

## Biomimetic models of [NiFe] hydrogenase for electrocatalytic hydrogen evolution

Gezer, G.

### Citation

Gezer, G. (2017, October 10). *Biomimetic models of [NiFe] hydrogenase for electrocatalytic hydrogen evolution*. Retrieved from https://hdl.handle.net/1887/58770

Version:	Not Applicable (or Unknown)
License:	<u>Licence agreement concerning inclusion of doctoral thesis in the</u> <u>Institutional Repository of the University of Leiden</u>
Downloaded from:	https://hdl.handle.net/1887/58770

Note: To cite this publication please use the final published version (if applicable).

Cover Page



# Universiteit Leiden



The handle <u>http://hdl.handle.net/1887/58770</u> holds various files of this Leiden University dissertation

Author: Gezer, G. Title: Biomimetic models of [NiFe] hydrogenase for electrocatalytic hydrogen evolution Issue Date: 2017-10-10

#### **Propositions (Stellingen)**

#### Accompanying the thesis

#### Biomimetic Models of [NiFe] Hydrogenase for Electrocatalytic Hydrogen Evolution

- 1. Altering the backbone spacer of chelating ligands changes the electrocatalytic properties of the corresponding metal complexes. *This thesis, Chapter 2*
- 2. Unexpected results can give a new perspective and lead to the investigation of a different research subject. *This thesis, Chapter 4*
- 3. The different physical properties of sulfur and selenium do not have a significant effect on the electrochemical properties of the complexes with ligands that only differ in these donor atoms. *This thesis, Chapter 3 and 5*
- 4. Irreversibility of reduction processes makes the understanding of electrocatalytic reactions difficult. *This thesis, Chapter 3 and 5*
- 5. A difference of nearly 1 V between the reduction potentials of the thiolate and selenolate analogues of the complexes with a propyl spacer is not expected if the reduction potentials of the thiolate and selenolate analogues of the complexes with a xylyl spacer are the same.

Weber et al., Eur. J. Inorg. Chem. 2014, 148-155; This thesis, Chapter 2

- 6. One cannot state that a compound is active in electrocatalytic proton reduction without reporting the results of the appropriate blank experiments. *Wombwell et al., Chem. Eur. J.* **2015**, *21, 1-10 and Weber et al., J. Am. Chem. Soc.* **2012**, *134, 20745-20755*
- 7. Although cyclic voltammetry can be a very important tool to check the electrochemical activity of a compound, reporting a complex to be an excellent catalyst without any quantification of the results is not useful. *Yang et al., Eur. J. Inorg. Chem.* **2015**, 2965-2973
- 8. As nickel-ruthenium complexes generally are more stable than the corresponding nickel-iron complexes, they should be considered as alternatives in the design of new electrocatalysts for dihydrogen production. *Canaguier et al., Chem. Eur. J.* **2009**, *15*, *9350-9364; This thesis Chapter 5*
- 9. Time flies when you are in the lab but it stops when you are writing.
- 10. Preparing a suitable and reliable setup for GC measurements is more difficult than synthesizing a new catalyst.

Gamze Gezer *Leiden, October, 2017*