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Role of near-surface environment in tuning electrochemical CO₂ reduction reaction and H₂ evolution reaction

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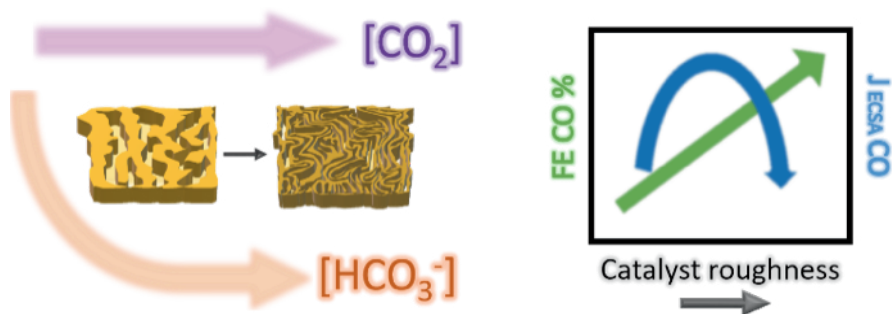
Chapter 5

Effect of pore diameter and length on electrochemical CO_2 reduction reaction at nanoporous gold catalysts



Abstract

In this work, we employ differential electrochemical mass spectrometry (DEMS) to track the real-time evolution of CO at nanoporous gold (NpAu) catalysts with varying pore parameters (diameter and length) during the electrochemical CO₂ reduction reaction (CO₂RR). We show that due to the increase in the local pH with increasing catalyst roughness, NpAu catalysts suppress bicarbonate-mediated hydrogen evolution reaction (HER) compared to a flat Au electrode. Additionally, the geometric current density for CO₂RR increases with the roughness of NpAu catalysts, which we attribute to the increased availability of active sites at NpAu catalysts. Together, the enhancement of CO₂RR and the suppression of competing HER results in a drastic increase in the Faradaic selectivity for CO₂RR with increasing pore length and decreasing pore diameter, reaching near 100% Faradaic efficiency for CO in the most extreme case. Interestingly, unlike the geometric current density, the specific current density for CO₂RR has a more complicated relation with the roughness of the NpAu catalysts. We show that this is due to the presence of Ohmic drop effects along the length of the porous channels. These Ohmic drop effects render the pores partially electrocatalytically inactive and hence, they play an important role in tuning the CO₂RR activity on nanoporous catalysts.



This chapter is based on the submitted manuscript:

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5.1. Introduction

Electrochemical reduction of CO₂ (CO₂RR) can be used to achieve a carbon-neutral energy cycle wherein carbon-based fuels can be produced with net zero emissions by using renewable electricity.¹⁻³ However, at present, the economic feasibility of this reaction remains an issue, primarily due to its low energy efficiency in commonly employed bicarbonate electrolytes, especially at high current densities.⁴⁻⁵ Consequently, significant research efforts have been made towards optimizing the catalyst design as well as the reaction process conditions for achieving better CO₂RR selectivities.⁶⁻⁹ Recently, meso- or nanoporous electrocatalyst materials have emerged as an interesting strategy to tune the selectivity of CO₂RR. To this end, most of the research has focused on the effect intrinsic properties of nanoporous structures (such as high, potentially active surface area, high relative density of stepped sites and grain boundaries) have on CO₂RR activity.¹⁰⁻¹⁵ However, the effect of porosity towards tuning the local reaction environment by regulating the near-surface concentration of different reactive species (such as CO₂, HCO₃⁻ and OH⁻) has received less scrutiny. Given that local concentration gradients play a significant role in tuning the competition between CO₂RR and the parasitic hydrogen evolution reaction (HER) on flat polycrystalline electrodes,^{8, 16-17} it can be expected that the mass transport limitations introduced by confinement effects in and around the nanoporous channels will also affect the competition between these two reactions.¹⁸⁻²⁰

In this respect, some recent studies have indeed emphasized the importance of local diffusional gradients in tuning the CO₂RR activity on nanoporous electrodes.²¹⁻²⁵ In general, it has been shown that with the increasing roughness/thickness of the nanoporous catalysts, the local pH at the surface also increases, resulting in the suppression of bicarbonate-mediated HER reaction ($\text{HCO}_3^- + 2e^- \rightarrow \text{H}_2 + 2\text{CO}_3^{2-}$). However, conflicting results have been obtained on the effect of local concentration gradients in tuning the rate of CO₂RR reaction on nanoporous catalysts. Surendranath and co-workers have shown the formation rate of CO on nanoporous gold (NpAu) is largely independent of the catalyst's pore dimensions, which they attributed to negligible concentration gradient for CO₂ between the bulk environment to the surface due to its slow acid-base equilibria in the electrolyte.^{22, 26} However, in the case of nanoporous silver catalysts, the same authors observe that increasing catalyst layer thickness results in higher partial current densities for CO₂RR.²¹ Meanwhile, Atwater and co-workers have observed a similar enhancement in the rate of CO₂RR in the case of NpAu catalysts, which they attributed to the presence of grain boundaries on nanoporous electrodes.²⁴ On the other hand, in a similar study, Cheng and co-workers have found the opposite effect of catalyst size on CO₂RR and attributed this to the decreasing solubility of CO₂ with increasing local pH at the surface.²⁵

One of the complications in unambiguously elucidating the role of nanoporous catalysts in tuning the activity of the electrocatalytic reactions (CO₂RR and HER) arises from the ancillary participation of homogenous acid-base equilibria ($\text{CO}_2(\text{aq.}) + \text{OH}^- \leftrightarrow \text{HCO}_3^-$; $\text{HCO}_3^- + \text{OH}^- \leftrightarrow \text{CO}_3^{2-} + \text{H}_2\text{O}$)

in regulating the local concentration gradients at the catalyst surface.²⁷⁻²⁹ Consequently, different concentration gradients can be generated near similar catalysts due to the convoluted interplay between the homogenous reactions, mass transport, and the heterogeneous reactions, especially in the case of long-term bulk electrolysis measurements.³⁰ Hence, in order to more reliably disentangle these effects, it is imperative to track the real-time evolution of the reaction products as well as the local concentration gradients by using an online detection technique. Moreover, cell geometries with well-defined mass transport conditions are also desirable in order to avoid the time-dependent concentration polarization effects.³¹ To circumvent these issues, we have recently developed two online methods, namely, a gold ring based rotating ring disk electrode (RRDE) technique and an online differential electrochemical mass spectrometry (DEMS) technique based on a dual thin-layer flow cell, for the quantitative detection of CO during CO₂RR under well-defined mass transport conditions.^{8, 17, 32-33}

In this paper, we present a study on the role of pore parameters (diameter and length) in tuning the competition between CO₂RR and HER on nanoporous Au (NpAu) catalysts by using our DEMS technique. We find that with decreasing pore diameter and increasing pore length the Faradaic selectivity for CO₂RR increases dramatically (up to ~100%), even under sub-optimal reaction conditions (0.5 atm. CO₂ in 0.1 M NaHCO₃ plus 0.4 M NaClO₄). We show that this enhancement arises due to two factors: (i) the suppression of bicarbonate mediated HER reaction with increasing local pH at the catalyst surface and (ii) the enhancement of CO₂RR at NpAu catalysts due to the increased availability of active sites with increasing catalyst roughness. Moreover, we show that the Ohmic drop effects across the porous channels play an important role in tuning the obtained specific current densities for CO₂RR as they render the pores partially inactive towards electrocatalysis. Hence, by using the online DEMS technique, we elucidate the convoluted interplay between catalyst geometry effects and local concentration gradient effects in tuning the activity/selectivity of CO₂RR at nanoporous catalysts.

5.2. Experimental section

5.2.1. Chemicals

The electrolytes were prepared from H₂SO₄ (98% by wt. solution, EMSURE, Merck), NaHCO₃ (≥ 99.7%, Honeywell Fluka), NaClO₄ (99.99%, trace metals basis, Sigma-Aldrich), NaH₂PO₄ (≥ 99.0%, EMSURE, Merck), Na₂HPO₄ (≥ 99.0%, EMSURE, Merck) and Ultrapure water (MilliQ gradient, ≥18.2 MΩcm, TOC < 5 ppb). CO (4.7 purity, Linde), Ar (6.0 purity, Linde) and CO₂ (4.5 purity, Linde) were used for purging the electrolytes. To prepare the nanoporous Au (NpAu) catalysts, KAu(CN)₂ (99.95 %, trace metals basis, Sigma-Aldrich), KAg(CN)₂ (Sigma-Aldrich), Na₂CO₃ (≥ 99.5%, Sigma-Aldrich) and HClO₄ (60% by wt. solution, EMSURE, Merck) were used.

5.2.2. NpAu catalyst synthesis and characterization

Before each experiment, the substrate electrode (polycrystalline Au) was mechanically polished on Buehler micro-polishing cloth (8 inches) with decreasing sizes of diamond polishing suspension, namely, 3 μ m, 1 μ m and 0.25 μ m. Next, it was sonicated in ultrapure water and acetone for 10 minutes to remove any organic/inorganic impurities. Thereafter, the NpAu catalyst layer was formed on the substrate electrode by following the electrochemical alloying-dealloying procedure outlined previously.³⁴⁻³⁶ Briefly, a binary alloy solution of 5:1 ratio of Ag/Au was prepared with KAg(CN)₂ and KAu(CN)₂ in 0.25 M Na₂CO₃ buffer solution and it was potentiostatically electrodeposited onto the flat Au substrate electrode at -1.2 V (vs Ag/AgCl) for 9 min. Thereafter, Ag was de-alloyed from the film by electrochemically cycling (20 cycles) the electrode in 1 M HClO₄ solution between 1.05 V to 1.2 V (vs Ag/AgCl) at 5 mVs⁻¹. The as-prepared NpAu catalyst (referred to as "NpAu4") was thermally coarsened to obtain varying pore diameter and height, at 523.15 K, for 120 sec., 60 sec. and 30 sec. to form NpAu1, NpAu2 and NpAu3, respectively. To avoid faulty activity measurements due to any residual CO₂ in the pores, a new catalyst layer was prepared for each measurement. To determine the roughness factor of the catalysts, electrochemical characterization was performed in 0.1 M H₂SO₄ (0.05 V to 1.75 V vs RHE at a scan rate of 50 mV s⁻¹) to obtain the electrochemically active surface area (ECSA) of the catalyst by calculating the charge from the reduction peak for the Au oxide in the characterization CV and dividing it by the specific charge of one monolayer of Au (390 μ C cm⁻²).³⁷

The morphology and the composition of the different NpAu catalysts was checked by using the ETD detector (for SEM imaging) and the EDS detector (for Energy Dispersive X-ray Spectroscopy) of Apreo SEM (Thermo Scientific), respectively. After the electrochemical measurements, the NpAu samples were rinsed with MilliQ water and upon drying placed on a custom-made SEM holder for analysis.

5.2.3. Differential electrochemical mass spectrometry (DEMS) measurements

All the electrochemical measurements were performed with a home-built DEMS set-up (described in detail elsewhere).³² The ionization potential of the ion source of the mass spectrometer was set to -27.5 V (vs the potential of the formation room and +72.5 V vs ground) to circumvent the fragmentation of CO₂ at mass 28 (CO⁺) which allowed for the ionic signal at mass 28 to be used for the quantitative detection of CO formed during CO₂RR.³² To achieve well-defined mass transport conditions during the CO₂RR studies, a dual thin layer cell with a flow cell geometry was used (described in detail elsewhere).³⁸ We note here that for all the CO₂RR measurements, instead of using CO₂ saturated conditions, we employed a 1:1 (vol.) ratio of Ar/CO₂. This was done in order to prevent bubble attachment in the microfluidic channels of the dual thin layer cell, which becomes unavoidable when high concentration of hydrophobic CO₂ is dissolved in the electrolyte, as it tends to salt out in the thin channels of the cell body (made of Kel-F) over the course of the measurements.¹⁷ Hence for

all the CO₂RR measurements, an Ar/CO₂ saturated electrolyte (0.1 M NaHCO₃ plus 0.4 M NaClO₄ or 0.1 M NaHCO₃ plus 2 M NaClO₄) was introduced into the top (electrochemical) compartment of the cell where the working electrode (NpAu or Flat Au) was placed. For all the CO₂RR studies, the cyclic voltammograms were recorded between -0.75 V and -1.2 V vs normal hydrogen electrode (NHE; where a Ag/AgCl reference was used in all the measurements and the potentials were converted to the NHE scale afterwards) at a scan rate of 5 mVs⁻¹. The flow of the electrolyte (5 μLs⁻¹) through the cell was achieved with a syringe pump which swept the electrolyte (along with the electrochemical products and the unreacted analyte) from the top compartment to the bottom compartment of the cell which was interfaced with the vacuum of the mass spectrometer with the help of a porous Teflon membrane and a porous steel frit. The products of HER and CO₂RR were detected at the ionic signal at *m/z* 2 (H₂) and *m/z* 28 (CO), respectively. The ionic current obtained with the DEMS is related to the molar flux of the species by eqn. 1:

$$I_{\text{ionic}(i)} = \frac{dn(i)}{dt} \times K^* \quad (1)$$

where, $I_{\text{ionic}(i)}$ is the ionic current due to a species i , $\frac{dn(i)}{dt}$ is the molar flux of i into the mass spectrometer and K^* is a proportionality constant that reflects the sensitivity of the DEMS setup for i , which among other things is dependent on the ionization probability of i .³⁹ Here, the molar flux $\frac{dn(i)}{dt}$ is associated to the Faradaic current $I_{\text{Faradaic}(i)}$ by Faraday's law:

$$\frac{dn(i)}{dt} = \frac{I_{\text{Faradaic}(i)} \times N}{zF} \quad (2)$$

where N is the transfer efficiency of species i in a given cell geometry/assembly, z is the number of electrons involved in the electrochemical reaction and F is Faraday's constant. By substituting eqn. 2 into eqn. 1, we obtain:

$$I_{\text{ionic}(i)} = \frac{I_{\text{Faradaic}(i)} \times K^\circ}{zF} \quad (3)$$

where $K^\circ = K^* \times N$ is the cell calibration constant. In order to quantify the ionic signals, in every new cell assembly, calibration measurements need to be performed to determine the value of the calibration constant (K°) for H₂ and CO, respectively. For the calibration of H₂, measurements were performed under Ar saturated conditions with a blank electrolyte (0.1 M phosphate buffer and 0.4 M NaClO₄), as shown in Fig. S1a of Appendix D, where H₂ was evolved with 100% Faradaic efficiency (scan window of -0.6 V to -1.8 V vs Ag/AgCl at a scan rate of 5 mVs⁻¹). Using eqn. 3, $K_{\text{H}_2}^\circ$ was determined ($K_{\text{H}_2}^\circ = \frac{I_{\text{ionic}(\text{H}_2)} \times zF}{I_{\text{Faradaic}(\text{HER})}}$). For the calibration of CO, bulk CO oxidation was performed under CO saturated conditions with a blank electrolyte (0.1

M phosphate buffer and 0.4 M NaClO₄), as shown in Fig. S1b in Appendix D, where CO₂ was evolved with 100% Faradaic efficiency (scan window of -0.6 V to 0.5 V vs Ag/AgCl at a scan rate of 5 mVs⁻¹). Using eqn. 3, K_{CO}° was determined ($K_{CO}^{\circ} = \frac{I_{\text{Ionic (CO)}} \times zF}{I_{\text{Faradaic (CO ox.)}}$). Next, the ionic signals obtained at mass 2 and mass 28 during CO₂RR measurements could be used to calculate the partial current densities for H₂ and CO formation from eqn. 3. The Faradaic selectivity for CO₂RR and HER could then be calculated as follows:

$$FE_{CO_2RR} = \frac{I_{\text{Faradaic (CO)}} \times 100}{(I_{\text{Faradaic (CO)}} + I_{\text{Faradaic (H}_2)})} = \frac{I_{\text{Faradaic (CO)}}}{I_{\text{Measured}}} \times 100 \quad (4)$$

and,

$$FE_{HER} = \frac{I_{\text{Faradaic (H}_2)}}{(I_{\text{Faradaic (CO)}} + I_{\text{Faradaic (H}_2)})} \times 100 = 100 - FE_{CO_2RR} \quad (5)$$

In summary, by determining the calibration constants ($K_{H_2}^{\circ}$ and K_{CO}°), the online DEMS technique can be used for quantitatively studying CO₂RR on NpAu catalysts.

5.3. Results and discussion

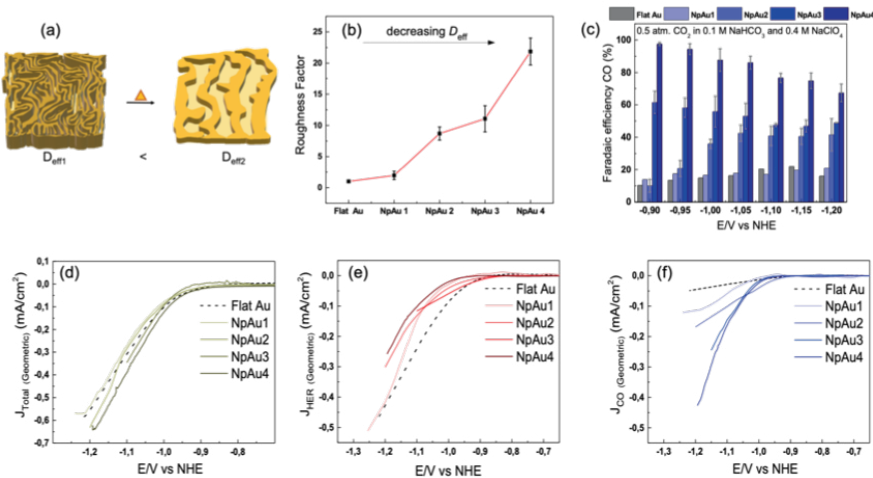


Figure 6 (a) Schematic representation of NpAu catalysts illustrating how thermal coarsening results in the self-similar growth of the NpAu ligaments such that the pore diameter increases and the effective pore length decreases with increasing coarsening. This also results in a decreasing effective diffusive flow (D_{eff}) through the porous channels. (b) Roughness factors of different Au catalysts as determined by the ECSA values obtained from Fig. S2 (Appendix D), where D_{eff} decreases with increasing roughness. (c) Faradaic efficiency for CO formation on the different Au catalysts at different potentials (vs. NHE), as obtained from the DEMS measurements by using eqn. 4. (d) Total geometric current density on different Au catalysts during CO₂RR, (e) Partial geometric current density for HER

on different Au catalysts obtained from the ionic current at m/z 2 by using eqn. 3 and (f) Partial geometric current density for CO formation on different Au catalysts obtained from the ionic current at m/z 28 by using eqn. 3. All the measurements were done in 0.5 atm. CO_2 , in 0.1 M NaHCO_3 plus 0.4 M NaClO_4 containing electrolyte with a scan rate of 5 mVs^{-1} and a flow rate of $300 \mu\text{Lmin}^{-1}$.

In this work we studied the role of pore parameters (diameter and length) in tuning the activity/selectivity of electrochemical CO_2 reduction reaction (CO2RR). For this we prepared four different nanoporous Au (NpAu) samples via an electrochemical alloying-dealloying and subsequent thermal coarsening procedure, as outlined in section 2.2. Electrochemical characterization (see Fig. S2 in Appendix D) and scanning electron microscopy (SEM) images (see Fig. S3 in Appendix D) show that the thermal treatment of the NpAu samples results in the systematic decrease of the ligament diameter of the porous network. Moreover, in accordance with the previous studies,^{34-36, 40} we expect that thermal coarsening of the NpAu samples results in a more or less self-similar growth of the porous structures, i.e. an increase in the pore/ligament diameter and a corresponding decrease in pore length (also shown in Fig. 1a). Thus, in going from NpAu1 to NpAu4 the pore diameter decreases and the pore length increases. This in turn leads to decreasing effective diffusion (D_{eff}) through the porous channels due to the generation of additional diffusional gradients along the length of the pores.⁴¹ In Fig. 1b, we show the roughness factor of the different NpAu catalysts with respect to the polycrystalline Au electrode (Flat Au), calculated by using the as-determined ECSA (shown in Fig. S2 in Appendix D) of the different surfaces. As expected, the roughness factor of the catalyst surface increases with decreasing thermal coarsening. We note here that in all the NpAu samples, we detect around 7 atomic-% of residual Ag with EDX spectroscopy (shown in Fig. S3 in Appendix D). Thereafter, we studied CO2RR at these catalysts by using the online DEMS, where the ionic signals at m/z 2 and 28 were used to quantify the amount of H_2 and CO, respectively, evolved during CO2RR on the different catalyst surfaces (for further details see section 2.3). In Fig. 1c, we show the Faradaic efficiency for CO2RR on the different Au catalysts as derived from the DEMS measurements by using eqn. 4. In agreement with the previous studies on nanoporous Au catalysts,^{17, 22, 24} we see that the Faradaic efficiency for CO2RR increases with increasing roughness factor of the catalysts, i.e. with decreasing pore diameter and increasing pore length (Flat Au < NpAu1 < NpAu2 < NpAu3 < NpAu4). Remarkably, in the case of NpAu4, we achieve ~100% Faradaic efficiency for CO2RR even with 0.5 atm. of CO_2 . Here, the sub-optimal reaction conditions for CO2RR (0.5 atm. of CO_2 instead of 1 atm. CO_2) also explain the uncharacteristically low Faradaic efficiency for CO in the case of flat polycrystalline Au (Flat Au) catalyst. We note that in addition to the low partial pressure of CO_2 , the high Na^+ ion concentration in the electrolyte also contributes to a lower Faradaic selectivity towards CO2RR. We have shown previously with our RRDE technique that with increasing cation concentration in the electrolyte (at a fixed bicarbonate concentration) the Faradaic efficiency towards CO2RR decreases on flat Au polycrystalline electrodes.^{8, 33} In that work we achieved Faradaic selectivities close to 60% in 1 atm. CO_2 with 0.5 M Na^+ ion

containing electrolytes.³³ Hence, the Faradaic efficiency of around 20% in the case of flat Au catalysts (as shown in Fig. 1c) can be attributed to the combination of low CO₂ partial pressure (0.5 atm.) and a high cation concentration in the electrolyte (0.5 M) in our measurements.

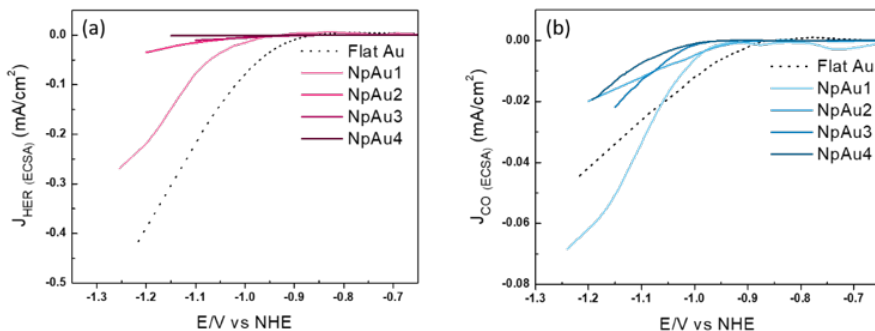


Figure 2 (a) Partial specific current density for HER on different Au catalysts (obtained from the data used in Fig. 1e) and (b) Partial specific current density for CO formation on different Au catalysts (obtained from the data used in Fig. 1f). All the measurements were done in 0.5 atm. CO₂ in 0.1 M NaHCO₃ plus 0.4 M NaClO₄ containing electrolyte with a scan rate of 5 mVs⁻¹ and a flow rate of 300 μLmin⁻¹.

In Fig. 1d, we plot the experimentally measured total geometric current density, and in Figs. 1e and 1f we plot the corresponding partial geometric current densities for HER and CO₂RR calculated from the ionic currents measured at mass 2 and 28, respectively. We see that the geometric current density for HER decreases in going from Flat Au to NpAu4 (as shown in Fig. 1e) and hence, the enhancement in the CO₂RR Faradaic efficiency with the increasing roughness of the NpAu catalysts can, at least in part, be attributed to the suppression of competing HER reaction.^{21-22, 25} We have shown previously with the RRDE technique^{8, 33} that in bicarbonate containing electrolytes, HER can either be mediated by HCO₃⁻ ions (eqn. 6) or by H₂O molecules (eqn. 7):



and that these two branches show a divergent dependence on electrolyte pH. It was observed that while HER due to the discharge of HCO₃⁻ ions decreased with increasing pH, for water reduction reaction the opposite takes place.^{8, 42} Moreover, the branch of HER that competes with CO₂RR is largely dependent on the identity of the electrolyte that is employed in the measurements. In brief, the dominating branch of HER switches from water-mediated to HCO₃⁻-mediated with increasing bicarbonate/cation concentration in the electrolyte.^{8, 33} In accordance with these results, under the experimental conditions of our measurements (0.1 M NaHCO₃ plus 0.4 M NaClO₄) we expect bicarbonate mediated HER to be the main branch of hydrogen evolution that competes with CO₂RR.

Moreover, we expect the partial current due to HER to be sensitive to the local pH which, among other things, is determined by the morphology of the electrode surface. In the case of nanoporous catalysts, with increasing catalyst roughness, the effective diffusion through the porous channels will decrease. As a result, the local pH will increase at the catalyst surface due to the hindered mass transport of locally generated OH^- ions away from the electrode and at the same time, due to the hindered mass transport of HCO_3^- ions and CO_2 (aq.) to the surface. Hence, we attribute the decreasing HER activity with increasing catalyst roughness, as observed in Fig. 1e, to the corresponding increase in local pH in going from Flat Au to NpAu4. This is because the increasing local alkalinity accelerates the homogeneous consumption of bicarbonate near the electrode surface:



which in turn leads to a paucity of HCO_3^- ions for participating in the HER reaction.⁸ We note here that the partial current density for HER decreases both with respect to the geometric surface area (Fig. 1e) and with respect to the ECSA of the catalysts (Fig. 2a).

On the other hand, for CO₂RR, we see that the partial geometric current density increases with increasing pore length and decreasing pore diameter i.e. with increasing catalyst roughness, as shown in Fig. 1f. This is partially counter-intuitive since increasing local pH at the catalyst surface should also lead to a corresponding increase in the homogeneous consumption of CO_2 :



Hence, similar to HER, we would expect that increasing local pH due to increasing catalyst roughness will result in the suppression of CO₂RR. However, we see the opposite trend in Fig. 1f. This can be rationalized by the fact that the rate of homogeneous CO_2 consumption ($k_{\text{q}}^+ = 2.23 \times 10^3 \text{ Kg mol}^{-1} \text{ s}^{-1}$) is six orders of magnitude slower than the rate of homogeneous HCO_3^- consumption ($k_{\text{q}}^+ = 6 \times 10^9 \text{ Kg mol}^{-1} \text{ s}^{-1}$).^{27, 43} Consequently, the mass transport rate of CO_2 (aq.) to the interface can outpace the rate of its homogeneous reaction, especially under the conditions of convection control (e.g. rotation control, flow control).

To validate this further, we performed measurements where the different NpAu catalysts were first exposed to a CO_2 -saturated electrolyte and thereafter, currents due to CO formation were measured under Ar-saturated conditions (shown in Fig. 3). Interestingly, we see that all the NpAu catalysts show appreciable currents due to CO formation even in the absence of a continuous CO_2 supply (Fig. 3a). Moreover, these currents persist after ten subsequent scans (Fig. 3b) even on the catalyst with the highest roughness factor (NpAu4). Together, these results show that despite the increasing mass transport limitations and with this an increasing local alkalinity in the porous channels, NpAu catalysts are able to furnish an appreciable reservoir of CO_2 . Moreover, the local CO_2 reservoir is able to take part in CO₂RR without completely getting homogeneously consumed. In fact, with

increasing catalyst roughness we obtain higher currents for CO formation (Fig. 3a), thus indicating that with increasing pore length we generate a larger CO₂ reservoir at the catalyst surface.

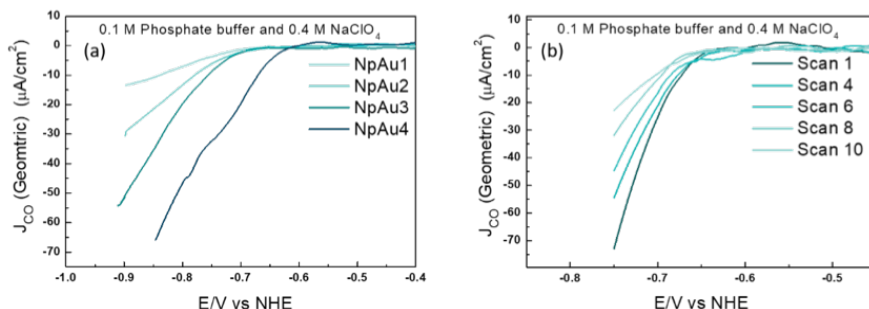


Figure 3 (a) Partial geometric current density for CO formation on different Au catalysts (Flat Au, NpAu1, NpAu2, NpAu3 and NpAu4) where the dual thin layer cell (along with the catalysts) was first flushed with a CO₂ sat. 0.1 M Phosphate buffer plus 0.4 M NaClO₄ containing electrolyte; the measurements were done in Ar sat. conditions (with the same electrolyte) after flushing the cell with Ar sat. electrolyte for 10 min. and in (b) the 10 subsequent scans under the same experimental conditions described above on NpAu4 catalyst. Scan rate: 5 mVs⁻¹ and flow rate: 300 μ Lmin⁻¹.

In Table S1 (Appendix D), we give calculated theoretical currents as limited by mass transport for CO₂RR and they are much larger than the experimentally obtained current densities on any of the catalysts employed in our measurements (see Fig. 1f). This shows that under the flow conditions of our measurements (300 μ L min⁻¹), we are below any mass transport limitations for CO₂RR. Hence, the slow homogenous reaction of CO₂ and the lack of any mass transport limitation for CO₂RR explains why the CO₂RR current density does not decrease in going from Flat Au to NpAu4. On the other hand, the availability of electrochemically active sites at the catalyst surface increases with increasing catalyst roughness, as a higher surface roughness automatically comes with more surface defects which are commonly considered sites of increased reactivity. Together, the increase in the number of active sites and/or increased surface defects explains the observed enhancement in CO₂RR current density with increasing catalyst roughness.

Interestingly, unlike HER, we see diverging trends when the partial geometric current density (current with respect to the geometric surface area) and partial specific current density (current with respect to the ECSA) for CO₂RR are compared. In Fig. 2b we plot the partial specific current density for CO₂RR and we see that while it increases in going from Flat Au to NpAu1, a further increase in the catalyst roughness results in a lower specific current density for CO₂RR, with no clear trend between NpAu2 to NpAu4. Given that the CO₂RR activity is tied to the availability of active sites at the catalyst surface, the fact that it scales with the geometric surface area (Fig. 1f) but does not scale with the ECSA (Fig. 2b) suggests that not the entire surface of the NpAu catalysts participates in electrocatalysis.

This is understandable because the thin porous channels of NpAu catalysts generate an additional electrical resistance which scales with increasing pore length.⁴⁴⁻⁴⁵ As shown in Fig. 4a, at the orifices of the porous sample an

Ohmic drop originates solely due to the electrolyte resistance (R_{sol}) which can be mitigated by positive feedback compensation. However, the distributed capacitance (C) and the charge transfer resistance (R_{ct}) that represent the electrocatalytic reaction at the interface are coupled to additional uncompensated resistance (r_{pore}) along the length of the pores. This means that there is an additional barrier for the current to pass through the pores. Hence, with increasing distance from the orifice it becomes more difficult for the porous channels to participate in the Faradaic reaction. Essentially, these cumulative uncompensated Ohmic drop effects can result in an electrocatalytically inactive zone at the bottom of the porous channels which results in the decoupling of the ECSA and the specific activity at these catalysts. Hence, the observed trend for the partial specific CO₂RR current density as a function of the catalyst roughness (shown in Fig. 2b) can be attributed to the fact that the increasing availability of active sites for CO₂RR from Flat Au to NpAu4 catalysts, is countered by a corresponding increase in the Ohmic drop effects with increasing pore length. As a result, while at relatively low catalyst roughness (NpAu1; Fig. 2b) the increased availability of the active sites overshadows these Ohmic drop effects, as the catalyst roughness (pore length) increases further and Ohmic drop effects dominate the overall activity for CO₂RR at NpAu catalysts.

In order to validate the presence of the Ohmic drop effects at NpAu catalysts, we performed additional CO₂RR measurements (shown in Fig. 4) by increasing the supporting electrolyte concentration in the system. Since increasing supporting electrolyte concentration will increase the conductivity of the electrolyte, it should result in a proportional decrease in the Ohmic drop (r_{pore}) along the porous channels. Hence, we would expect the partial specific current density for CO₂RR to increase with increasing ionic strength of the electrolyte at NpAu catalysts.

The results in Fig. 4 confirm this expectation: on all the NpAu catalysts (Fig. 4d and 4f) the partial specific current density for CO₂RR increases with increasing concentration of the supporting electrolyte, while on flat Au catalysts, the partial current density for CO₂RR slightly decreases with the increasing ionic strength (Fig. 4b). This confirms that Ohmic drop effects along the porous channels play a significant role in determining the obtained specific current densities for CO₂RR on NpAu catalysts.

Notably, the presence of these Ohmic drop effects on NpAu catalysts also implies even though the total amount of CO₂ inside the pores increases with increasing pore size, this CO₂ does not necessarily react everywhere within the pore. Because it takes longer for CO₂ to diffuse out of a more porous electrode, this "unreacted" CO₂ can lead to an enhanced current for CO₂RR under non-steady-state conditions, until the pores are depleted of CO₂ (as is the case in the results shown in Fig. 3b). Hence, it should be kept in mind that this local CO₂ reservoir can affect the observed CO₂RR activities on NpAu catalysts, however it is a non-steady-state effect.

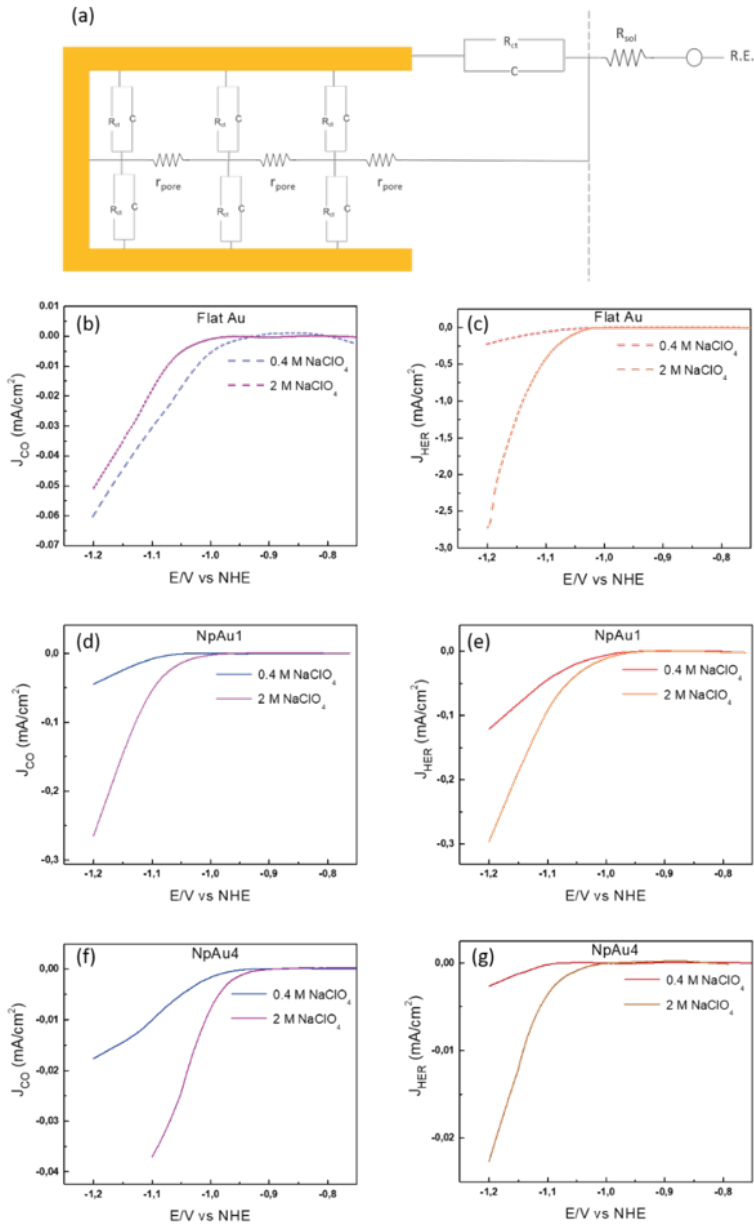


Figure 4 (a) Schematic representation of the circuit at NpAu catalysts where R_{sol} is the solution resistance between the reference electrode and the working electrode, R_{ct} , C are the distributed charge transfer resistance and the distributed capacitance associated with the Faradaic reaction at the catalyst surface and r_{pore} is the distributed resistance of the pore per unit length and per unit surface area of the porous film. Partial specific current density for CO formation as obtained from the ionic current at m/z 28 by using eqn. 3 with 0.5 atm. CO₂ in 0.1 M NaHCO₃ plus 0.4 M NaClO₄ containing electrolyte (blue) and in 0.1 M NaHCO₃ plus 2 M NaClO₄ containing electrolyte (pink) on (b) Flat Au, (d) NpAu1 and (f) NpAu 4. Partial current density for HER as obtained from the ionic current at m/z 2 by using eqn. 3 with 0.5 atm. CO₂ in 0.1 M NaHCO₃ plus 0.4 M NaClO₄ containing electrolyte (red) and in 0.1 M NaHCO₃ plus 2 M NaClO₄ containing electrolyte (orange) on (c) Flat Au, (e) NpAu1 and (g) NpAu 4. Scan rate: 5 mV s⁻¹ and flow rate: 300 μ L min⁻¹.

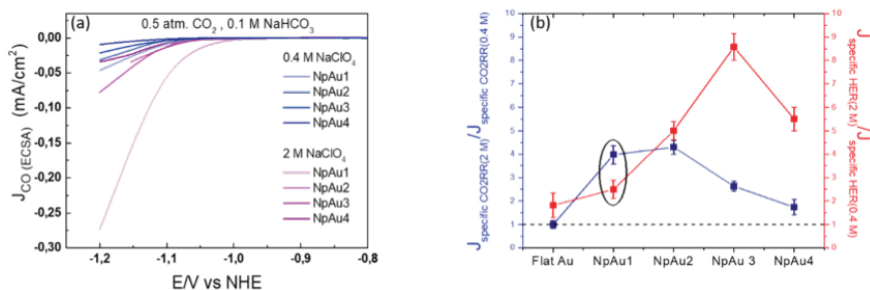


Figure 5 (a) Comparison of the partial specific current density for CO formation on different NpAu catalysts (NpAu1 to NpAu4) as obtained from the ionic current at m/z 28 by using eqn. 3 with 0.5 atm. CO₂ and 0.1 M NaHCO₃ in 0.4 M NaClO₄ (blue) and 2 M NaClO₄ (pink) containing electrolyte at a scan rate of 5 mVs⁻¹ and a flow rate of 300 μL min⁻¹. (b) Ratio of the specific current density in 2 M NaClO₄ containing electrolyte and 0.4 M NaClO₄ containing electrolyte at a fixed NaHCO₃ (0.1 M) concentration towards CO₂RR (blue) and HER (red) at -1.2 V vs NHE on different Au catalysts. Here a ratio greater than 1 indicates a higher specific current density in 2 M NaClO₄ containing electrolyte compared to 0.4 M NaClO₄ containing electrolyte.

Moreover, we note that unlike CO₂RR, the HER activity increases with increasing concentration of the supporting electrolyte (Fig. 4c, 4e and 4f) regardless of the catalyst surface. This agrees with our previous studies on flat polycrystalline Au electrodes where we have shown that HER shows a positive reaction order for the cation concentration in the electrolyte. In that work we ascribed this to the changes in the double layer composition with changing cation concentration in the electrolyte.^{33, 42} Moreover, in agreement with the current study, we also observed near-zero reaction order in cation concentration for CO₂RR at flat polycrystalline Au electrodes, which we ascribed to the enhancement of HER at the expense of CO₂RR with the increasing cation concentration in the electrolyte.³³

Additionally, in Fig. 5a we plot the partial specific current density for CO₂RR on the different Au catalysts in 0.4 M and 2 M NaClO₄ containing electrolyte and we see that while in the case of NpAu1, the increasing ionic strength of the electrolyte is able to compensate for the Ohmic drop effects completely, for the other NpAu catalysts, only a slight improvement is observed. This shows that by increasing the total ionic strength of the electrolyte from 0.6 M to 2.1 M we are still not able to completely compensate for the Ohmic drop effects, especially in catalysts with higher roughness factors. Nevertheless, increasing the ionic strength of the electrolyte does improve the obtained specific current density for CO₂RR on all the NpAu catalysts. Importantly, there is a trade-off here between the increase in the CO₂RR activity (specific current density) and CO₂RR selectivity (Faradaic efficiency) with the increasing ionic strength of the electrolyte. As shown in Fig. 5b, the enhancement in the HER activity with increasing ionic strength overshadows the enhancement in the CO₂RR activity on all the catalysts except NpAu1. This is also reflected in the Faradaic selectivity for CO₂RR, as shown in Fig. 6, where we see that except NpAu1, in all the other NpAu catalysts, the Faradaic selectivity towards CO₂RR decreases with increasing ionic strength of the electrolyte. Hence, these results show that we can partially compensate for the Ohmic drop effects occurring inside nanoporous

catalysts by increasing the supporting electrolyte concentration which enables higher CO₂RR activity. However, depending on the roughness/porosity of the catalysts, the concomitant enhancement of the competing HER reaction with increasing cation concentration in the electrolyte can overshadow the enhancement in the CO₂RR activity, which may therefore result in a lower Faradaic selectivity for CO₂RR.

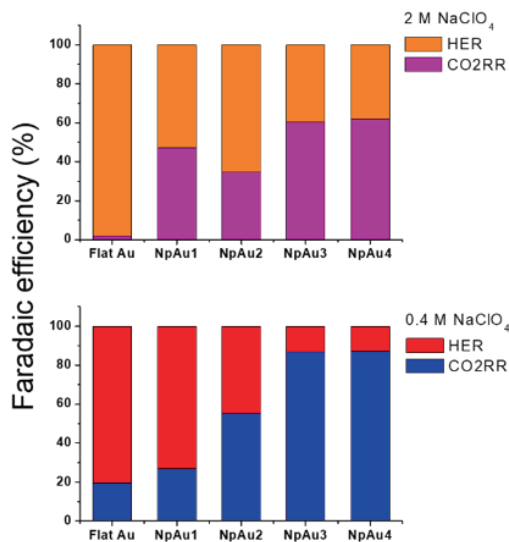


Figure 6 Comparison of the Faradaic efficiency for CO₂RR and HER in 2 M NaClO₄ containing electrolyte (top panel) and 0.4 M NaClO₄ containing electrolyte (bottom panel) at a fixed NaHCO₃ concentration (0.1 M) at -1 V vs NHE in 0.5 atm. CO₂.

5.4. Conclusions

We studied the role of pore parameters (diameter and pore length) in tuning CO₂RR on nanoporous Au (NpAu) catalysts by using the online DEMS technique. We find a higher CO₂RR Faradaic selectivity is achieved with increasingly narrower and longer pores (increasing catalyst roughness), and in the most extreme case (NpAu4) we achieve near 100% efficiency even in an under-saturated CO₂ solution. This enhancement is attributed to two factors, the enhancement of the CO₂RR currents due to the presence of more active sites with increasing catalyst roughness and the suppression of competing HER reaction due to the increasing local pH at the catalyst surface. Interestingly, the partial specific current density for CO₂RR shows a more complicated relationship with catalyst roughness than the partial geometric current density, as it does not scale with increasing catalyst roughness. We show that this is due to the presence of additional Ohmic drop across the length of the thin porous channels which plays an important role in determining the specific activity of nanoporous catalysts.

These results have important implications as they show that with increasing catalyst porosity, there is a trade-off between the number of active sites and the Ohmic drop effects at the catalyst surface. These two effects counter-act each other for tuning the CO₂RR activity. This is relevant as previous studies on nanoporous catalysts have not considered these offsets

and some of the discrepancies in the current literature on the role of pore parameters in tuning CO₂RR activity can be explained by the presence of these counter-acting effects. Additionally, these results also imply that the increase in the local alkalinity due to the generation of OH⁻ ions via CO₂RR and HER, will diminish along the length of the pores due to the increasing Ohmic drop effects. Hence, exclusion of these effects can lead to erroneous estimation/modelling of local pH gradients at the surface of nanoporous catalysts.

In conclusion, this work provides important new insights into the intricate role of local pH gradients and intrinsic geometric effects in tuning the activity/selectivity of CO₂RR at the surface of nonporous Au catalysts and we believe that these insights will be instrumental in the rational optimization of CO₂RR on practical electrode geometries.

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References

1. Nitopi, S.; Bertheussen, E.; Scott, S. B.; Liu, X.; Engstfeld, A. K.; Horch, S.; Seger, B.; Stephens, I. E. L.; Chan, K.; Hahn, C.; Nørskov, J. K.; Jaramillo, T. F.; Chorkendorff, I., Progress and Perspectives of Electrochemical CO₂ Reduction on Copper in Aqueous Electrolyte. *Chemical Reviews* **2019**, *119* (12), 7610-7672.
2. Birdja, Y. Y.; Pérez-Gallent, E.; Figueiredo, M. C.; Göttle, A. J.; Calle-Vallejo, F.; Koper, M. T. M., Advances and challenges in understanding the electrocatalytic conversion of carbon dioxide to fuels. *Nature Energy* **2019**, *4* (9), 732-745.
3. Liu, A.; Gao, M.; Ren, X.; Meng, F.; Yang, Y.; Gao, L.; Yang, Q.; Ma, T., Current progress in electrocatalytic carbon dioxide reduction to fuels on heterogeneous catalysts. *Journal of Materials Chemistry A* **2020**, *8* (7), 3541-3562.
4. Zhao, S.; Jin, R.; Jin, R., Opportunities and Challenges in CO₂ Reduction by Gold- and Silver-Based Electrocatalysts: From Bulk Metals to Nanoparticles and Atomically Precise Nanoclusters. *ACS Energy Letters* **2018**, *3* (2), 452-462.
5. Burdyny, T.; Smith, W. A., CO₂ reduction on gas-diffusion electrodes and why catalytic performance must be assessed at commercially-relevant conditions. *Energy & Environmental Science* **2019**, *12* (5), 1442-1453.
6. Kas, R.; Hummadi, K. K.; Kortlever, R.; de Wit, P.; Milbrat, A.; Luiten-Olieman, M. W. J.; Benes, N. E.; Koper, M. T. M.; Mul, G., Three-dimensional porous hollow fibre copper electrodes for efficient and high-rate electrochemical carbon dioxide reduction. *Nature Communications* **2016**, *7*, 10748.
7. Kas, R.; Kortlever, R.; Yilmaz, H.; Koper, M. T. M.; Mul, G., Manipulating the Hydrocarbon Selectivity of Copper Nanoparticles in CO₂ Electroreduction by Process Conditions. *ChemElectroChem* **2015**, *2* (3), 354-358.
8. Goyal, A.; Marcandalli, G.; Mints, V. A.; Koper, M. T. M., Competition between CO₂ Reduction and Hydrogen Evolution on a Gold Electrode under Well-Defined Mass Transport Conditions. *Journal of the American Chemical Society* **2020**, *142* (9), 4154-4161.
9. Wang, Y.; Liu, J.; Wang, Y.; Al-Enizi, A. M.; Zheng, G., Tuning of CO₂ Reduction Selectivity on Metal Electrocatalysts. *Small* **2017**, *13* (43), 1701809.
10. Zhang, L.; Wang, Z. Y.; Mehio, N. D.; Jin, X. B.; Dai, S., Thickness- and Particle-Size-Dependent Electrochemical Reduction of Carbon Dioxide on Thin-Layer Porous Silver Electrodes. *ChemSusChem* **2016**, *9* (5), 428-432.

11. Lu, Q.; Rosen, J.; Zhou, Y.; Hutchings, G. S.; Kimmel, Y. C.; Chen, J. G.; Jiao, F., A selective and efficient electrocatalyst for carbon dioxide reduction. **2014**, *5*, 3242.
12. Kumar, B.; Atla, V.; Brian, J. P.; Kumari, S.; Nguyen, T. Q.; Sunkara, M.; Spurgeon, J. M., Reduced SnO₂ Porous Nanowires with a High Density of Grain Boundaries as Catalysts for Efficient Electrochemical CO₂-into-HCOOH Conversion. *Angewandte Chemie International Edition* **2017**, *56* (13), 3645-3649.
13. Rosen, J.; Hutchings, G. S.; Lu, Q.; Rivera, S.; Zhou, Y.; Vlachos, D. G.; Jiao, F., Mechanistic Insights into the Electrochemical Reduction of CO₂ to CO on Nanostructured Ag Surfaces. *ACS Catalysis* **2015**, *5* (7), 4293-4299.
14. Jiang, X.; Cai, F.; Gao, D.; Dong, J.; Miao, S.; Wang, G.; Bao, X., Electrocatalytic reduction of carbon dioxide over reduced nanoporous zinc oxide. *Electrochemistry Communications* **2016**, *68* (Supplement C), 67-70.
15. Peng, Y.; Wu, T.; Sun, L.; Nsanzimana, J. M. V.; Fisher, A. C.; Wang, X., Selective Electrochemical Reduction of CO₂ to Ethylene on Nanopores-Modified Copper Electrodes in Aqueous Solution. *ACS Applied Materials & Interfaces* **2017**, *9* (38), 32782-32789.
16. Clark, E. L.; Bell, A. T., Direct Observation of the Local Reaction Environment during the Electrochemical Reduction of CO₂. *Journal of the American Chemical Society* **2018**, *140* (22), 7012-7020.
17. Bondue, C. J.; Graf, M.; Goyal, A.; Koper, M. T. M., Suppression of Hydrogen Evolution in Acidic Electrolytes by Electrochemical CO₂ Reduction. *Journal of the American Chemical Society* **2021**, *143* (1), 279-285.
18. Park, S.; Kim, H. C.; Chung, T. D., Electrochemical analysis based on nanoporous structures. *Analyst* **2012**, *137* (17), 3891-3903.
19. Luc, W.; Jiao, F., Nanoporous Metals as Electrocatalysts: State-of-the-Art, Opportunities, and Challenges. *ACS Catalysis* **2017**, *7* (9), 5856-5861.
20. Galinsky, M.; Chal, U.; Breitung, C., The Impact of Microstructure Geometry on the Mass Transport in Artificial Pores: A Numerical Approach. *Modelling and Simulation in Engineering* **2014**, *2014*, 7.
21. Yoon, Y.; Hall, A. S.; Surendranath, Y., Tuning of Silver Catalyst Mesostructure Promotes Selective Carbon Dioxide Conversion into Fuels. *Angewandte Chemie International Edition* **2016**, *55* (49), 15282-15286.
22. Hall, A. S.; Yoon, Y.; Wuttig, A.; Surendranath, Y., Mesostructure-Induced Selectivity in CO₂ Reduction Catalysis. *Journal of the American Chemical Society* **2015**, *137* (47), 14834-14837.
23. Luo, W.; Zhang, J.; Li, M.; Züttel, A., Boosting CO Production in Electrocatalytic CO₂ Reduction on Highly Porous Zn Catalysts. *ACS Catalysis* **2019**, *9* (5), 3783-3791.
24. Welch, A. J.; DuChene, J. S.; Tagliabue, G.; Davoyan, A.; Cheng, W.-H.; Atwater, H. A., Nanoporous Gold as a Highly Selective and Active Carbon Dioxide Reduction Catalyst. *ACS Applied Energy Materials* **2019**, *2* (1), 164-170.
25. Chen, C.; Zhang, B.; Zhong, J.; Cheng, Z., Selective electrochemical CO₂ reduction over highly porous gold films. *Journal of Materials Chemistry A* **2017**, *5* (41), 21955-21964.
26. Wuttig, A.; Yoon, Y.; Ryu, J.; Surendranath, Y., Bicarbonate Is Not a General Acid in Au-Catalyzed CO₂ Electroreduction. *Journal of the American Chemical Society* **2017**, *139* (47), 17109-17113.
27. Dunwell, M.; Yang, X.; Setzler, B. P.; Anibal, J.; Yan, Y.; Xu, B., Examination of Near-Electrode Concentration Gradients and Kinetic Impacts on the Electrochemical Reduction of CO₂ using Surface-Enhanced Infrared Spectroscopy. *ACS Catalysis* **2018**, *8*, 3999-4008.
28. Dunwell, M.; Lu, Q.; Heyes, J. M.; Rosen, J.; Chen, J. G.; Yan, Y.; Jiao, F.; Xu, B., The Central Role of Bicarbonate in the Electrochemical Reduction of Carbon Dioxide on Gold. *Journal of the American Chemical Society* **2017**, *139* (10), 3774-3783.
29. Wuttig, A.; Yaguchi, M.; Motobayashi, K.; Osawa, M.; Surendranath, Y., Inhibited proton transfer enhances Au-catalyzed CO₂-to-fuels selectivity. *Proceedings of the National Academy of Sciences* **2016**, *113* (32), E4585-E4593.
30. Zhang, B. A.; Ozel, T.; Elias, J. S.; Costentin, C.; Nocera, D. G., Interplay of Homogeneous Reactions, Mass Transport, and Kinetics in Determining Selectivity of the Reduction of CO₂ on Gold Electrodes. *ACS Central Science* **2019**, *5* (6), 1097-1105.

31. Clark, E. L.; Resasco, J.; Landers, A.; Lin, J.; Chung, L.-T.; Walton, A.; Hahn, C.; Jaramillo, T. F.; Bell, A. T., Standards and Protocols for Data Acquisition and Reporting for Studies of the Electrochemical Reduction of Carbon Dioxide. *ACS Catalysis* **2018**, *8* (7), 6560-6570.
32. Bondue, C. J.; Koper, M. T. M., A DEMS approach for the direct detection of CO formed during electrochemical CO₂ reduction. *Journal of Electroanalytical Chemistry* **2020**, 113842.
33. Marcandalli, G.; Goyal, A.; Koper, M. T. M., Electrolyte Effects on the Faradaic Efficiency of CO₂ Reduction to CO on a Gold Electrode. *ACS Catalysis* **2021**, *11* (9), 4936-4945.
34. Graf, M.; Vonbun-Feldbauer, G. B.; Koper, M. T. M., Direct and Broadband Plasmonic Charge Transfer to Enhance Water Oxidation on a Gold Electrode. *ACS Nano* **2021**, *15* (2), 3188-3200.
35. Ji, C.; Searson, P. C., Fabrication of nanoporous gold nanowires. *Applied Physics Letters* **2002**, *81* (23), 4437-4439.
36. Graf, M.; Jalas, D.; Weissmüller, J.; Petrov, A. Y.; Eich, M., Surface-to-Volume Ratio Drives Photoelectron Injection from Nanoscale Gold into Electrolyte. *ACS Catalysis* **2019**, *9* (4), 3366-3374.
37. Łukaszewski, M.; Soszko, M.; Czerwiński, A., Electrochemical Methods of Real Surface Area Determination of Noble Metal Electrodes – an Overview. *International Journal of Electrochemical Science* **2016**, *11*, 4442-4469.
38. Abd-El-Latif, A. A.; Bondue, C. J.; Ernst, S.; Hegemann, M.; Kaul, J. K.; Khodayari, M.; Mostafa, E.; Stefanova, A.; Baltruschat, H., Insights into electrochemical reactions by differential electrochemical mass spectrometry. *TrAC Trends in Analytical Chemistry* **2015**, *70*, 4-13.
39. Baltruschat, H., Differential electrochemical mass spectrometry. *Journal of the American Society for Mass Spectrometry* **2004**, *15* (12), 1693-1706.
40. Ding, Y.; Kim, Y.-J.; Erlebacher, J., Nanoporous Gold Leaf: "Ancient Technology"/Advanced Material. *Advanced Materials* **2004**, *16* (21), 1897-1900.
41. Haensch, M.; Balboa, L.; Graf, M.; Silva Olaya, A. R.; Weissmüller, J.; Wittstock, G., Mass Transport in Porous Electrodes Studied by Scanning Electrochemical Microscopy: Example of Nanoporous Gold. *ChemElectroChem* **2019**, *6* (12), 3160-3166.
42. Goyal, A.; Koper, M. T. M., The Interrelated Effect of Cations and Electrolyte pH on the Hydrogen Evolution Reaction on Gold Electrodes in Alkaline Media. *Angewandte Chemie International Edition* **2021**, *60* (24), 13452-13462.
43. Schulz, K. G.; Riebesell, U.; Rost, B.; Thoms, S.; Zeebe, R. E., Determination of the rate constants for the carbon dioxide to bicarbonate inter-conversion in pH-buffered seawater systems. *Marine Chemistry* **2006**, *100* (1), 53-65.
44. Haverkort, J. W., A theoretical analysis of the optimal electrode thickness and porosity. *Electrochimica Acta* **2019**, *295*, 846-860.
45. Andrieux, C. P.; Costentin, C.; Di Giovanni, C.; Savéant, J.-M.; Tard, C., Conductive Mesoporous Catalytic Films. Current Distortion and Performance Degradation by Dual-Phase Ohmic Drop Effects. Analysis and Remedies. *The Journal of Physical Chemistry C* **2016**, *120* (38), 21263-21271.

