT2). *J. Chem. Phys.* **2016**, *145* (12). <https://doi.org/10.1063/1.4963019>.
- (247) McCreery, J. H.; Wolken, G. A Model Potential for Chemisorption: H₂ + W(001). *J. Chem. Phys.* **1975**, *63* (6), 2340–2349. <https://doi.org/10.1063/1.431663>.
- (248) Martin-Gondre, L.; Crespos, C.; Larrégaray, P.; Rayez, J. C.; Conte, D.; van Ootegem, B. Detailed Description of the Flexible Periodic London–Eyring–Polanyi–Sato Potential Energy Function. *Chem. Phys.* **2010**, *367* (2–3), 136–147. <https://doi.org/10.1016/j.chemphys.2009.11.012>.
- (249) Wei, Z.; Martínez, J. M. P.; Carter, E. A. Introducing the Embedded Random Phase Approximation: H₂ Dissociative Adsorption on Cu(111) as an Exemplar. *J. Chem. Phys.* **2023**, *159* (19), 194108. <https://doi.org/10.1063/5.0181229>.
- (250) Mori-Sánchez, P.; Cohen, A. J.; Yang, W. Many-Electron Self-Interaction Error in Approximate Density Functionals. *J. Chem. Phys.* **2006**, *125* (20). <https://doi.org/10.1063/1.2403848>.
- (251) Mori-Sánchez, P.; Cohen, A. J.; Yang, W. Localization and Delocalization Errors in Density Functional Theory and Implications for Band-Gap Prediction. *Phys. Rev. Lett.* **2008**, *100* (14), 146401. <https://doi.org/10.1103/PhysRevLett.100.146401>.
- (252) Zheng, J.; Zhao, Y.; Truhlar, D. G. The DBH24/08 Database and Its Use to Assess Electronic Structure Model Chemistries for Chemical Reaction Barrier Heights. *J. Chem. Theory Comput.* **2009**, *5* (4), 808–821. <https://doi.org/10.1021/ct800568m>.
- (253) Sim, E.; Song, S.; Burke, K. Quantifying Density Errors in DFT. *J. Phys. Chem. Lett.* **2018**, *9* (22), 6385–6392. <https://doi.org/10.1021/acs.jpclett.8b02855>.
- (254) Karikorpi, M.; Holloway, S.; Henriksen, N.; Nørskov, J. K. Dynamics of Molecule-Surface Interactions. *Surf. Sci. Lett.* **1987**, *179* (1), L41–L48. [https://doi.org/10.1016/0167-2584\(87\)90235-0](https://doi.org/10.1016/0167-2584(87)90235-0).
- (255) Perdew, J. P.; Ruzsinszky, A.; Tao, J.; Staroverov, V. N.; Scuseria, G. E.; Csonka, G. I. Prescription for the Design and Selection of Density Functional Approximations: More Constraint Satisfaction with Fewer Fits. *J. Chem. Phys.* **2005**, *123* (6), 062201. <https://doi.org/10.1063/1.1904565>.
- (256) Wijzenbroek, M.; Kroes, G. J. The Effect of the Exchange–Correlation Functional on H₂ Dissociation on Ru(0001). *J. Chem. Phys.* **2014**, *140* (8). <https://doi.org/10.1063/1.4865946>.
- (257) Cohen, A. J.; Mori-Sánchez, P.; Yang, W. Challenges for Density Functional Theory. *Chem. Rev.* **2012**, *112* (1), 289–320. <https://doi.org/10.1021/cr200107z>.
- (258) Kresse, G.; Hafner, J. Ab Initio Molecular Dynamics for Liquid Metals. *Phys. Rev. B* **1993**, *47* (1), 558–561. <https://doi.org/10.1103/PhysRevB.47.558>.
- (259) Kresse, G.; Hafner, J. Ab Initio Molecular-Dynamics Simulation of the Liquid-Metal–Amorphous-Semiconductor Transition in Germanium. *Phys. Rev. B* **1994**, *49* (20), 14251–14269. <https://doi.org/10.1103/PhysRevB.49.14251>.
- (260) Kresse, G.; Hafner, J. Norm-Conserving and Ultrasoft Pseudopotentials for First-Row and Transition Elements. *J. Phys. Condens. Matter* **1994**, *6* (40), 8245–8257. <https://doi.org/10.1088/0953-8984/6/40/015>.
- (261) Kresse, G.; Furthmüller, J. Efficiency of Ab-Initio Total Energy Calculations for Metals

- and Semiconductors Using a Plane-Wave Basis Set. *Comput. Mater. Sci.* **1996**, 6 (1), 15–50. [https://doi.org/10.1016/0927-0256\(96\)00008-0](https://doi.org/10.1016/0927-0256(96)00008-0).
- (262) Kresse, G.; Furthmüller, J. Efficient Iterative Schemes for Ab Initio Total-Energy Calculations Using a Plane-Wave Basis Set. *Phys. Rev. B* **1996**, 54 (16), 11169–11186. <https://doi.org/10.1103/PhysRevB.54.11169>.
- (263) Wijzenbroek, M.; Klein, D. M.; Smits, B.; Somers, M. F.; Kroes, G. J. Performance of a Non-Local van Der Waals Density Functional on the Dissociation of H₂ on Metal Surfaces. *J. Phys. Chem. A* **2015**, 119 (50), 12146–12158. <https://doi.org/10.1021/acs.jpca.5b06008>.
- (264) Kuebler, N. A.; Robin, M. B.; Yang, J. J.; Gedanken, A.; Herrick, D. R. Fully Resolved Zeeman Pattern in the Stern-Gerlach Deflection Spectrum of O₂ (3g-, K=1). *Phys. Rev. A* **1988**, 38 (2), 737–749. <https://doi.org/10.1103/PhysRevA.38.737>.
- (265) Powell, M. J. *Numerical Methods for Nonlinear Algebraic Equations*; Rabinowitz, P., Ed.; Gordon and Breach: New York, 1970.
- (266) Garbow, B. S.; Hillstorn, K. E.; More, J. J. MINPACK Subroutine HYBRD. **1980**.
- (267) Schmid, M.; Leonardelli, G.; Tscheließnig, R.; Biedermann, A.; Varga, P. Oxygen Adsorption on Al(111): Low Transient Mobility. *Surf. Sci.* **2001**, 478 (3), L355–L362. [https://doi.org/10.1016/S0039-6028\(01\)00967-0](https://doi.org/10.1016/S0039-6028(01)00967-0).
- (268) Wahnström, G.; Lee, A. B.; Strömquist, J. Motion of “Hot” Oxygen Adatoms on Corrugated Metal Surfaces. *J. Chem. Phys.* **1996**, 105 (1), 326–336. <https://doi.org/10.1063/1.471877>.
- (269) Binetti, M.; Weiße, O.; Hasselbrink, E.; Komrowski, A. J.; Kummel, A. C. Abstractive Chemisorption of O₂ on Al(111). *Faraday Discuss.* **2000**, 117 (111), 313–320. <https://doi.org/10.1039/b004006l>.
- (270) Binetti, M.; Weiße, O.; Hasselbrink, E.; Katz, G.; Kosloff, R.; Zeiri, Y. The Role of Nonadiabatic Pathways and Molecular Rotations in the Oxygen Abstraction Reaction on the Al(111) Surface. *Chem. Phys. Lett.* **2003**, 373 (3–4), 366–371. [https://doi.org/10.1016/S0009-2614\(03\)00586-4](https://doi.org/10.1016/S0009-2614(03)00586-4).
- (271) Binetti, M.; Hasselbrink, E. Abstraction of Oxygen from Dioxygen on Al(111) Revealed by Resonant Multiphoton Ionization Laser Spectrometry. *J. Phys. Chem. B* **2004**, 108 (38), 14677–14684. <https://doi.org/10.1021/jp049197y>.
- (272) Strandberg, M. W. P.; Tinkham, M. Theory of the Fine Structure of the Molecular. *Phys. Rev.* **1955**, 97 (4), 937–966.
- (273) Polanyi, J. C. Some Concepts in Reaction Dynamics. *Science (80-)*. **1987**, 236 (4802), 680–690. <https://doi.org/10.1126/science.236.4802.680>.
- (274) Tiwari, A. K.; Nave, S.; Jackson, B. Methane Dissociation on Ni(111): A New Understanding of the Lattice Effect. *Phys. Rev. Lett.* **2009**, 103 (25), 253201. <https://doi.org/10.1103/PhysRevLett.103.253201>.
- (275) Tiwari, A. K.; Nave, S.; Jackson, B. The Temperature Dependence of Methane Dissociation on Ni(111) and Pt(111): Mixed Quantum-Classical Studies of the Lattice Response. *J. Chem. Phys.* **2010**, 132 (13), 134702. <https://doi.org/10.1063/1.3357415>.
- (276) Luntz, A. C.; Williams, M. D.; Bethune, D. S. The Sticking of O₂ on a Pt(111) Surface. *J. Chem. Phys.* **1988**, 89 (7), 4381–4395. <https://doi.org/10.1063/1.454824>.
- (277) Jacobse, L.; den Dunnen, A.; Juurlink, L. B. F. The Molecular Dynamics of Adsorption and Dissociation of O₂ on Pt(553). *J. Chem. Phys.* **2015**, 143 (1). <https://doi.org/10.1063/1.4923006>.
- (278) Gostein, M.; Sitz, G. O. Rotational State-Resolved Sticking Coefficients for H₂ on Pd(111): Testing Dynamical Steering in Dissociative Adsorption. *J. Chem. Phys.* **1997**, 106 (17), 7378–7390. <https://doi.org/10.1063/1.473699>.
- (279) Perdew, J. P. Climbing the Ladder of Density Functional Approximations. *MRS Bull.*

- 2013**, 38 (9), 743–750. <https://doi.org/10.1557/mrs.2013.178>.
- (280) van Bree, R. A. B.; Gerrits, N.; Kroes, G.-J. Dissociative Chemisorption of O₂ on Al(111): Dynamics on a Potential Energy Surface Computed with a Non-Self-Consistent Screened Hybrid Density Functional Approach. *Faraday Discuss.* **2024**, 251, 361–381. <https://doi.org/10.1039/D3FD00165B>.
- (281) Juaristi, J. I.; Alducin, M.; Muiño, R. D.; Busnengo, H. F.; Salin, A. Role of Electron-Hole Pair Excitations in the Dissociative Adsorption of Diatomic Molecules on Metal Surfaces. *Phys. Rev. Lett.* **2008**, 100 (11), 116102. <https://doi.org/10.1103/PhysRevLett.100.116102>.
- (282) Berland, K.; Cooper, V. R.; Lee, K.; Schröder, E.; Thonhauser, T.; Hyldgaard, P.; Lundqvist, B. I. Van Der Waals Forces in Density Functional Theory: A Review of the VdW-DF Method. *Reports Prog. Phys.* **2015**, 78 (6), 066501. <https://doi.org/10.1088/0034-4885/78/6/066501>.
- (283) Moldabekov, Z. A.; Lokamani, M.; Vorberger, J.; Cangi, A.; Dornheim, T. Non-Empirical Mixing Coefficient for Hybrid XC Functionals from Analysis of the XC Kernel. *J. Phys. Chem. Lett.* **2023**, 14 (5), 1326–1333. <https://doi.org/10.1021/acs.jpclett.2c03670>.
- (284) van Bree, R. A. B.; Kroes, G. J. O₂ Dissociation on Cu(111) Dynamics on a Novel Screened Hybrid van Der Waals DFT Potential Energy Surface. *J. Phys. Chem. C* **2024**, 128 (45), 19182–19196. <https://doi.org/10.1021/acs.jpcc.4c05466>.
- (285) Klimeš, J.; Bowler, D. R.; Michaelides, A. A Critical Assessment of Theoretical Methods for Finding Reaction Pathways and Transition States of Surface Processes. *J. Phys. Condens. Matter* **2010**, 22 (7), 074203. <https://doi.org/10.1088/0953-8984/22/7/074203>.
- (286) Klimeš, J.; Bowler, D. R.; Michaelides, A. Van Der Waals Density Functionals Applied to Solids. *Phys. Rev. B* **2011**, 83 (19), 195131. <https://doi.org/10.1103/PhysRevB.83.195131>.
- (287) Teter, M. P.; Payne, M. C.; Allan, D. C. Solution of Schrödinger’s Equation for Large Systems. *Phys. Rev. B* **1989**, 40 (18), 12255–12263. <https://doi.org/10.1103/PhysRevB.40.12255>.
- (288) Bylander, D. M.; Kleinman, L.; Lee, S. Self-Consistent Calculations of the Energy Bands and Bonding Properties of B12C3. *Phys. Rev. B* **1990**, 42 (2), 1394–1403. <https://doi.org/10.1103/PhysRevB.42.1394>.
- (289) *PRECFOCK - VASP Wiki*. <https://www.vasp.at/wiki/index.php/PRECFOCK> (accessed 2025-01-13).
- (290) Haas, P.; Tran, F.; Blaha, P. Calculation of the Lattice Constant of Solids with Semilocal Functionals. *Phys. Rev. B* **2009**, 79 (8), 085104. <https://doi.org/10.1103/PhysRevB.79.085104>.
- (291) Noonan, J. R.; Davis, H. L. Confirmation of an Exception to the “General Rule” of Surface Relaxations. *J. Vac. Sci. Technol. A Vacuum, Surfaces, Film.* **1990**, 8 (3), 2671–2676. <https://doi.org/10.1116/1.576692>.
- (292) Schöchlin, J.; Bohnen, K. P.; Ho, K. M. Structure and Dynamics at the Al(111)-Surface. *Surf. Sci.* **1995**, 324 (2–3), 113–121. [https://doi.org/10.1016/0039-6028\(94\)00710-1](https://doi.org/10.1016/0039-6028(94)00710-1).
- (293) Marcus, R. A. On the Analytical Mechanics of Chemical Reactions. Classical Mechanics of Linear Collisions. *J. Chem. Phys.* **1966**, 45 (12), 4500–4504. <https://doi.org/10.1063/1.1727529>.
- (294) McCullough, E. A.; Wyatt, R. E. Quantum Dynamics of the Collinear (H, H₂) Reaction. *J. Chem. Phys.* **1969**, 51 (3), 1253–1254. <https://doi.org/10.1063/1.1672133>.
- (295) Oudot, B.; Doblhoff-Dier, K. Reaction Barriers at Metal Surfaces Computed Using the Random Phase Approximation: Can We Beat DFT in the Generalized Gradient Approximation? *J. Chem. Phys.* **2024**, 161 (5), 054708.

- <https://doi.org/10.1063/5.0220465>.
- (296) Goerigk, L.; Grimme, S. Efficient and Accurate Double-Hybrid-Meta-GGA Density Functionals—Evaluation with the Extended GMTKN30 Database for General Main Group Thermochemistry, Kinetics, and Noncovalent Interactions. *J. Chem. Theory Comput.* **2011**, 7 (2), 291–309. <https://doi.org/10.1021/ct100466k>.
 - (297) Baule, B. Theoretische Behandlung Der Erscheinungen in Verdünnten Gasen. *Ann. Phys.* **1914**, 349 (9), 145–176. <https://doi.org/10.1002/andp.19143490908>.
 - (298) Golibrzuch, K.; Bartels, N.; Auerbach, D. J.; Wodtke, A. M. The Dynamics of Molecular Interactions and Chemical Reactions at Metal Surfaces: Testing the Foundations of Theory. *Annu. Rev. Phys. Chem.* **2015**, 66 (1), 399–425. <https://doi.org/10.1146/annurev-physchem-040214-121958>.
 - (299) Head-Gordon, M.; Tully, J. C. Molecular Dynamics with Electronic Frictions. *J. Chem. Phys.* **1995**, 103 (23), 10137–10145. <https://doi.org/10.1063/1.469915>.
 - (300) Alducin, M.; Díez Muiño, R.; Juaristi, J. I. Non-Adiabatic Effects in Elementary Reaction Processes at Metal Surfaces. *Prog. Surf. Sci.* **2017**, 92 (4), 317–340. <https://doi.org/10.1016/j.progsurf.2017.09.002>.
 - (301) White, J. D.; Chen, J.; Matsiev, D.; Auerbach, D. J.; Wodtke, A. M. Conversion of Large-Amplitude Vibration to Electron Excitation at a Metal Surface. *Nature* **2005**, 433 (7025), 503–505. <https://doi.org/10.1038/nature03213>.
 - (302) Shenvi, N.; Roy, S.; Tully, J. C. Nonadiabatic Dynamics at Metal Surfaces: Independent-Electron Surface Hopping. *J. Chem. Phys.* **2009**, 130 (17), 174107. <https://doi.org/10.1063/1.3125436>.
 - (303) Shakouri, K.; Behler, J.; Meyer, J.; Kroes, G.-J. Accurate Neural Network Description of Surface Phonons in Reactive Gas–Surface Dynamics: N₂ + Ru(0001). *J. Phys. Chem. Lett.* **2017**, 8 (10), 2131–2136. <https://doi.org/10.1021/acs.jpclett.7b00784>.
 - (304) Liu, Q.; Zhou, X.; Zhou, L.; Zhang, Y.; Luo, X.; Guo, H.; Jiang, B. Constructing High-Dimensional Neural Network Potential Energy Surfaces for Gas–Surface Scattering and Reactions. *J. Phys. Chem. C* **2018**, 122 (3), 1761–1769. <https://doi.org/10.1021/acs.jpcc.7b12064>.
 - (305) Scoles, G. *Atomic and Molecular Beam Methods*, v. 1.; Oxford University Press, 1988.
 - (306) Gattinoni, C.; Michaelides, A. Atomistic Details of Oxide Surfaces and Surface Oxidation: The Example of Copper and Its Oxides. *Surf. Sci. Rep.* **2015**, 70 (3), 424–447. <https://doi.org/10.1016/j.surfrep.2015.07.001>.
 - (307) Over, H.; Kim, Y. D.; Seitsonen, A. P.; Wendt, S.; Lundgren, E.; Schmid, M.; Varga, P.; Morgante, A.; Ertl, G. Atomic-Scale Structure and Catalytic Reactivity of the RuO₂ (110) Surface. *Science* (80-.). **2000**, 287 (5457), 1474–1476. <https://doi.org/10.1126/science.287.5457.1474>.
 - (308) Berger, H. F.; Leisch, M.; Winkler, A.; Rendulic, K. D. A Search for Vibrational Contributions to the Activated Adsorption of H₂ on Copper. *Chem. Phys. Lett.* **1990**, 175 (5), 425–428. [https://doi.org/10.1016/0009-2614\(90\)85558-T](https://doi.org/10.1016/0009-2614(90)85558-T).
 - (309) Rettner, C. T.; Michelsen, H. A.; Auerbach, D. J. Quantum-State-Specific Dynamics of the Dissociative Adsorption and Associative Desorption of H₂ at a Cu(111) Surface. *J. Chem. Phys.* **1995**, 102 (11), 4625–4641. <https://doi.org/10.1063/1.469511>.
 - (310) Hou, H.; Guldung, S. J.; Rettner, C. T.; Wodtke, A. M.; Auerbach, D. J. The Stereodynamics of a Gas-Surface Reaction. *Science* (80-.). **1997**, 277 (5322), 80–82. <https://doi.org/10.1126/science.277.5322.80>.
 - (311) Murphy, M. J.; Hodgson, A. Adsorption and Desorption Dynamics of H₂ and D₂ on Cu(111): The Role of Surface Temperature and Evidence for Corrugation of the Dissociation Barrier. *J. Chem. Phys.* **1998**, 108 (10), 4199–4211. <https://doi.org/10.1063/1.475818>.

- (312) Kaufmann, S.; Shuai, Q.; Auerbach, D. J.; Schwarzer, D.; Wodtke, A. M. Associative Desorption of Hydrogen Isotopologues from Copper Surfaces: Characterization of Two Reaction Mechanisms. *J. Chem. Phys.* **2018**, *148* (19), 194703. <https://doi.org/10.1063/1.5025666>.
- (313) Habraken, F. H. P. M.; Kieffer, E. P.; Bootsma, G. A. A Study of the Kinetics of the Interactions of O₂ and N₂O with a Cu(111) Surface and of the Reaction of CO with Adsorbed Oxygen Using Aes, LEED and Ellipsometry. *Surf. Sci.* **1979**, *83* (1), 45–59. [https://doi.org/10.1016/0039-6028\(79\)90479-5](https://doi.org/10.1016/0039-6028(79)90479-5).
- (314) JUNELL, P.; AHONEN, M.; HIRSIMÄKI, M.; VALDEN, M. INFLUENCE OF SURFACE MODIFICATION ON THE ADSORPTION DYNAMICS OF O₂ ON Cu {100}. *Surf. Rev. Lett.* **2004**, *11* (04n05), 457–461. <https://doi.org/10.1142/S0218625X0400627X>.
- (315) Hodgson, A.; Lewin, A. K.; Nesbitt, A. Dissociative Chemisorption of O₂ on Cu(110). *Surf. Sci.* **1993**, *293* (3), 211–226. [https://doi.org/10.1016/0039-6028\(93\)90315-B](https://doi.org/10.1016/0039-6028(93)90315-B).
- (316) Zhu, L.; Zhang, Y.; Zhang, L.; Zhou, X.; Jiang, B. Unified and Transferable Description of Dynamics of H₂ Dissociative Adsorption on Multiple Copper Surfaces via Machine Learning. *Phys. Chem. Chem. Phys.* **2020**, *22* (25), 13958–13964. <https://doi.org/10.1039/D0CP02291H>.
- (317) Galparsoro, O.; Kaufmann, S.; Auerbach, D. J.; Kandratsenka, A.; Wodtke, A. M. First Principles Rates for Surface Chemistry Employing Exact Transition State Theory: Application to Recombinative Desorption of Hydrogen from Cu(111). *Phys. Chem. Chem. Phys.* **2020**, *22* (31), 17532–17539. <https://doi.org/10.1039/D0CP02858D>.
- (318) Martin-Gondre, L.; Crespos, C.; Larrégaray, P. Dynamics of Dissociative Chemisorption of O₂ on Cu(100) Surface: A Theoretical Study. *Surf. Sci.* **2019**, *688*, 45–50. <https://doi.org/10.1016/j.susc.2019.05.006>.
- (319) Ramos, M.; Díaz, C.; Martínez, A. E.; Busnengo, H. F.; Martín, F. Dissociative and Non-Dissociative Adsorption of O₂ on Cu(111) and Cu ML /Ru(0001) Surfaces: Adiabaticity Takes Over. *Phys. Chem. Chem. Phys.* **2017**, *19* (16), 10217–10221. <https://doi.org/10.1039/C7CP00753A>.
- (320) Fallaque, J. G.; Ramos, M.; Busnengo, H. F.; Martín, F.; Díaz, C. Normal and Off-Normal Incidence Dissociative Dynamics of O₂ (v , J) on Ultrathin Cu Films Grown on Ru(0001). *Phys. Chem. Chem. Phys.* **2021**, *23* (13), 7768–7776. <https://doi.org/10.1039/D0CP03979A>.
- (321) Mondal, A.; Wijzenbroek, M.; Bonfanti, M.; Díaz, C.; Kroes, G. J. Thermal Lattice Expansion Effect on Reactive Scattering of H₂ from Cu(111) at Ts = 925 K. *J. Phys. Chem. A* **2013**, *117* (36), 8770–8781. <https://doi.org/10.1021/jp4042183>.
- (322) Sueyoshi, T.; Sasaki, T.; Iwasawa, Y. Molecular and Atomic Adsorption States of Oxygen on Cu(111) at 100–300 K. *Surf. Sci.* **1996**, *365* (2), 310–318. [https://doi.org/10.1016/0039-6028\(96\)00738-8](https://doi.org/10.1016/0039-6028(96)00738-8).
- (323) Xu, Y.; Mavrikakis, M. Adsorption and Dissociation of O₂ on Cu(111): Thermochemistry, Reaction Barrier and the Effect of Strain. *Surf. Sci.* **2001**, *494* (2), 131–144. [https://doi.org/10.1016/S0039-6028\(01\)01464-9](https://doi.org/10.1016/S0039-6028(01)01464-9).
- (324) Minniti, M.; Farías, D.; Perna, P.; Miranda, R. Enhanced Selectivity towards O₂ and H₂ Dissociation on Ultrathin Cu Films on Ru(0001). *J. Chem. Phys.* **2012**, *137* (7), 074706. <https://doi.org/10.1063/1.4746942>.
- (325) Zhang, D.; Jansen, C.; Kleyn, A. W.; Juurlink, L. B. F. Adsorption Dynamics of O₂ on Cu(111): A Supersonic Molecular Beam Study. *Phys. Chem. Chem. Phys.* **2023**, *25* (21), 14862–14868. <https://doi.org/10.1039/D3CP01215H>.
- (326) Hall, J.; Saksager, O.; Chorkendorff, I. Dissociative Chemisorption of O₂ on Cu(100). Effects of Mechanical Energy Transfer and Recoil. *Chem. Phys. Lett.* **1993**, *216* (3–6), 413–417. [https://doi.org/10.1016/0009-2614\(93\)90119-L](https://doi.org/10.1016/0009-2614(93)90119-L).

- (327) Busnengo, H. F.; Di Césare, M. A.; Dong, W.; Salin, A. Surface Temperature Effects in Dynamic Trapping Mediated Adsorption of Light Molecules on Metal Surfaces: H₂ on Pd(111) and Pd(110). *Phys. Rev. B* **2005**, 72 (12), 125411. <https://doi.org/10.1103/PhysRevB.72.125411>.
- (328) Zhang, Y.; Yang, W. A Challenge for Density Functionals: Self-Interaction Error Increases for Systems with a Noninteger Number of Electrons. *J. Chem. Phys.* **1998**, 109 (7), 2604–2608. <https://doi.org/10.1063/1.476859>.
- (329) Lynch, B. J.; Fast, P. L.; Harris, M.; Truhlar, D. G. Adiabatic Connection for Kinetics. *J. Phys. Chem. A* **2000**, 104 (21), 4811–4815. <https://doi.org/10.1021/jp000497z>.
- (330) Chae, K. H.; Lu, H. C.; Gustafsson, T. Medium-Energy Ion-Scattering Study of the Temperature Dependence of the Structure of Cu(111). *Phys. Rev. B* **1996**, 54 (19), 14082–14086. <https://doi.org/10.1103/PhysRevB.54.14082>.
- (331) Wu, L. Y.; Roman, M. J.; Heazlewood, B. R.; Kurahashi, M. Steric Effects in the Adsorption of O₂ on a Cu(111) Surface. *Phys. Chem. Chem. Phys.* **2025**. <https://doi.org/10.1039/D4CP04595E>.
- (332) Taleb, A. Al; Schiller, F.; Vyalikh, D. V.; María Pérez, J.; Auras, S. V.; Farías, D.; Ortega, J. E. Simulating High-Pressure Surface Reactions with Molecular Beams. *Phys. Chem. Chem. Phys.* **2023**, 26 (3), 1770–1776. <https://doi.org/10.1039/d3cp05071h>.
- (333) Zhou, Y.; Zhou, L.; Hu, X.; Xie, D. Dynamics Studies of O₂ Collision on Pt(111) Using a Global Potential Energy Surface. *J. Phys. Chem. C* **2020**, 124 (19), 10573–10583. <https://doi.org/10.1021/acs.jpcc.0c01247>.
- (334) Goikoetxea, I.; Juaristi, J. I.; Alducin, M.; Díez Muiño, R. Dissipative Effects in the Dynamics of N₂ on Tungsten Surfaces. *J. Phys. Condens. Matter* **2009**, 21 (26), 264007. <https://doi.org/10.1088/0953-8984/21/26/264007>.
- (335) Hand, M.; Harris, J. Recoil Effects in Surface Dissociation. *J. Chem. Phys.* **1990**, 92 (12), 7610–7617. <https://doi.org/10.1063/1.458198>.
- (336) Behler, J.; Parrinello, M. Generalized Neural-Network Representation of High-Dimensional Potential-Energy Surfaces. *Phys. Rev. Lett.* **2007**, 98 (14), 146401. <https://doi.org/10.1103/PhysRevLett.98.146401>.
- (337) Behler, J. Perspective: Machine Learning Potentials for Atomistic Simulations. *J. Chem. Phys.* **2016**, 145 (17), 170901. <https://doi.org/10.1063/1.4966192>.
- (338) Piessens, R.; de Doncker-Kapenga, E.; Überhuber, C. W.; Kahaner, D. K. *Quadpack*; Springer Series in Computational Mathematics; Springer Berlin Heidelberg: Berlin, Heidelberg, **1983**; Vol. 1. <https://doi.org/10.1007/978-3-642-61786-7>.

List of abbreviations, acronyms, and symbols

A list of abbreviations, acronyms, and symbols commonly used in the thesis:

AES:	<i>Auger electron spectroscopy</i>
AIMD:	<i>Ab initio molecular dynamics</i>
BOA:	<i>Born-Oppenheimer approximation</i>
BOMD:	<i>Born-Oppenheimer molecular dynamics</i>
BOSS:	<i>Born-Oppenheimer static surface</i>
BuSA:	<i>Burlisch–Stoer algorithm</i>
CRP:	<i>Corrugation reducing procedure</i>
COM:	<i>Centre-of-mass</i>
DC:	<i>Dissociative chemisorption</i>
DCM:	<i>Dynamic corrugation model</i>
DF:	<i>Density functional</i>
DFT:	<i>Density functional theory</i>
DOF:	<i>Degrees of freedom</i>
EA:	<i>Electron affinity</i>
E_{av}:	<i>Average energy of the molecular beam</i>
E_{ct}:	<i>Charge transfer energy; Φ - EA</i>
ECW:	<i>Embedded correlated wavefunction</i>
E^{HF}:	<i>Hartree Fock exchange energy</i>
Ehp:	<i>Electron-hole-pair</i>
E_i:	<i>Incidence energy</i>
$E_I^{\perp}$:	<i>Normal incidence energy</i>
E_{kin}:	<i>Kinetic energy</i>
E_{ro}:	<i>Rotational energy</i>
E_{rovib}:	<i>Rovibrational energy</i>
E_s:	<i>Stream energy of the molecular beam</i>
E_{vib}:	<i>Vibrational energy</i>
E_{xc}:	<i>Exchange-correlation energy</i>
FCC:	<i>Face-centred cubic</i>
FFT:	<i>Fast-Fourier transform</i>
GGA:	<i>Generalised gradient approximation</i>
GLO:	<i>Generalised Langevin oscillator</i>
HDNN:	<i>High dimensional neural-network</i>
HK:	<i>Hohenberg-Kohn</i>
HF:	<i>Hartree-Fock</i>
HREELS:	<i>High-resolution electron energy loss spectroscopy</i>

K&W:	<i>King & Wells</i>
KS:	<i>Kohn-Sham</i>
LDA:	<i>Local density approximation</i>
LEED:	<i>Low energy electron diffraction</i>
MD:	<i>Molecular dynamics</i>
mGGA:	<i>Meta-generalised gradient approximation</i>
N:	<i>Normal orientation; $\theta = 0^\circ$</i>
NSCF:	<i>Non-self-consistent field</i>
NES:	<i>Normal energy scaling</i>
P:	<i>Parallel orientation; $\theta = 90^\circ$</i>
PAW:	<i>Projector augmented wave</i>
PBC:	<i>Periodic Boundary Conditions</i>
P_E:	<i>Event probability</i>
PES:	<i>Potential energy surface</i>
P_r:	<i>Reaction probability</i>
P_r^D:	<i>Direct reaction probability</i>
P_r^I:	<i>Indirect reaction probability</i>
P_r^T:	<i>Total reaction probability</i>
P_s:	<i>Scattering probability</i>
P_t:	<i>Trapping probability</i>
QCT:	<i>Quasi-classical trajectory</i>
QD:	<i>Quantum dynamics</i>
Ref.:	<i>Reference</i>
RMSE:	<i>Root mean squared error</i>
S_0:	<i>Sticking probability; $1 - P_s$</i>
SCF:	<i>Self-consistent field</i>
SCM:	<i>Static corrugation model</i>
S_H:	<i>S_0 of helicoptering molecules</i>
S_P:	<i>S_0 of perpendicular molecules</i>
T:	<i>Tilted orientation; $\theta = 45^\circ$</i>
T_N:	<i>Nozzle temperature</i>
TOF:	<i>Time-of-flight</i>
T_s:	<i>Surface temperature</i>
TS:	<i>Transitions state</i>
VASP:	<i>Vienna Ab Initio Simulation Package</i>
V_{av}:	<i>Average velocity of the molecular beam</i>
VdW:	<i>Van der Waals</i>
V_s:	<i>Stream velocity of the molecular beam</i>
XC:	<i>Exchange-correlation</i>
XPS:	<i>X-ray photon spectroscopy</i>

ZPE:	<i>Zero-point energy</i>
ΔE_s:	<i>Distribution width of the energy of the molecular beam</i>
ΔV_s:	<i>Distribution width of the velocity of the molecular beam</i>
Θ:	<i>Incidence angle of the molecular beam</i>
Φ:	<i>Work function</i>