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The sky is made of lava: how lava worlds reveal their interiors through their atmospheres

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2 | LAVATMOS: AN OPEN-SOURCE CHEMICAL EQUILIBRIUM VAPORIZATION CODE FOR LAVA WORLDS

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Abstract

To date, over 500 short-period rocky planets with equilibrium temperatures above 1500 K have been discovered. Such planets are expected to support magma oceans, providing a direct interface between the interior and the atmosphere. This provides a unique opportunity to gain insight into their interior compositions through atmospheric observations. A key process in doing such work is the vapor outgassing from the lava surface. LavAtmos is an open-source code that calculates the equilibrium chemical composition of vapor above a dry melt for a given composition and temperature. Results show that the produced output is in good agreement with the partial pressures obtained from experimental laboratory data as well as with other similar codes from literature. LavAtmos allows for the modeling of vaporization of a wide range of different mantle compositions of hot rocky exoplanets. In combination with atmospheric chemistry codes, this enables the characterization of interior compositions through atmospheric signatures.

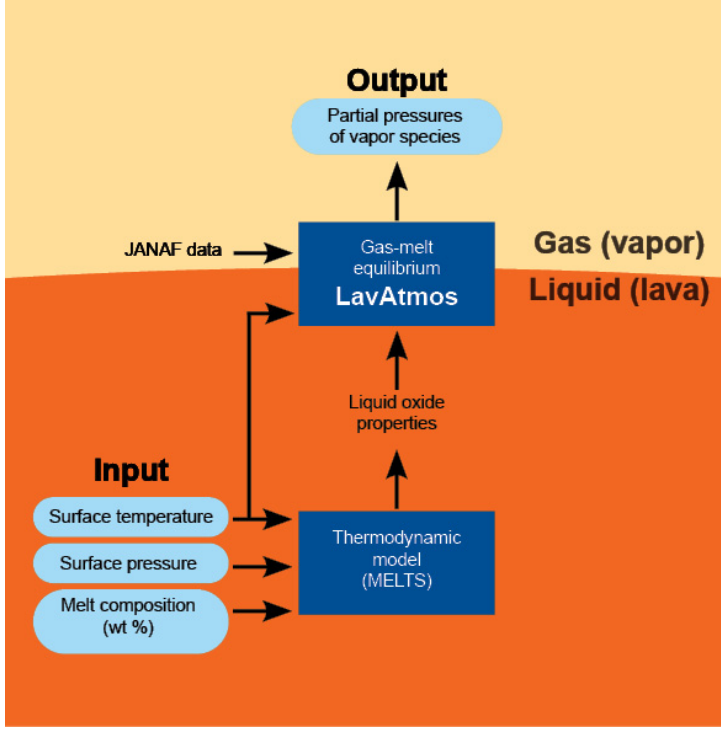


Figure 2.1: LavAtmos schematic: LavAtmos is a code that calculates the composition of the gas above lava at certain surface temperature, pressure, and composition. The MELTS code developed by Ghiorso and co-authors is used to calculate liquid oxide properties for the given input. This is then used alongside data from the JANAF tables in order to perform gas-melt equilibrium calculations. The final output is partial pressure for each of the included vapor species. *Graphical abstract included in publication.*

2.1 Introduction

With an ever-growing catalog of newly discovered exoplanets, we have moved from the discovery phase well into the characterization phase. An emerging category of specific interest is that of the so-called hot rocky exoplanets.

These planets are exposed to extreme stellar irradiation that leads to surface temperatures hot enough to prevent the planet from cooling and creating a crust (Boukaré et al. 2022; Henning et al. 2018), exposing the silicate mantle directly to the atmosphere. Furthermore, since the atmosphere is a direct product of the outgassing from the magma ocean and is in equilibrium with it, the atmospheric compositions of these planets are directly influenced by the interior composition (Dorn & Lichtenberg 2021; Ito et al. 2015; Kite et al.

2016, 2020; Miguel et al. 2011; Nguyen et al. 2020). This provides a unique opportunity to derive interior properties from atmospheric observations. In addition, there is growing evidence that planetesimals, as well as the rocky planets and moons that form from their accretion, were covered by magma early in their evolution (Elkins-Tanton 2012; Greenwood et al. 2005; Hin et al. 2017; Norris & Wood 2017; Schaefer & Elkins-Tanton 2018). Therefore, constraints on the atmospheric products of interior-atmosphere interactions on hot rocky exoplanets might also provide us with a window to the conditions in the early solar system and early Earth (Hirschmann 2012).

Hot rocky planets have been the targets of several observing programs throughout the past several years on a range of different ground- and space-based telescopes (Deibert et al. 2021; Demory et al. 2011; Esteves et al. 2017; Keles et al. 2022). Conclusions drawn from these observations remain uncertain and have yet to give definitive proof of an atmosphere. However, some tentative evidence has been given for K2-141 b (Zieba et al. 2022) and 55 Cnc e (Angelo & Hu 2017; Demory et al. 2016b; Zilinskas et al. 2020, 2021), and the advent of the new generation of telescopes, such as JWST and Ariel, may allow for the characterization of the chemical composition of hot rocky-exoplanet atmospheres in the near future (Ito et al. 2021; Zilinskas et al. 2022).

To know what to look for in such observations and to interpret the data once it arrives, accurate atmospheric models are required. For atmospheres on hot rocky exoplanets, this involves modeling the degassing from lava at the surface of the planet. Since we yet have to gain a good understanding of the possible types of rocky-exoplanet compositions, we do not yet know what kind of compositions we should expect for these melts. Recent work (Brugman et al. 2021; Putirka et al. 2021; Putirka & Rarick 2019) indicates that we should expect a wide range of different possible silicate compositions. Hence, open-source vaporization codes that can work with a wide range of compositions are necessary to enable modeling potential atmospheres as our understanding of hot rocky planets develops.

To date, a limited number of codes have been used to calculate the chemical composition of vapors degassing from lava at a given temperature and the composition of an atmosphere in equilibrium with lava of a given temperature. The MAGMA code (Fegley & Cameron 1987) was written to study the fractional vaporization of Mercury. The same code was used for the study of other solar system bodies (Schaefer & Fegley 2007; Schaefer & Fegley Jr. 2004) and exoplanets (Kite et al. 2016; Miguel et al. 2011; Schaefer & Fegley 2009; Schaefer et al. 2012; Visscher & Fegley 2013). This code makes use of the Ideal Mixing of Complex Components (IMCC) model, developed by Hastie and co-authors (Hastie & Bonnell 1985, 1986; Hastie et al. 1982b), to calculate the activity of the oxide components in the melt.

In more recent years, the MELTS code (Ghiorso & Sack 1995) has seen increased use for modeling the thermodynamics for outgassing codes (Ito et al. 2015, 2021; Jäggi et al. 2021; Wolf et al. 2023). Wolf and coauthors have devel-

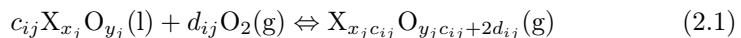
oped the code named VapoRock which has been used to model the early atmosphere of Mercury (Jäggi et al. 2021) and to explore how relative abundances of SiO and SiO₂ could be used to infer the O₂ fugacity of a volatile-depleted mantle (Wolf et al. 2023). The main difference between VapoRock and LavAtmos is the manner in which the O₂ partial pressure is determined. In the discussion (“Discussion” Section), we include a more in-depth comparison of the two codes. Other approaches that calculate the condensate compositions from an initial gas composition, as opposed to calculating vaporization reactions from an existing melt reservoir, have also been developed in the literature (Herbort et al. 2020, 2022).

In this paper, we present a new open-source code, which we named LavAtmos, that calculates the equilibrium composition of a vapor above a melt of a given composition, at a given temperature and at a given melt–vapor interface pressure. As shown in the graphical table of contents, a general overview of the workflow of the code is presented. Just as the abovementioned previous works (Ito et al. 2015, 2021; Wolf et al. 2023), we use MELTS to calculate the oxide component properties of a melt. These properties are then combined with thermochemical data available in the JANAF tables (Chase 1998) to perform gas–melt equilibrium calculations. The oxygen fugacity (f_{O_2}) is derived from the law of mass action, similarly to the approach used for thermodynamic calculations for pure silica and alumina (Krieger 1965) and for the MAGMA code Fegley & Cameron (1987). LavAtmos currently takes 9 oxide species into account (SiO₂, MgO, Al₂O₃, TiO₂, Fe₂O₃, FeO, CaO, Na₂O, and K₂O), 31 different vapor species with corresponding vaporization reactions (shown in Table 1) and is suitable for calculations between 1500 and 4000 K. LavAtmos is written in Python for ease of use and integration with the MELTS Python wrapper named Thermoengine.1 It is released as an open-source code under the GNU General Public License version 3.2 LavAtmos is available on <https://github.com/cvbuchem/LavAtmos>.

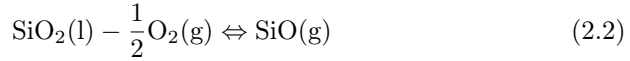
In this paper, we provide an in-depth look at the methods used in “Methodology” Section. We compare its performance to laboratory data and the results calculated by other similar codes where those are available in the public domain in “Validation” Section. Finally, we discuss assumptions made in the method, the advantages and limitations of the code, and highlight a set of potential applications in “Discussion” Section, rounding off with a conclusion in “Conclusion” Section.

2.2 Methodology

In this section, we cover how the partial pressures of included species are calculated. Consider the generalized form of a vaporization reaction of a liquid oxide j , gaseous O₂, and the resulting vapor species i :



X is the cation of the species, x_j the number of cation atoms, y_j the number of oxygen atoms, and c_{ij} and d_{ij} are the stoichiometric coefficients for this reaction. The gaseous species (atmosphere) are indicated using (g) and the liquid species (melt) using (l). As an example, the vaporisation reaction of liquid SiO_2 to form gaseous SiO can be written as:



Assuming that the reaction is in equilibrium, the partial pressure of the vapor species i can be calculated by adhering to the law of mass action as follows:

$$P_{ij} = K_{r_{ij}} a_j^{c_{ij}} P_{\text{O}_2}^{d_{ij}} \quad (2.3)$$

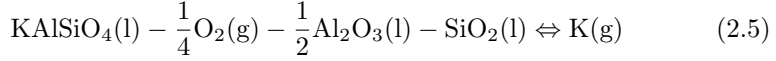
where P_{ij} is the partial pressure of vapor species i as formed from liquid species j , $K_{r_{ij}}$ is the chemical equilibrium constant of the reaction, a_j is the activity of liquid oxide j , and P_{O_2} is the partial pressure of O_2 (also known as the oxygen fugacity $f\text{O}_2$). Some more elaboration on the derivation is shown in the appendix (section 2.A). For the example reaction shown in equation 2.2, the partial pressure of SiO can be determined using:

$$P_{\text{SiO}} = K_r a_{\text{SiO}_2} P_{\text{O}_2}^{-1/2} \quad (2.4)$$

The variables that must be known to calculate the partial pressure of a vapor species are the stoichiometric coefficients c_i and d_i , the chemical equilibrium constant of reaction $K_{r_{i,j}}$ for the vapor species i and the melt oxide j , the chemical activity a_j of the melt oxide j involved in the reaction, and the oxygen partial pressure P_{O_2} . Stoichiometric coefficients are determined by writing out balanced reaction equations (see Table 2.1). The chemical equilibrium constant of each reaction $K_{r_{ij}}$ is determined using the data available in the JANAF tables (Chase 1998).

The activity of the oxides in the melt is determined using the MELTS code. Developed over the course of the past two and a half decades, MELTS has been consistently updated and expanded (Asimow & Ghiorso 1998; Ghiorso & Gualda 2015; Ghiorso et al. 2002; Ghiorso & Sack 1995; Gualda et al. 2012). It performs internally consistent modeling of liquid–solid equilibria in magmatic systems at elevated temperatures and pressures. For a given temperature, pressure, and composition (in terms of oxide weight percentages, see Table 2.2 for example compositions), MELTS is able to calculate the thermochemical properties of the end-member component species that it accounts for in each included phase (Ghiorso & Sack 1995). In the current version of LavAtmos, only the liquid phase is used. Hence, as of yet, partial melting is not taken into account. The properties, specifically the activities, of these end-member components are used by LavAtmos to calculate the energetics of the vaporization reactions of the oxides in the melt. The end-member species included in our MELTS calculations are SiO_2 , Al_2O_3 , TiO_2 , Fe_2O_3 , Fe_2SiO_4 , Mg_2SiO_4 , CaSiO_3 , Na_2SiO_3 ,

and KAlSiO_4 ¹. Due to the fact that some of these species include more than one cation, we write vaporisation reactions that produce a vapor species for one of the cations and liquid species for the others, similarly to the approach taken in work done on condensates from impacts between silicate-rich bodies Fedkin et al. (2006). Using KAlSiO_4 as example, the equilibrium vaporisation reaction may be written as:



Therefore, the partial pressure for K is calculated using:

$$P_{\text{K}} = K_r a_{\text{KAlSiO}_4} P_{\text{O}_2}^{-1/4} a_{\text{Al}_2\text{O}_3}^{-1/4} a_{\text{SiO}_2}^{-1} \quad (2.6)$$

Since Al_2O_3 and SiO_2 are also end-member species, their activities are known. In the cases where a gas species has more than one reaction producing it, such as for Fe (reactions 16 and 18 in Table 2.1) and FeO (reactions 17 and 19 in Table 2.1), the total partial pressure of the species is equal to the sum of the partial pressures calculated for each reaction.

MELTS also requires a pressure as input. For the results shown in this paper, we assume a surface pressure of 0.1 bar. When testing the results for a range of pressure from 10^{-4} up to 100 bar there were no significant differences in the resulting partial pressures. This pressure range includes the atmospheric pressures expected at the lava-atmosphere interface on hot rocky exoplanets (Zilinskas et al. 2022). We assume the silicate compositions are fully molten and for this reason only include the liquid phase in the MELTS calculations. The Python wrapper named Thermoengine is used to interact with the MELTS code (written in C). At the moment of submission of this manuscript, LavAtmos works with MELTS version 1.0.2. This version includes rhyolites-MELTS (Gualda et al. 2012) allowing for the modeling of hydrous silicic systems which may become relevant in future iterations of LavAtmos.

Looking back at equation 2.3, we still need to determine the O_2 partial pressure (P_2). For this, we assume that the metal to oxygen ratio in the vapor is the same as the stoichiometries of the melt oxides, as is also done in the vapor calculations done by Krieger (1965). This gives us a mass balance equation in the form of:

$$n_{\text{O}_2} = - \sum_i \sum_j d_{ij} n_{ij} \quad (2.7)$$

The definition of partial pressure ($P_{ij} = \frac{n_{ij}P}{n}$ for vapor species i produced from liquid species j) can be applied in order to rewrite equation 2.7 as:

$$P_{\text{O}_2} = - \sum_i \sum_j d_{ij} P_{ij} \quad (2.8)$$

¹The mole fraction of each of these end-member species is derived by MELTS from the melt composition given in terms of weight percentages of SiO_2 , MgO , Al_2O_3 , TiO_2 , Fe_2O_3 , FeO , CaO , Na_2O , and K_2O .

Table 2.1: Overview of the 31 vaporization reactions included in LavAtmos - Reaction 1 is the gas to gas reaction of O_2 to O . The rest of the reactions (2-31) describe the vaporization reactions. Each vapor species has a unique reaction that includes a MELTS end-member species, O_2 , and (if necessary) residue liquid metal oxides.

End-member	#	Reactants	Vapor
–	1	$1/2 O_2(g)$	$\rightleftharpoons O(g)$
SiO_2	2	$SiO_2(l) + O_2(g)$	$\rightleftharpoons Si(g)$
	3	$2SiO_2(l) + 2O_2(g)$	$\rightleftharpoons Si_2(g)$
	4	$3SiO_2(l) + 3O_2(g)$	$\rightleftharpoons Si_3(g)$
	5	$SiO_2(l) + 1/2 O_2(g)$	$\rightleftharpoons SiO(g)$
	6	$SiO_2(l)$	$\rightleftharpoons SiO_2(g)$
Al_2O_3	7	$1/2 Al_2O_3(l) + 3/4 O_2(g)$	$\rightleftharpoons Al(g)$
	8	$Al_2O_3(l) + 3/2 O_2(g)$	$\rightleftharpoons Al_2(g)$
	9	$1/2 Al_2O_3(l) + 1/4 O_2(g)$	$\rightleftharpoons AlO(g)$
	10	$Al_2O_3(l) + O_2(g)$	$\rightleftharpoons Al_2O(g)$
	11	$1/2 Al_2O_3(l) + 1/2 O_2(g)$	$\rightleftharpoons AlO_2(g)$
TiO_2	12	$Al_2O_3(l) + 1/2 O_2(g)$	$\rightleftharpoons Al_2O_2(g)$
	13	$TiO_2(l) + O_2(g)$	$\rightleftharpoons Ti(g)$
	14	$TiO_2(l) + 1/2 O_2(g)$	$\rightleftharpoons TiO(g)$
	15	$TiO_2(l)$	$\rightleftharpoons TiO_2(g)$
Fe_2O_3	16	$1/2 Fe_2O_3(l) + 3/4 O_2(g)$	$\rightleftharpoons Fe(g)$
	17	$1/2 Fe_2O_3(l) + 1/4 O_2(g)$	$\rightleftharpoons FeO(g)$
Fe_2SiO_4	18	$1/2 Fe_2SiO_4(l) + 1/2 O_2(g) + 1/2 SiO_2(l)$	$\rightleftharpoons Fe(g)$
	19	$1/2 Fe_2SiO_4(l) + 1/2 SiO_2(l)$	$\rightleftharpoons FeO(g)$
Mg_2SiO_4	20	$1/2 Mg_2SiO_4(l) + 1/2 O_2(g) + 1/2 SiO_2(l)$	$\rightleftharpoons Mg(g)$
	21	$Mg_2SiO_4(l) + O_2(g) + SiO_2(l)$	$\rightleftharpoons Mg_2(g)$
	22	$1/2 Mg_2SiO_4(l) + 1/2 SiO_2(l)$	$\rightleftharpoons MgO(g)$
$CaSiO_3$	23	$CaSiO_3(l) + 1/2 O_2(g) + SiO_2(l)$	$\rightleftharpoons Ca(g)$
	24	$2CaSiO_3(l) + O_2(g) + 2SiO_2(l)$	$\rightleftharpoons Ca_2(g)$
	25	$CaSiO_3(l) + SiO_2(l)$	$\rightleftharpoons CaO(g)$
Na_2SiO_3	26	$1/2 Na_2SiO_3(l) + 1/4 O_2(g) + 1/2 SiO_2(l)$	$\rightleftharpoons Na(g)$
	27	$Na_2SiO_3(l) + 1/2 O_2(g) + SiO_2(l)$	$\rightleftharpoons Na_2(g)$
	28	$1/2 Na_2SiO_3(l) + 1/4 O_2(g) + 1/2 SiO_2(l)$	$\rightleftharpoons NaO(g)$
$KAlSiO_4$	29	$KAlSiO_4(l) + 1/4 O_2(g) + 1/2 Al_2O_3(l) + SiO_2(l)$	$\rightleftharpoons K(g)$
	30	$2KAlSiO_4(l) + 1/2 O_2(g) + Al_2O_3(l) + 2SiO_2(l)$	$\rightleftharpoons K_2(g)$
	31	$KAlSiO_4(l) + 1/4 O_2(g) + 1/2 Al_2O_3(l) + SiO_2(l)$	$\rightleftharpoons KO(g)$

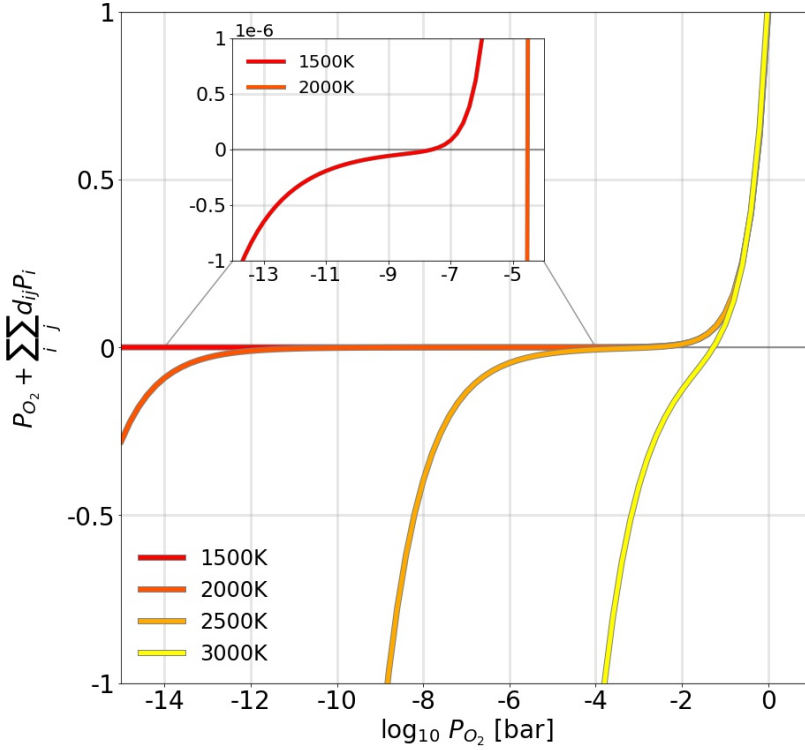


Figure 2.2: Behavior of the mass law function as a function of P_{O_2} at different temperatures - LavAtmos determines the P_{O_2} value by solving the mass law function shown in 2.9. Here we see the behavior of the function for different P_{O_2} values and at different temperatures. An inset is used to show the behavior of the function at a melt temperature of 1500 K due to the difference in scale within which the changes occur relative to higher temperatures.

It is here that we see the importance of tracking the partial pressure of a vapor species i released from a specific reaction with a liquid oxide j . For vapor species that have more than one reaction producing them, such as the aforementioned Fe and FeO, it is necessary to couple the corresponding partial pressures of released Fe/FeO (P_{ij}) to the right stoichiometry (d_{ij}). Substituting equation 2.3 and setting one side to 0 gives:

$$0 = P_{O_2} + \sum_i \sum_j d_{ij} K_{r_{ij}} a_j^{c_{ij}} P_{O_2}^{d_{ij}} \quad (2.9)$$

Finding the P_{O_2} for which equation 2.9 holds, allows one to calculate the partial pressures of all other vapor species (using equation 2.3) while ensuring that the mass action law is being upheld. In Figure 2.2, the behavior of the right side of equation 2.9 as a function of P_{O_2} is shown for different melt temperatures. The

higher the melt temperature, the greater the P_{O_2} value at which equation 2.9 holds true (where the y-axis is zero in Figure 2.2). If a P_{O_2} value is used that corresponds to a (non-zero) negative value on the y-axis at a given temperature, this implies that not all of the O_2 released during vaporization reactions is included in the vapor. If a P_{O_2} value is used that corresponds to a (non-zero) positive value on the y-axis at a given temperature, this implies that extra oxygen would have to be added to the system in order to reach the calculated equilibrium.

In LavAtmos, equation 2.9 is solved for P_{O_2} using the Scipy optimization function `fsolve`². A vaporization reaction is defined for each vapor species included in LavAtmos (as shown in Table 2.1).

2.3 Validation

Table 2.2: Melt compositions used for Figures REFERENCE - The compositions given are for SRM glass (Hastie & Bonnell 1985), Illite (Hastie et al. 1982b), bulk silicate earth (BSE) (Palme & O'Neill 2003), a CB chondrite chondrule with bulk Na and K mass balanced for chondrules to fit bulk meteorite iron-silicate ratio (Weisberg et al. 1990, 2001), an analog to enstatite chondrites (EH4) (Wiik 1956), and Northern Smooth Planes (NSP) lava for both source and lava (Namur et al. 2016; Nittler & Weider 2019).

Oxide wt%	SRM glass	Illite	BSE	CB	EH4	NSP source	NSP lava
SiO ₂	71.390	60.200	45.400	50.700	62.730	53.670	58.700
MgO	0.270	2.100	36.760	36.900	30.240	36.890	13.900
Al ₂ O ₃	2.780	26.000	4.480	4.600	2.580	4.750	13.800
TiO ₂	—	—	0.210	—	—	—	—
Fe ₂ O ₃	0.040	4.400	—	—	—	—	—
FeO	—	—	8.100	3.500	—	0.020	0.040
CaO	10.750	—	3.650	3.300	1.990	2.260	5.810
Na ₂ O	12.750	0.200	0.349	0.190	1.710	1.970	7.000
K ₂ O	2.020	7.400	0.031	0.050	0.200	0.050	0.200
Total	100.000	100.300	98.980	99.240	99.450	99.610	99.450

In this section LavAtmos results are compared to both experimental data and the results of other similar codes. In Figure 2.3 we compare the output of the LavAtmos code to that of the data gathered by Hastie and co-authors (Hastie et al. 1982a; Hastie & Bonnell 1985), with output from MAGMA (Schaefer & Fegley Jr. 2004), and output from VapoRock (Wolf et al. 2023) where available. Note that for the laboratory data no uncertainties were provided by the authors.

²<https://docs.scipy.org/doc/scipy/reference/generated/scipy.optimize.fsolve.html>

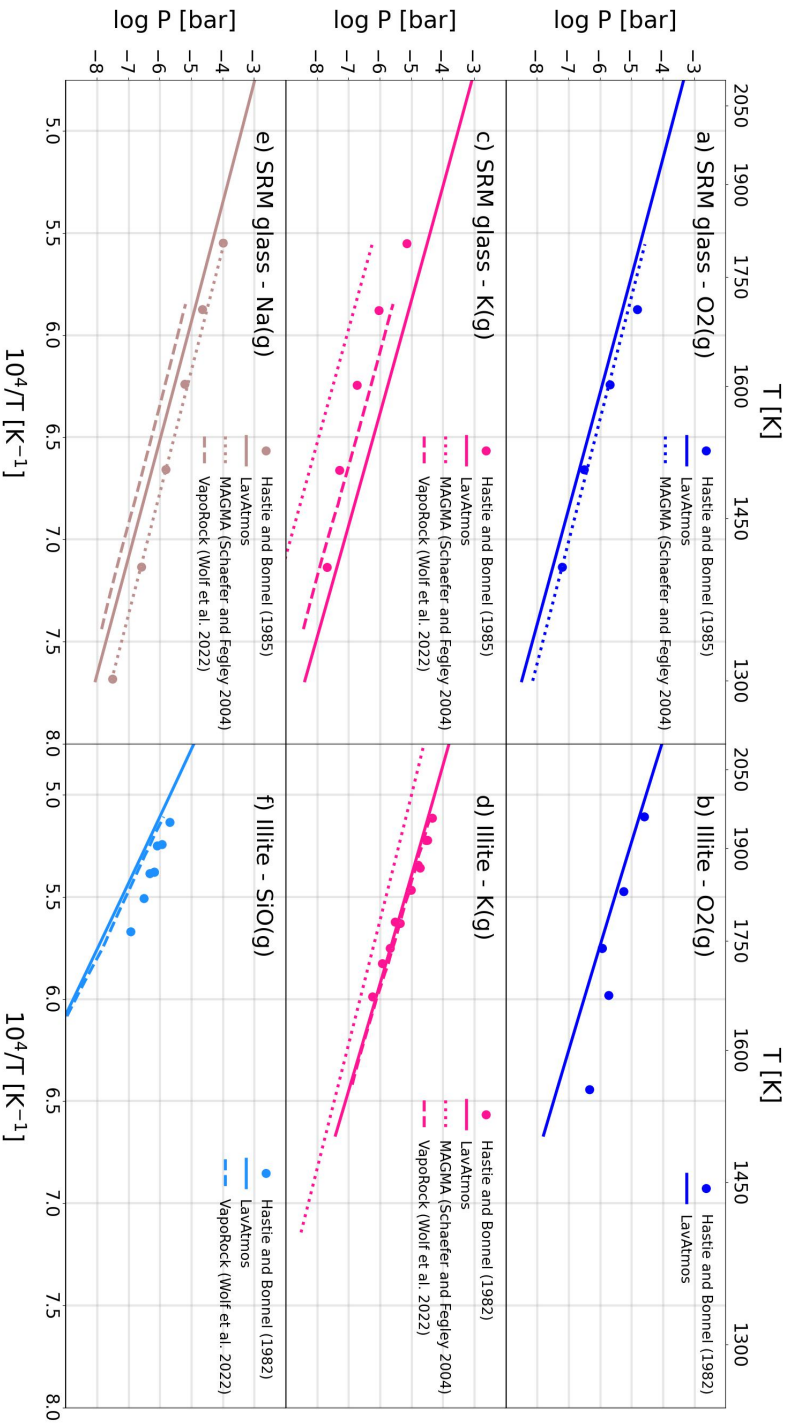


Figure 2.3: Comparison between data from Hastie and co-authors and code output - Each plot shows the partial pressure of a vapor species as experimentally measured by Hastie and co-authors Hastie et al. (1982b); Hastie & Bonnell (1985) (dots), calculated using LavAtmos (solid line), calculated using MAGMA (dotted line) Schaefer & Fegley Jr. (2004), and as calculated using VapoRock (dashed line) ?. Plots a, c, show the results for SRM glass, while plots b, d, and f show the results for illite. The compositions of the melts are shown in Table ???. Not all species had available data from all codes, hence some of the codes are missing in some of the plots.

In the panels on the left side of the figure (a, c, and e), the partial pressures of O_2 , K, and Na above a synthetic (SRM) glass are shown. The bulk composition of this glass is given in Table 2.2. For O_2 we see that the results of both MAGMA and LavAtmos align well with the laboratory data. VapoRock is not compared in the O_2 plots due to it requiring fO_2 as an input (more on this below), for the calculation of the partial pressures of the other species VapoRock used the values given by Hastie and co-authors as input. For K the partial pressure predicted by LavAtmos and VapoRock differ by about a factor 2, with VapoRock lying closer to the experimental data. The partial pressures predicted by MAGMA lie about one order of magnitude below that of VapoRock. For Na it appears that MAGMA output falls closer to the laboratory data while LavAtmos and VapoRock slightly underestimate the Na partial pressures.

For Illite (see Table 2.2 for its bulk composition), on the right side of Figure 2.3 (b, d, and f), we see that the predicted partial pressure for O_2 is in agreement with the data at temperatures of 1750 K and above. We also see a much better agreement between the LavAtmos output and the data for K than we did for SRM glass, with VapoRock giving almost identical values while MAGMA tends to underestimate the K partial pressure by about a factor two at lower temperatures (~1550K) and about an order of magnitude at higher temperatures (above 2000K). The partial pressure for SiO predicted by LavAtmos and VapoRock is very similar with both underestimating the experimental data by about a factor two.

Overall we find good agreement between the calculated partial pressures and the experimental data. We do recognize that this is a limited amount of data to which we are able to compare our output and that the lack of experimental uncertainties makes it difficult to quantify the results. The difference in the values predicted for the K partial pressures by MAGMA and LavAtmos is due to the difference in the thermodynamic models used by the two codes. As explained in previous work (Ito et al. 2015) where similar discrepancies with the MAGMA code were found in their MELTS-based calculations, the IMCC model (used by MAGMA to calculate activity values) is calibrated by experimental results on synthetic high- K_2O melts. This differs from MELTS, which is calibrated on experimental data drawn from natural, low- K_2O melts. This could also help explain why the K partial pressures predicted by LavAtmos appear to be closer to the experimental data for Illite (a naturally occurring composition) than for the SRM glass (synthetic composition). Another reason for the difference in partial pressures could be the much higher percentage of CaO and Na_2O in synthetic melts than in natural compositions. This would also explain why VapoRock is showing similar behavior to that of LavAtmos. The difference between the values calculated by LavAtmos and VapoRock are likely due to the different treatments of fO_2 . We will elaborate further upon this below.

To illustrate the use of LavAtmos and compare with other codes, we mod-

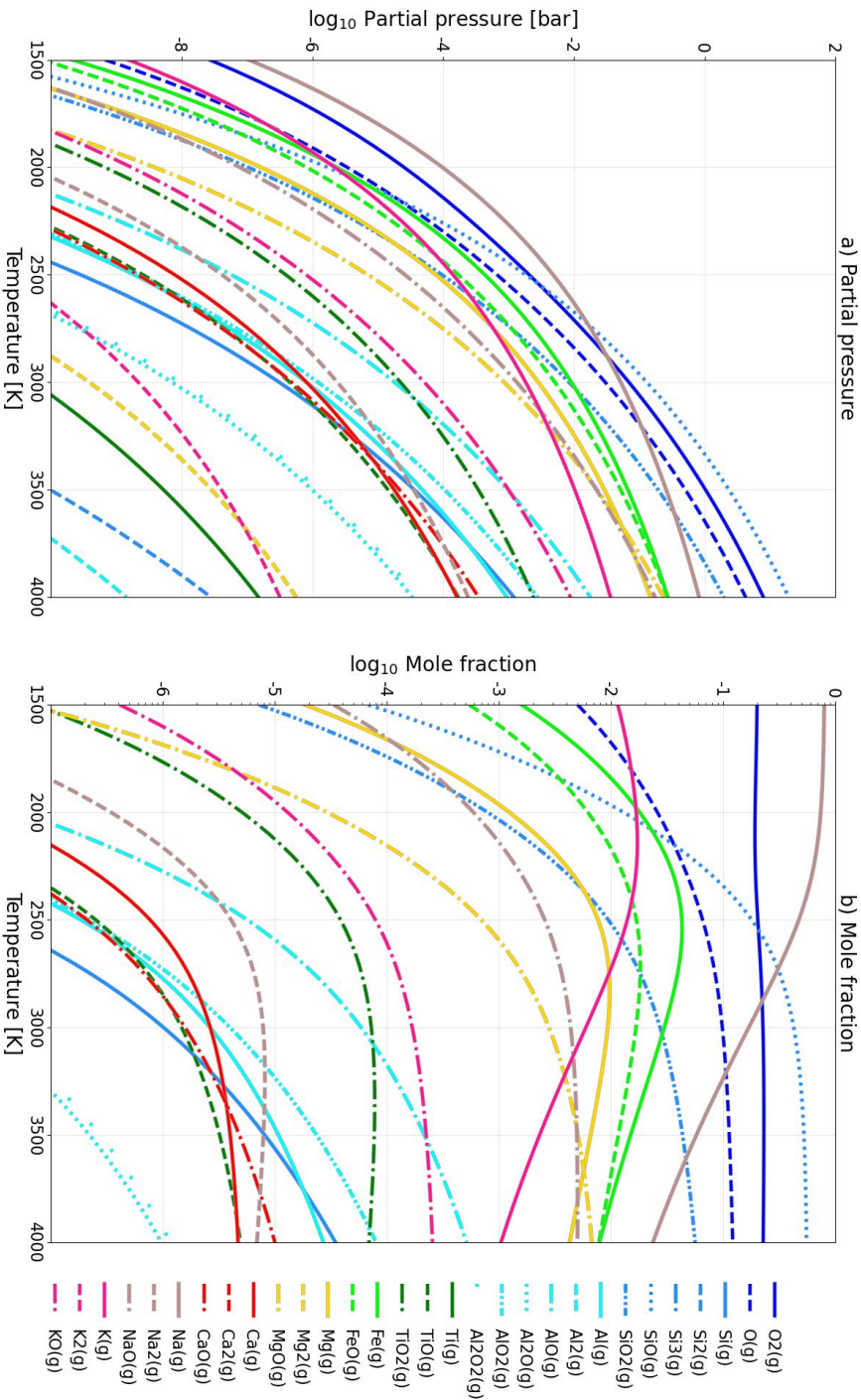


Figure 2.4: LavAtmos output for BSE composition between 1500 and 4000K - The left panel (a) shows the partial pressures and the right panel (b) shows the mole fractions (partial pressure divided by the total pressure) of the included vapor species. These calculations were done for a BSE composition (see Table 2.2). A surface pressure of 0.01 bar was assumed.

eled the vapour output for a bulk silicate earth (BSE) composition (see Table 2.2) at temperatures between 1500 K and 4000 K. The resulting partial pressures and mole fractions are shown in Figure 2.4. The dominant species up to about 2700 K is Na vapor, at which point SiO vapor takes over as the dominant species. O₂ and O have a strong presence as well throughout the majority of this temperature range. The general trend is that the more volatile species, such as Na, K, Fe, and Mg make up a larger fraction of the total gas pressure at lower temperatures than they do at higher temperatures. The opposite is true for the less volatile species, such as Si, Al, and Ti.

In Figure 2.5 we can see a comparison of a selection of the calculated mole fractions for a BSE composition (Table 2.2) as calculated by LavAtmos, Ito et al. (2015), and MAGMA (Visscher & Fegley 2013). The BSE composition used for the Ito et al. (2015) results differs slightly (a few decimal percentages) from the BSE composition used for the other calculations. Tests using LavAtmos however show that the difference is so small that the effects are negligible. Due to lack of details on the Ito et al. (2015) gas-melt equilibrium calculations we are not able to comment on differences in methodology. In Figure 2.6 we also show the corresponding total pressures. The partial pressures calculated by Ito et al. (2015) (a in Figure 2.5) are very close to the partial pressures calculated by LavAtmos. This is likely due to the fact that both codes make use of MELTS (Ghiorso & Sack 1995) to calculate the thermodynamic properties of the melt. In addition to potential differences in the approach, our calculations include a wider range of vapor species. When comparing LavAtmos to the MAGMA code (Fegley & Cameron 1987) data in recently published work (Visscher & Fegley 2013) (c in Figure 2.5) the most significant difference is between the calculated K partial pressures, as we also saw in Figure 2.3. As explained earlier in this section, this is due to the difference in calibration of the thermodynamic models used for the modeling of the liquid oxide activities (MELTS for LavAtmos and IMCC for MAGMA).

The total vapor pressures calculated for each code (Figure 2.6) show similar results. At lower temperatures it appears that LavAtmos and Ito et al. (2015) predict a lower total pressure than MAGMA. For higher temperatures MAGMA predicts a lower total pressure compared to the other two codes. Note that the mole fractions of the vapor species are calculated by normalizing the partial pressures by the total pressure calculated by the codes. Hence, if two codes have the same mole fractions as output for certain vapor species, this may not translate to the same partial pressure values depending on if the calculated total pressures are the same or not.

In Figure 2.7 we compare the output of LavAtmos (solid lines) with that of VapoRock (dashed lines) for 4 different compositions (see Table 2.2) as published in the work by Jäggi et al. (2021). VapoRock is a gas-melt equilibrium code that is also based on the thermodynamics calculated by MELTS. One of the ways in which this is apparent is the similar values calculated for the K partial pressures (pink) for all compositions. The greatest difference between

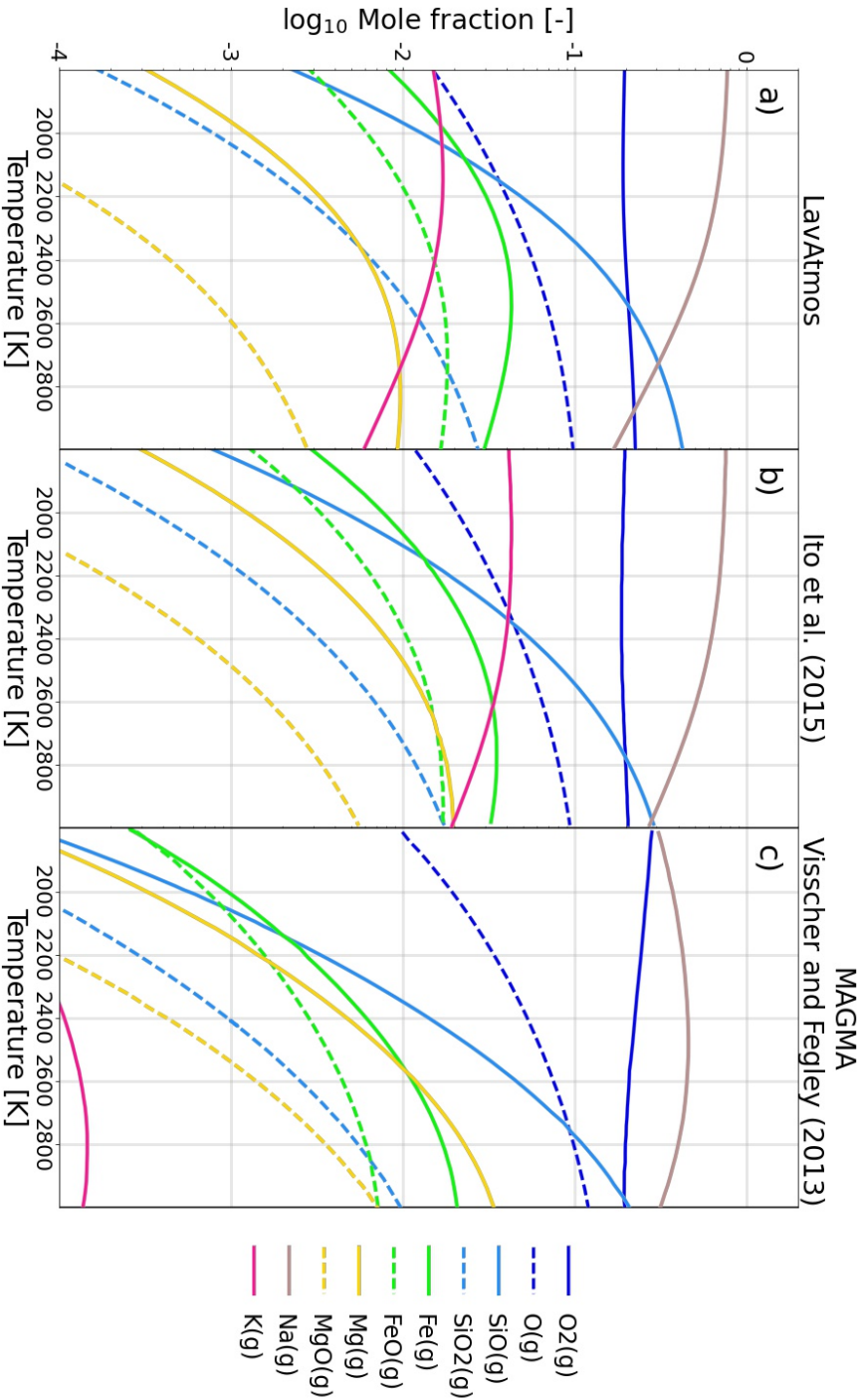


Figure 2.5: Comparison of mole fractions above a BSE melt as calculated by different codes - A selection of vapor species was made based on the data available in other works. The results shown are from LavAtmos (a), Ito et al. (2015) (b) and MAGMA as published in Vischer & Fegley (2013) (c).

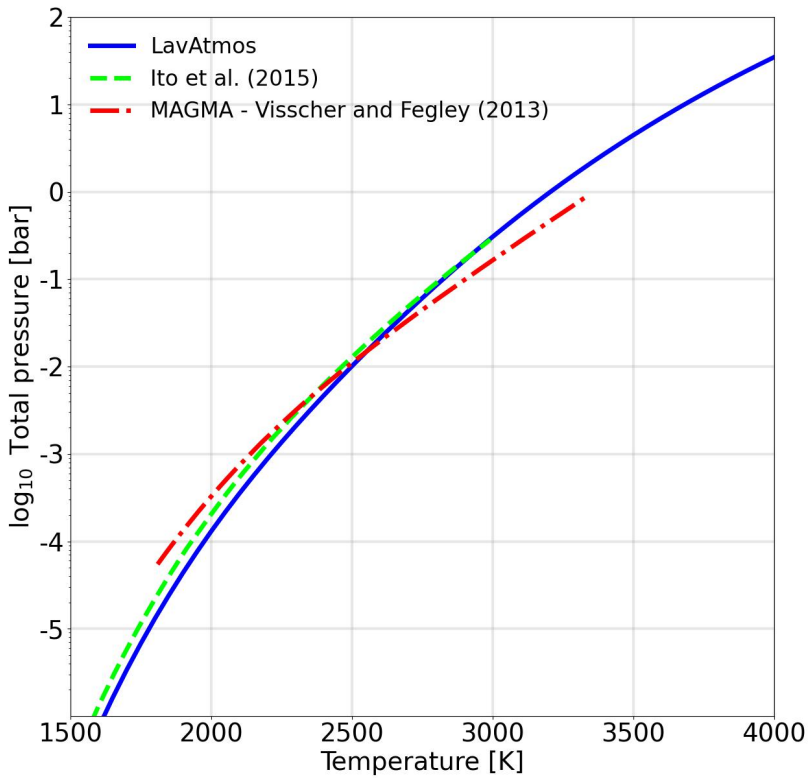


Figure 2.6: Total vapor pressures above a BSE melt as calculated by different codes - LavAtmos (solid dark-blue line), Ito et al. (2015) (dashed lime line), and MAGMA as published in Visscher & Fegley (2013) (dash dotted red line)

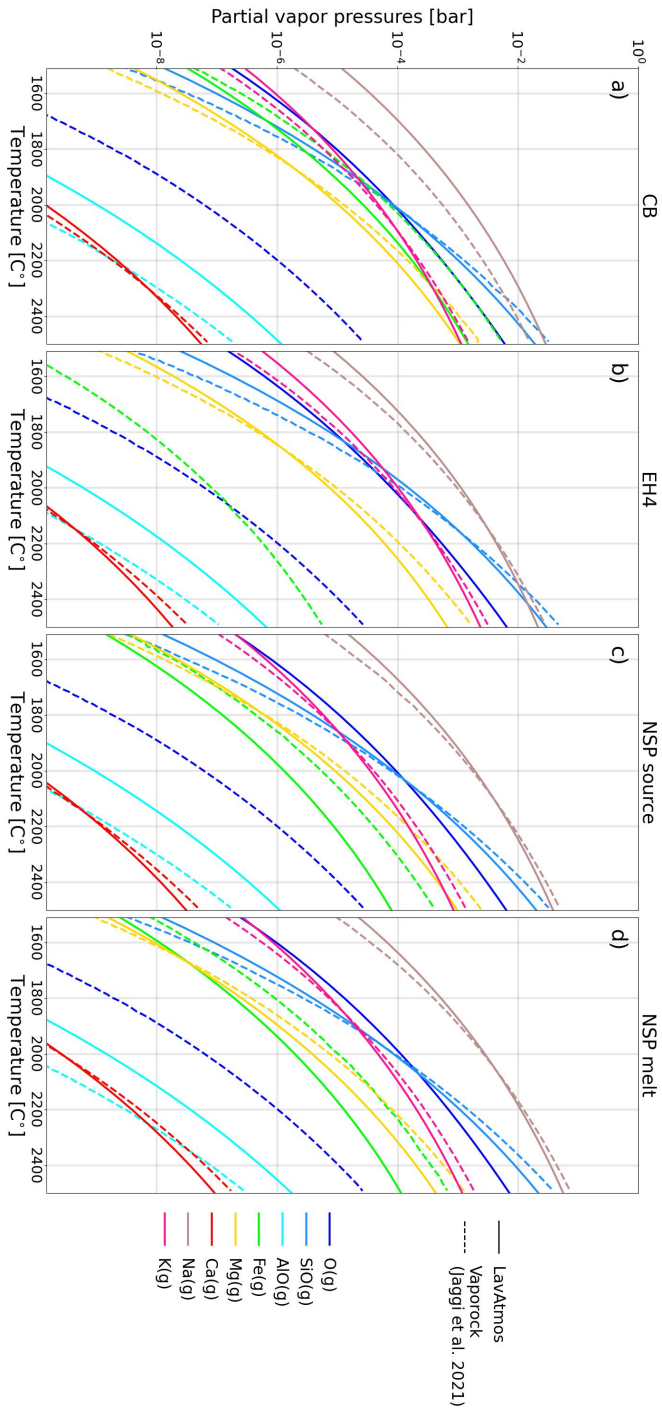


Figure 2.7: LavAtmos and VapoRock comparison - LavAtmos (solid lines) and VapoRock (Jäger et al. 2021) (dashed lines) partial pressure output for a selection of species for melt compositions CB (a), EH4 (b), NSP source (c), and NSP melt (d). See Table 2.2 for the compositions.

the results of the two codes is seen in the values calculated for oxygen partial pressure. This can be attributed to the main difference in the approach taken by VapoRock and LavAtmos: the manner in which the oxygen fugacity of the system is determined.

For VapoRock, the oxygen fugacity of the melt is one of the user inputs (alongside temperature, pressure, and melt composition). For the results shown in Figure 2.7, and oxygen fugacity of IW-1 was used (Jäggi et al. 2021) (1 log unit less than the oxygen fugacity of the known iron-wustite buffer). This provides the user the option to fix the oxygen fugacity value to a certain value, or to test the influence of different oxygen fugacity values on the outgassing of the other species.

LavAtmos, however, does not take oxygen fugacity as an input and instead calculates it internally through the dual constraints of mass action and mass balance, similarly to MAGMA (Fegley & Cameron 1987; Schaefer & Fegley Jr. 2004), as shown in section 2.2 (equations 2.3 and 2.7). This allows the user to calculate a value for the oxygen fugacity based on the assumed congruent vaporization of the oxide species in the melt. As shown in Figure 2.3, this method is able to reproduce the oxygen fugacity values seen in experimental data. The same has been shown repeatedly for MAGMA (Fegley & Cameron 1987; Schaefer & Fegley Jr. 2004; Sossi & Fegley 2018).

VapoRock makes use of the thermodynamic data from the work done by Lamoreaux and co-authors (Lamoreaux & Hildenbrand 1984; Lamoreaux et al. 1987) and from the JANAF tables (Chase 1998), while LavAtmos sources its thermodynamic data from the JANAF tables only. Lamoreaux and co-authors found that the thermodynamic values that they found were in good agreement with those found in the JANAF database. Hence, we expect that this difference in approach does not have a significant impact on the results. VapoRock includes degassing of chromium species (Cr), while LavAtmos does not. We expect that future iterations of LavAtmos will also include this species.

2.4 Discussion

LavAtmos assumes that thermochemical equilibrium is reached. Hence, whenever applying this code to a system one should keep in mind whether or not the timescales of the system allow for chemical equilibrium to be reached (at least on a local level). Due to the high temperatures (> 1500 K) for which the code is meant, it can generally be safely assumed that an equilibrium state is reached at the surface of hot rocky exoplanets (Miguel et al. 2011; Ito et al. 2015).

The value of the oxygen fugacity is derived from the law of mass balance (see section 2.2). This assumes that the ratio of O atoms to cations must be the same in the out-gassed vapor as in the melt. This is a valid assumption if we assume that the system is dominated by the melt and that besides the vapor released from the lava, no other atmosphere is present. Figure 2.3 shows

that P_{O_2} values calculated using this method are in agreement with laboratory data.

The temperature range within which it is advised to use LavAtmos is between 1500 and 4000 K. For our method we assume that the planetary surface is fully molten (above liquidus temperature). The temperature at which a surface rock composition is fully molten varies strongly with composition and pressure as shown in the work by Hirschmann and co-authors (Hirschmann 2000) who illustrate the effects on melting behavior of typical terrestrial peridotitic mantle compositions by changing the alkali content or Mg/Fe ratio of the composition. It is estimated that the liquidus of terrestrial mantle peridotite at 1 atmosphere is around 1700 K based on extrapolation of higher pressure data (Zhang & Herzberg 1994). The liquidus temperature for the Earth's mantle at 1 atmosphere is estimated to be around 2000 K (Andrault et al. 2011). The liquidus of a basalt is down at around 1500 K (Cohen et al. 1967). Recently, it has been shown that elevated oxygen partial pressure can lower the basalt liquidus further (Lin et al. 2021). We therefore advise future users of LavAtmos to look into the estimated liquidus temperature of the composition for which to calculate degassing if a temperature below 2000 K is used. Below the liquidus temperature, crystals will form changing the composition of the magma away from the initial bulk composition. Thanks to the possibility of including mineral phases in MELTS (Ghiorso & Sack 1995), this limit may be removed in future work, with LavAtmos incorporating information about the thermodynamic properties of the melt as it cools, and changes in melt composition due to the crystallization of minerals.

As mentioned in section 2.2, the output of LavAtmos is pressure independent up to about 100 bar. Therefore, if the vaporization calculations are performed for a lava surface pressure below 100 bar, the default value of 0.1 bar can be used. If a surface pressure higher than 100 bar is expected, this should be specified to when calling LavAtmos. As mentioned before, higher pressures may change the liquidus temperature of the melt, something which should be kept in mind when selecting the temperature range for which the code is run.

The validation (section 2.3) with respect to laboratory data relies entirely on the data gathered by Hastie and co-authors (Hastie et al. 1982b; Hastie & Bonnell 1985). The first issue with this is the fact that no uncertainties were included in the published lab data. This has as consequence that the comparisons are mostly qualitative in nature. This makes it difficult to judge which of the tested codes is more representative of the experimental data. There are also only a limited number of vapor species for which partial pressure data is available. Besides the species for which data was published by Hastie and colleagues (O_2 , K, Na and SiO), no suitable experimental data appears to be available for the remaining species to be compared to. We are aware of experimental data on the vaporization of two lunar basalt samples (de Maria et al. 1971). Nevertheless, due to uncertainties about changes in the bulk composition of these samples during the evaporation measurements, as well as the presence of iron

metal and reported interaction of oxygen with the capsule used to hold the samples (de Maria et al. 1971), we do not consider it appropriate to include these data in our comparison. This makes it difficult to judge the accuracy of the code in predicting partial pressures for which we do not have vaporization data.

Another source of uncertainty is that we are only able to compare our data to two different types of melt compositions (SRM glass and illite, see Table 2.2). More extensive lab measurements of gas partial pressures above complex melts of natural compositions are sorely needed.

2.5 Conclusion

We developed a open-source code for gas-melt equilibrium calculations that can be used to predict the composition of a vapor above a melt of a given composition and temperature. Thermochemical values are drawn from the JANAF tables (Chase 1998). The geothermodynamic code MELTS (Ghiorso & Sack 1995) is employed to model the thermodynamic activities of oxide species in the melt, which provides an approach based on more recent work than the IMCC model used by the MAGMA code. The P_{O_2} , necessary for determining the partial pressures of all other species, is determined self-consistently using the dual constraints of the laws of mass-action and mass-balance, which sets this code apart from VapoRock, which takes P_{O_2} as an input. We have shown that the output of the code is in line with the available lab data as well as other commonly used codes from literature. LavAtmos is applicable under the condition that one can assume that chemical equilibrium is reached, that no volatiles (e.g. H_2O , CO_2 , N_2) are included in the melt, and within a temperature range of 1500 K and 4000 K.

With the new generation of telescopes allowing exoplanet research to push towards observations of the atmospheres of smaller planets (JWST, Ariel, ELTs), we hope that an open source lava-degassing code will be of use when interpreting spectra of planets potentially supporting lava oceans. Future work will involve including more volatile species so as to be applicable to a wider range of possible exoplanets.

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Data Availabilty

The code presented in this work is openly available on Github at <https://github.com/cvbuchem/LavAtmos.git>.

Appendix

2.A Deriving the partial pressure equation

For an equilibrium reaction of the type shown in equation 2.1, the relation between the chemical potentials (μ) of the liquid oxide species (j) and vapor species (i) involved can be written as:

$$c_{ij}\mu_j + d_{ij}\mu_{\text{O}_2} = \mu_i \quad (2.10)$$

The chemical potential of a species is given by:

$$\mu = \mu^\circ + RT \ln(a) \quad (2.11)$$

Where μ° is the standard chemical potential of the species, a is the activity³, and R the gas constant. For ideal gasses at low pressures we assume that the activity is equal to the partial pressure of the gas. The chemical equilibrium constant for a reaction r at temperature T is given by:

$$\ln K_r = -\frac{\Delta_r G^\circ(T)}{RT} \quad (2.12)$$

Where $\Delta_r G^\circ$ is the difference in standard chemical potentials of the products and the reactants. Using this, equation 2.10 can be rewritten as:

$$c_{ij}(\mu_j^\circ + RT \ln(a)) + d_{ij}(\mu_{\text{O}_2}^\circ + RT \ln(P_{\text{O}_2})) = \mu_i^\circ + RT \ln(P_i)$$

Equation 2.12 can be rewritten as:

$$-\Delta G^\circ = \mu_i^\circ - c_{ij}\mu_j^\circ - d_{ij}\mu_{\text{O}_2}^\circ = RT \ln(K_{r_{ij}})$$

Which can then be substituted into the previous equation, giving us:

$$RT \ln(P_i) = RT \ln(K_{r_{ij}}) + c_{ij}RT \ln(a_j) + d_{ij}RT \ln(P_{\text{O}_2})$$

Which simplifies to equation 2.3.

³Activity is a term that indicates the effective concentration of a species. One can consider it as a measure of "how much" of a certain species is available for reaction.



